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Materials & Molecular Research Division

Presented at the International Conference on
Photochemistry and Photobiology, Alexandria, Egypt,
January 5-10, 1983

MOLECULAR BEAM STUDIES OF PRIMARY
PHOTOCHEMICAL PROCESSES

Yuan T. Lee

December 1982



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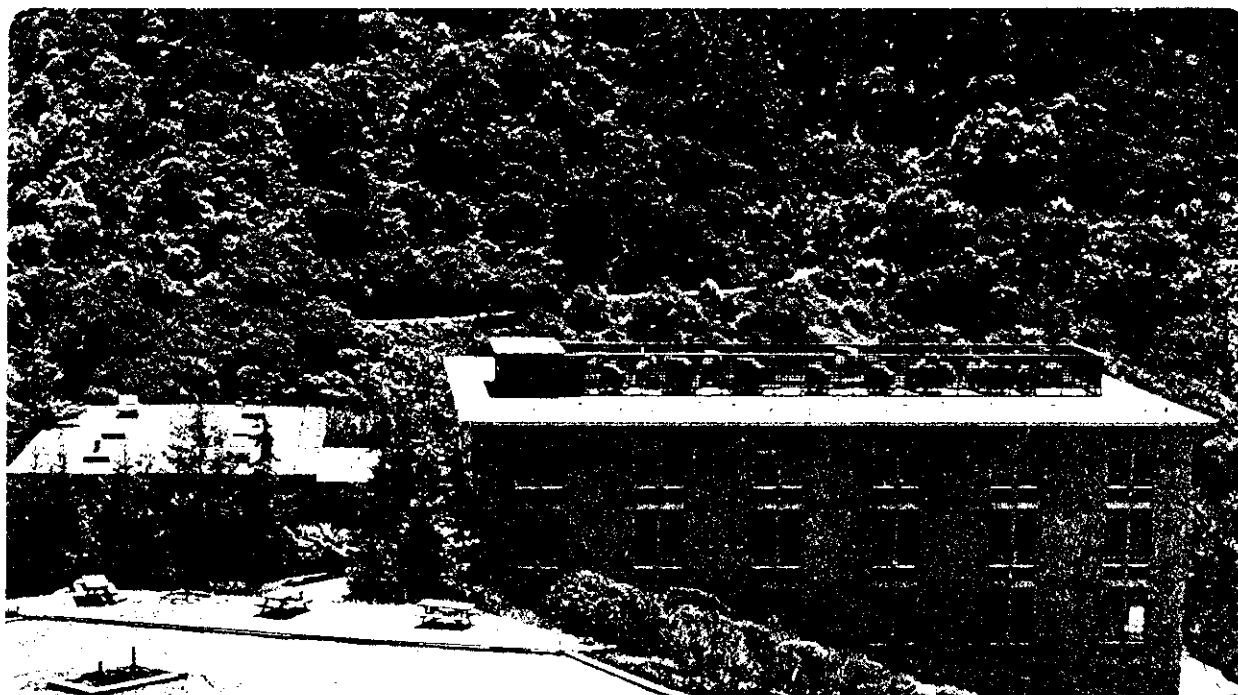
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ABSTRACT

Application of the method of molecular-beam photofragmentation translational spectroscopy to the investigation of primary photochemical processes of polyatomic molecules is described.

Examples will be given to illustrate how information concerning the energetics, dynamics, and mechanism of dissociation processes can be obtained from the precise measurements of angular and velocity distributions of products in an experiment in which a well-defined beam of molecules is crossed with a laser.

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INTRODUCTION

There are many questions concerning the primary photochemical processes of simple polyatomic molecules which still remain to be answered. Some of the questions most frequently asked concern the mechanism and the relative importance of various dissociation channels, the nature of excited states involved, the energetics of product formation and the detailed dynamics of dissociation processes.

In many cases, the lack of a clear understanding of the primary photochemical processes is due to the experimental complication caused by either the collisions of excited molecules prior to dissociation or the secondary reactions of primary radical products. Sometimes, the difficulty arises because of the lack of a suitable experimental method to identify the primary polyatomic radical products which are highly internally excited. These highly excited polyatomic radicals often fragment extensively during the electron bombardment ionization and do not yield the parent ions in mass spectrometric detection. The lack of either sufficient spectroscopic information or distinct features in absorption or emission spectra also make it difficult to detect these excited radicals spectroscopically.

The complication caused by collisional processes can be eliminated by carrying out the experiment at a very low pressure or under collision free conditions of a molecular beam. The positive identification of primary radical products, which is also problematic in the elucidation of the mechanism of some bimolecular

elementary chemical reactions, has also been shown to be possible, if high resolution molecular beam experiments are carried out using a mass spectrometric detector. From a careful comparison of angular and velocity distributions of products measured at various mass numbers which yield observable signals, and using kinematic relations together with the requirement of the conservation of energy and momentum, the mechanisms of a series of reactions involving oxygen atoms and unsaturated hydrocarbons have been successfully elucidated^{1,2} in spite of the severe fragmentation of products in the mass spectrometer.

The measurement of the velocity and angular distributions of photodissociation products in a molecular beam experiment has been referred to as the method of molecular beam photofragmentation translational spectroscopy. The usefulness of this method in the elucidation of dissociation dynamics and the characterization of repulsive electronic states has been fully demonstrated in pioneering research work carried out in the laboratories of Wilson³ and Bersohn.⁴ Recent advances in the development of high power lasers and improvements in the high intensity molecular beam source and the detection of scattered molecules make this method an extremely useful and important tool for the investigation of primary photochemical processes. Much of the initial confusion concerning the dynamics of infrared multiphoton dissociation processes have been clarified by this method.⁵

In this article some examples and results of our recent molecular beam photofragmentation studies will be described.

EXPERIMENTAL

The experimental arrangement used in our molecular beam photofragmentation translational spectroscopy has been described previously.^{6,7} Typically, a supersonic beam of molecules is formed in a nozzle-skimmer arrangement by expanding either a neat or a rare gas carried vapor through a 0.1 mm diameter nozzle. With a set of defining slits located on the walls of the two stage differential pumping chamber, the beam is defined to a 2° angular divergence. The molecular beam is crossed by the output of a pulsed high power laser in the main chamber which is maintained at 1×10^{-7} torr.

Photofragments are detected by a triply differentially pumped mass spectrometer which rotates in the plane of the two beams about the crossing volume. Product time-of-flight spectra are recorded at various angles using a multichannel scaler triggered by the laser. The ionizer of the mass spectrometer is located, depending on the apparatus used, either 20.8 cm or 34.1 cm from the intersection region of the laser and the molecular beam. Because of the defining slits on the detector chambers, only photofragments originating in the crossing volume of approximately 3 mm on a side are obtained.

EXAMPLES OF RECENT EXPERIMENTAL STUDIESA) The Dissociation Dynamics of Formaldehyde, H_2CO .⁸

One of the intriguing questions in the photodissociation of formaldehyde is the nature of the dissociation process and the energy distribution of product molecules ($\text{H}_2 + \text{CO}$) when the formaldehyde is excited to an energy level near the zero point energy of the S_1 state. It became clear in recent years that the internal conversion to S_0 precedes the dissociation of formaldehyde, although some questions concerning the "mysterious intermediate state"⁹ linger on.

Ab initio quantum mechanical calculations have shown that the transition state of the dissociation of H_2CO to $\text{H}_2 + \text{CO}$ in the S_0 potential energy surface is coplanar with both hydrogen atoms lying on one side of the CO axis.¹⁰ The barrier for the dissociation was estimated to be about 80 kcal/mole¹¹⁻¹³ which is almost identical to the excitation energy of the S_1 electronic state at the zero point. When the formaldehyde is excited near the zero point energy of S_1 , almost all the energy deposited in the molecules, 80 kcal/mole, will become potential energy at the transition state. This potential energy will be released in the production of $\text{H}_2 + \text{CO}$ after passing through the potential energy barrier.

The results of our recent molecular beam photofragmentation studies indicate that when formaldehyde is excited near the S_1 threshold, a large fraction of the available energy, 75 percent or 60 kcal/mole, appears as translational energy of the $\text{H}_2 + \text{CO}$ products, with a relatively narrow energy distribution.⁸ Since the CO

produced was shown by Houston and Moore⁹ to be mostly in the ground vibrational state, even if most of the remaining energies were to be in the vibrational excitation of H₂, the average vibrational quantum number of H₂ should not exceed $v' = 2$.⁸ The high translational energy in the dissociation products implies that most of the available energy is released repulsively between H₂ and CO after the H₂ bond is nearly fully developed and both C-H bonds are broken.

The experimental observation of high translational energy release with a relatively narrow distribution in products has important consequences in understanding the nature of the "mysterious intermediate state" in the dissociation of formaldehyde. From the geometry of the transition state and the observed product translational energy distribution, the rotational distribution of CO, which could be inverted, is expected to have an average quantum number of around $J' = 23$. This level of rotational excitation in CO contains only about 4 kcal/mole of energy, but the relatively high quantum number implies that the previous observation of the "delayed" production of CO, which was attributed to the existence of the "mysterious intermediate state" might be in error.⁸ Because, the production of CO in that experiment was monitored by the absorption of CO in $J' = 10$ using a CW CO laser, the observed delay in the production of CO ($J' = 10$) was likely to be the rotational relaxation of CO which was produced "promptly" in higher rotational states rather than due to the existence of the "mysterious intermediate state."

The high rotational state distribution of the CO product concluded from the geometry of the transition state and the results of the measured translational energy distributions has been verified by Ho and Smith¹⁴ in a recent experiment in which CO rotational state distributions were measured directly by the laser induced fluorescence method.

B) Dissociation Mechanism of Glyoxal, $\text{H}_2\text{C}_2\text{O}_2$.¹⁵

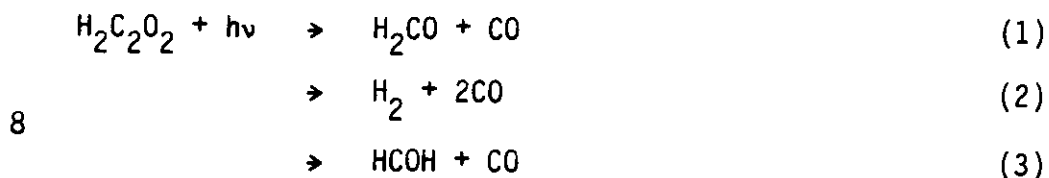
Glyoxal is another molecule for which the vibronic energy levels are well resolved spectroscopically and are also fully characterized. Recent investigations of Loge and Parmenter¹⁶ have shown that glyoxal dissociates in the absence of collision to produce H_2CO (formaldehyde), CO and H_2 as products when it is excited to the threshold of the S_1 electronic state (~ 60 kcal/mole). Although the formation of H_2CO and CO is understood to be the result of the 1,2 hydrogen atom migration after the electronically excited molecules are internally converted to the highly vibrationally excited ground state molecules, the production of H_2 is somewhat puzzling. Since the maximum possible internal excitation of H_2CO formed after the excitation of glyoxal to near the S_1 threshold is only around 60 kcal/mole, which is significantly less than the potential energy barrier (~ 80 kcal/mole)¹¹⁻¹³ of the dissociation of H_2CO , the secondary decomposition of vibrationally excited H_2CO is not likely. There is also no known radical dissociation channel at this excitation energy which could eventually yield H_2 as secondary products.

Extensive ab initio calculations intended for the understanding of the decomposition of glyoxal by Osamura et al.¹⁷ have shown that a concerted reaction involving four-center elimination of H₂ with a simultaneous rupture of the C-C bond to form two CO molecules from a cis configuration of glyoxal has an unexpectedly low potential energy barrier of 55 kcal/mole, which is comparable to that of 1,2 hydrogen atom migration forming H₂CO and CO. It suggests that the concerted reaction, $\text{H}_2\text{C}_2\text{O}_2 \rightarrow \text{H}_2 + 2\text{CO}$, is probably as important as the formation of H₂CO and CO in the photodissociation of glyoxal and H₂ molecules are in fact one of the primary products.

In order to verify the existence and the importance of the concerted reaction forming H₂ + 2CO, molecular beam photofragmentation experiments were carried out¹⁵ using a Nd-YAG laser pumped tunable dye laser system for excitation. The time-of-flight spectrum monitored at m/e = 30 shows the existence of two peaks which correspond to two dynamically distinct channels forming CH₂O + CO. From a consideration of the energies available for various possible channels and the measured translational energy distributions, it is concluded that the major peak with higher translational energy is due to the formation of H₂CO (formaldehyde) and CO, and the minor peak with lower translational energy is consistent with the formation of HCOH (hydroxymethylene) and CO. The time-of-flight spectra obtained at m/e = 29 (HCO⁺) contain the same two peaks observed in the spectra of m/e = 30. Apparently, HCO⁺ ions are daughter ions of m/e = 30 and not from the formation of the two HCO radicals from glyoxal. In fact, the excitation energy near the threshold of the

S_1 state of glyoxal is not enough to break the C-C bond to form two HCO radicals.

In the time-of-flight spectra of $m/e = 28$, in addition to the two features which correspond to the formation of $H_2CO + CO$ and $HCOH + CO$, there is an additional feature lying between those from $H_2CO + CO$ and $HCOH + CO$. This new channel is an indication of the existence of a third CO producing channel which has nothing to do with the formation of the CH_2O species. Although experimental difficulties prevented us from observing H_2 molecules directly, there is no doubt that this third CO producing channel is indeed due to a concerted reaction forming $H_2 + 2CO$. The branching ratios derived from the analysis of experimental results for the following dissociation channels



are 9:3.5:1 for reactions (1):(2):(3).

The formation of hydroxymethylene from glyoxal, which involves the 1,3 migration of the hydrogen atom from carbon atom to oxygen atom in the trans-configuration of glyoxal, was not suspected in previous experimental and theoretical studies, although there were some indications in earlier experimental studies of the possible existence of a radical channel which leads to the formation of some condensable products.¹⁶ Hydroxymethylene has not yet been observed

in the gas phase experimental investigations. But, the stability of this radical has been reported in an earlier ab initio calculations carried out by Goddard and Schaefer.¹⁰

C) Energy Splitting Between Singlet and Triplet Methylene, CH_2 .⁷

Although photoelectron spectra obtained from the photodetachment of CH_2^- gave the energy splitting between singlet methylene [$\text{CH}_2(\text{S})$] and triplet methylene [$\text{CH}_2(\text{T})$] to be 19 kcal/mole,¹⁸ most other experimental results and more rigorous ab initio quantum mechanical calculations gave the value around 9 to 10 kcal/mole.⁷ But, since the photodetachment experiment on CH_2^- carried out by Lineberger and coworkers has been the only microscopic experiment making a direct observation on the quantities related to the energy splitting, the significant difference of as much as 10 kcal/mole on this splitting became one of the most interesting controversies in chemical physics in recent years.

In order to provide a second microscopic experiment using a different approach for the resolution of this controversy, the molecular beam photofragmentation translational spectroscopy on ketene, CH_2CO , was carried out recently.⁷

In the photofragmentation of an isolated molecule, because of the conservation of energy, the formation of products in the ground vibrational rotational state of a given electronic state should accompany the release of all the available energy in translational motion of products. In other words, the difference between the

energy of exciting photons, $h\nu$, and the maximum translational energy released, E_m' , in the dissociation, should provide the bond dissociation energy. Thus, if two laser frequencies, ν_S and ν_T , were used for the dissociation of CH_2CO to produce $\text{CH}_2(\text{S}) + \text{CO}$ and $\text{CH}_2(\text{T}) + \text{CO}$, respectively, and the maximum translational energies released for both cases, $E_m'(\text{S})$ and $E_m'(\text{T})$, are measured accurately in a molecular beam experiment, the energy splitting between the singlet and triplet methylene can be derived simply from the difference between the quantities $h\nu_S - E_m'(\text{S})$ and $h\nu_T - E_m'(\text{T})$.

This method avoids any complication which might be caused by the existence of exit potential energy barriers, since it does not depend on measurements of the thresholds for production of singlet and triplet methylene. However, it requires formation of some of the products in the ground vibrational-rotational state. The selection of CH_2CO for this experiment has this consideration in mind. Because, the bond lengths of CO and CH in CH_2CO are almost identical to that of CO and CH_2 , and the in plane bending vibrational motion of CH_2 in CH_2CO is directly coupled to the dissociation coordinate, even if a large fraction of excess energy appears as product internal excitation, some of the products are likely to be in the ground vibrational-rotational state.

Exciting CH_2CO at 351 nm, the only energetically allowed channel is $\text{CH}_2(\text{T}) + \text{CO}$. At 308 nm, although the production of both singlet and triplet methylene are energetically allowed, almost all

the products are the formation of $\text{CH}_2(\text{S}) + \text{CO}$. Our results of the measurements of product translational energy distributions show that $E_m^{\text{I}}(\text{S})$ and $E_m^{\text{I}}(\text{T})$ are 6.7 and 3.8 kcal/mole respectively, which, in turn, yields the singlet-triplet splitting of 8.5 ± 0.8 kcal/mole.⁷ If the high value of 19 kcal/mole obtained in the photodetachment experiments is correct, the electron affinity of CH_2 obtained in the photodetachment experiment should also be increased by ~ 10 kcal/mole.

CONCLUDING REMARKS

The examples given above are only a small fraction of the photodissociation studies carried out by the molecular beam method. The works on the investigation of the dynamics of infrared multi-photon dissociation processes and the vibrational predissociation of hydrogen bonded molecules are some other examples.

It is also worth noting the important contributions made by the ab initio quantum mechanical calculations in the understanding of all three examples give above. The ability to accurately calculate the geometry and the energy levels of transition states and radical species will become more and more important in all areas of chemistry in the future.

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