



# National Institute of Standards & Technology

## Certificate of Analysis

### Standard Reference Material 3127a

#### Spectrometric Standard Solution

#### Lanthanum

Batch Code 490712

This Standard Reference Material (SRM) is intended for use in atomic absorption spectrometry, inductively coupled plasma spectrometry, spectrophotometry, or any other analytical technique that requires aqueous standard solutions for calibrating instruments. SRM 3127a is a single element solution prepared gravimetrically to contain a nominal 10 mg/mL of lanthanum with a nitric acid concentration (v/v) of 10 percent. The certified value (V) is based on replicate titrations against a reference solution of lanthanum metal of known purity. The value has been adjusted upward by 0.1% relative based on estimated transpiration losses of solvent through the container walls of 0.2% relative per year. The density of the solution at 22 °C is 1.069 g/mL.

Metal	Concentration (mg/mL)	Source Purity, %	Acid Conc. (v/v) Approximate
La	9.95 ± 0.02	La <sub>2</sub> O <sub>3</sub> (99.99) <sup>a</sup>	HNO <sub>3</sub> , 10%

<sup>a</sup>This high-purity material was analyzed by optical emission spectrometry and found to contain less than 100 µg/g total impurities.

The uncertainty in the certified value is calculated as

$$U = (2u_c + 0.001V) \text{ mg/mL}$$

where  $u_c$  is the "combined uncertainty" calculated according to the ISO Guide [1]. The value of  $u_c$  is intended to represent, at the level of one standard deviation, the combined effect of uncertainty components associated with volumetric, gravimetric, and titrimetric factors, as well as the purity of the lanthanum metal. The additional quantity, 0.001V, is an allowance for transpiration of the solution through the container walls, which is estimated to be ± 0.1% of the certified value during the one-year period of validity of the certification.

The combined uncertainty consists of Type B components due to uncertainties in material purity, balance reading, material handling, and dilution.

SRM 3127a was prepared by T.A. Butler of the NIST Analytical Chemistry Division. Titrimetric analyses were made by C.M. Beck II of the Analytical Chemistry Division.

The support aspects involved in the preparation, certification, and issuance of this SRM were coordinated through the Standard Reference Materials Program by J.S. Kane.

Gaithersburg, MD 20899  
January 24, 1995

Thomas E. Gills, Chief  
Standard Reference Materials Program

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## Procedures for Use

**Stability:** This certification is valid for one-year from the shipping date, provided the solution is kept tightly capped and stored under normal laboratory conditions. NIST will monitor the stability of representative solutions from the SRM lot, and if any changes occur that invalidate this certification, NIST will notify purchasers.

**Preparation of Working Standard Solutions:** Solutions should be brought to  $22 \pm 1$  °C before use and all glass or plastic surfaces coming into contact with the standard must have been previously cleaned. A working standard solution can be prepared from the SRM solution by serial dilution. Dilutions should be made with certified volumetric class A flasks and 5 or 10 mL class A pipets. All volumetric transfers of solutions should be performed using a proven analytical technique. Each dilution should be acidified with an appropriate high-purity acid and diluted to calibrated volume using high-purity water. The stability of the working standard solution will depend on the final acid concentration; therefore, care should be exercised to ensure that the final acid concentration of the dilution closely approximates that of the SRM. To achieve the highest accuracy, the analyst should prepare daily working solutions from 100  $\mu\text{g/mL}$  dilutions of the original SRM solution.

## REFERENCE

[1] "*Guide to the Expression of Uncertainty in Measurement*", ISBN 92-67-10188-9, 1st Ed. ISO, Geneva, Switzerland, 1993.