Electron Interactions With Cl_2

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Low-energy electron interactions with the Cl_2 molecule are reviewed. Information is synthesized and assessed on the cross sections for total electron scattering, total rotational excitation, total elastic electron scattering, momentum transfer, total vibrational excitation, electronic excitation, total dissociation into neutrals, total ionization, total electron attachment, and ion-pair formation. Similar data on the density-reduced ionization, density-reduced electron attachment, density-reduced effective ionization, electron transport coefficients, and electron attachment rate constant are also synthesized and critically evaluated. Cross sections are suggested for total electron scattering, total elastic electron scattering, total ionization, dissociation into neutrals, electron attachment, and ion-pair formation. A cross section is derived for the total vibrational excitation cross section via low-lying negative ion resonances. Data are suggested for the coefficients for electron attachment, ionization, and effective ionization, and for the rate constant for electron attachment. While progress has been made regarding our knowledge on electron–chlorine interactions at low energies (<100 eV), there is still a need for: (i) improvement in the uncertainties of all suggested cross sections; (ii) measurement of the cross sections for momentum transfer, vibrational excitation, electronic excitation, and dissociative ionization; and (iii) accurate measurement of the electron transport coefficients in pure Cl_2 and in mixtures with rare gases. Also provided in this paper is pertinent information on the primary Cl_2 discharge byproducts Cl_2^+, Cl_2^-, Cl, Cl^-, and Cl^+. [S0047-2689(99)00401-8]

Key words: chlorine; Cl_2; cross section; electron attachment; electron collisions; electron scattering; electron transport; ionization.

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1. Introduction

Molecular chlorine (Cl_2) is a plasma processing gas (e.g., see Refs. 1–24). It is used in plasma etching of semiconductors where the Cl atoms produced in a gas discharge efficiently etch a silicon surface. The dominant primary electron interaction processes are taken to be single-step electron-impact ionization of Cl_2 and Cl, dissociation of Cl_2 into neutrals, and dissociative attachment to Cl_2.[1,10,13,17] The basic species involved in Cl_2 plasmas, then, are the three molecular species: Cl_2, Cl_2^-, and Cl_2^+, and the three atomic species: Cl, Cl^-, and Cl^+. Although recent work on the interactions of Cl_2 with slow electrons is largely motivated by plasma etching technology, considerable work on electron interactions with the Cl_2 molecule was done in the 1970s and the 1980s motivated by gas ultraviolet (UV) laser applications.[25–31] In this latter application the fundamental process of interest is dissociative electron attachment producing halogen atomic negative ions (Cl^-) which efficiently recombine with the rare-gas positive ions[29,30] to form the lasing species (e.g., ArF*, KrF*, and XeCl* excimers) in rare-gas–halide lasers.[27,28]

Molecular chlorine is also of atmospheric and environmental interest.[32] It is a potential atmospheric reservoir of chlorine atoms[33] which are released photolytically,

TABLE 1. Definition of symbols

Symbol	Definition	Common scale and units
$\sigma_{pa,t}(\lambda)$	Total photoabsorption cross section	10^{-20} cm^2; 10^{-24} m^2
$\sigma_{pi,t}(\lambda)$	Total photoionization cross section	10^{-20} cm^2; 10^{-24} m^2
$\sigma_{pi,partial}(\lambda)$	Partial photoionization cross section	10^{-20} cm^2, 10^{-24} m^2
$\sigma_{sc,t}(\varepsilon)$	Total electron scattering cross section	10^{-16} cm^2; 10^{-20} m^2
$\sigma_{rot,t}(\varepsilon)$	Total rotational electron scattering cross section	10^{-16} cm^2; 10^{-20} m^2
$\sigma_{rot,j\leftarrow 0}(\varepsilon)$	Cross section for rotational excitation of the j rotational state integrated over angle	10^{-16} cm^2; 10^{-20} m^2
$\sigma_{rot,sum}(\varepsilon)$	Rotationally summed electron scattering cross section	10^{-16} cm^2 sr^{-1}
$d^2\sigma_{rot,sum}/d\Omega\, d\varepsilon$	Rotationally summed differential electron scattering cross section	10^{-16} cm^2 sr^{-1} eV^{-1}
$\sigma_{e,t}(\varepsilon)$	Total elastic electron scattering cross section	10^{-16} cm^2; 10^{-20} m^2
$\sigma_{m}(\varepsilon)$	Momentum transfer cross section (elastic)	10^{-16} cm^2; 10^{-20} m^2
$\sigma_{vib,t}(\varepsilon)$	Total vibrational excitation cross section	10^{-16} cm^2; 10^{-20} m^2
$\sigma_{vib,indir}(\varepsilon)$	Total indirect vibrational excitation cross section	10^{-16} cm^2; 10^{-20} m^2
$\sigma_{elec}(\varepsilon)$	Electronic excitation cross section	10^{-16} cm^2; 10^{-20} m^2
$\sigma_{i,t}(\varepsilon)$	Total ionization cross section	10^{-16} cm^2; 10^{-20} m^2
$\sigma_{diss,t}(\varepsilon)$	Total dissociation cross section	10^{-16} cm^2; 10^{-20} m^2
$\sigma_{diss,neut,t}(\varepsilon)$	Total cross section for electron impact dissociation into neutrals	10^{-16} cm^2; 10^{-20} m^2
$\sigma_{a,t}(\varepsilon)$	Total electron attachment cross section	10^{-16} cm^2; 10^{-20} m^2
$\sigma_{da,t}(\varepsilon)$	Total dissociative electron attachment cross section	10^{-16} cm^2; 10^{-20} m^2
$\sigma_{ip}(\varepsilon)$	Cross section for ion-pair formation	10^{-18} cm^2; 10^{-22} m^2
$\sigma_{pdest,Cl_2^-}(\lambda)$	Photodestruction cross section for Cl_2^-	10^{-18} cm^2; 10^{-22} m^2
$\sigma_{pi,Cl}(\lambda)$	Photoionization cross section of Cl	10^{-18} cm^2; 10^{-22} m^2
$\sigma_{sc,t,Cl}(\varepsilon)$	Total electron scattering cross section for Cl	10^{-16} cm^2; 10^{-20} m^2
$\sigma_{e,t,Cl}(\varepsilon)$	Total elastic electron scattering cross section for Cl	10^{-16} cm^2; 10^{-20} m^2
$\sigma_{m,Cl}(\varepsilon)$	Momentum transfer cross section for Cl	10^{-16} cm^2; 10^{-20} m^2
$\sigma_{exc,t,Cl}(\varepsilon)$	Total electron-impact excitation cross section of Cl	10^{-16} cm^2; 10^{-20} m^2
$\sigma_{i,t,Cl}(\varepsilon)$	Total electron-impact ionization cross section for Cl	10^{-16} cm^2; 10^{-20} m^2
$\sigma_{i,Cl}(\varepsilon)$	Electron-impact single ionization cross section for Cl	10^{-16} cm^2; 10^{-20} m^2
$\sigma_{pd,Cl^-}(\lambda)$	Photodetachment cross section for Cl^-	10^{-18} cm^2; 10^{-22} m^2
$\sigma_{cd,Cl^-}(\mathcal{E})$	Collisional detachment cross section involving Cl^-	10^{-16} cm^2; 10^{-20} m^2
$\sigma_{ct,Cl^-}(\mathcal{E})$	Cross section for charge transfer in collisions involving Cl^-	10^{-16} cm^2; 10^{-20} m^2
$\sigma_{i,Cl^+}(\varepsilon)$	Cross section for single ionization of Cl^+	10^{-16} cm^2; 10^{-20} m^2
α/N	Density-reduced ionization coefficient	10^{-22} m^2
$(\alpha-\eta)/N$	Density-reduced effective ionization coefficient	10^{-22} m^2
η/N	Density-reduced electron attachment coefficient	10^{-22} m^2
$k_{a,t}$	Total electron attachment rate constant	10^{-10} cm^3 s^{-1}
$(k_{a,t})_{th}$	Thermal electron attachment rate constant	10^{-10} cm^3 s^{-1}
w	Electron drift velocity	10^6 cm s^{-1}
D_T/μ	Transverse electron diffusion coefficient to electron mobility ratio	V

$$Cl_2 + h\nu(\lambda \leqslant 500 \text{ nm}) \rightarrow 2Cl.$$

In this paper a number of collision cross sections, coefficients, and rate constants are used to quantify the various processes which result from collisions of low-energy (mostly below about 100 eV) electrons with the Cl_2 molecule. These are identified in Table 1 along with the corresponding symbols and units. The procedure for assessing and recommending data followed in this paper is the same as in the previous five papers in this series.[34–38] As will be discussed throughout this paper, few of the available data sufficiently meet the criteria to be "recommended." This demonstrates the need for additional data for this molecule. We have, however, "suggested" the best available data for each collision process.

Since the Cl_2 molecule is one of the simplest reactive gases used in plasma processing (often in mixtures with Ar), we consider it desirable from the point of view of this application to also provide relevant information on the most likely discharge byproducts, namely Cl_2^-, Cl_2^+, Cl, Cl^-, and Cl^+. In this way, one may have more complete information about the key species and processes. Therefore, while the emphasis in this paper is on low-energy electron interactions with the neutral Cl_2 molecule, pertinent information is also provided for the Cl_2^+ and Cl_2^- molecular ions, and for the atomic species Cl, Cl^-, and Cl^+.

An early attempt to critically evaluate low-energy electron-impact cross section data for Cl_2 was made by Morgan.[31] In addition, a number of investigators have used

Boltzmann codes and electron transport data to calculate cross sections and rate coefficients for some electron collision processes in Cl_2.[1,10,39–41] The value of these results is questionable however, partly because of the limited measurements on electron transport parameters upon which they are based, and because of the model dependent nature of the calculated cross sections. The results of two such calculations by Rogoff *et al.*[1] and Pinhão and Chouki[40] are compared with other data in later sections of this paper.

2. Electronic and Molecular Structure

2.1. Cl_2

The electronic structure of the outermost (valence) shell of the Cl_2 molecule in its ground electronic state[42–44] is: $..(\sigma_g 3p)^2, (\pi_u 3p)^4, (\pi_g 3p)^4, (\sigma_u 3p)^0$ and has $^1\Sigma_g^+$ symmetry. The first four excited electronic states of Cl_2 listed by Huber and Herzberg[45] are $A'\,^3\Pi_{2u}$, $A\,^3\Pi_{1u}$, $B\,^3\Pi_{0u}^+$, and $^1\Pi_u$. A number of theoretical and experimental studies have located many other excited electronic states (see below).

There are three types of sources of information of interest to the present study regarding the electronic structure of the chlorine molecule: calculations, photoabsorption and photoelectron studies, and electron energy-loss investigations. Although our interest is focused on the third type of information, basic information provided by the other two types of investigations is included in the paper as complementary.

The most useful theoretical work concerning the electronic states of chlorine are the *ab initio* calculations of Peyerimhoff and Buenker.[46] These workers calculated potential energy curves for the ground and excited electronic states of the chlorine molecule and for its positive and negative ions using the multireference single and double excitation with configuration interaction (MRD-CI) method. They considered all states which correlate with the lowest atomic limit $Cl(^2\Pi_u)+Cl(^2P_u)$] and many others which go into ionic $Cl^+ + Cl^-$ or Rydberg $Cl^* + Cl$ asymptotes. All singlet states which correlate with the ground atomic products were found to be repulsive. Among the triplet states of Cl_2 which dissociate into the ground state atoms only the $^3\Pi_u$ state is not repulsive. The potential energy curves calculated by Peyerimhoff and Buenker[46] are reproduced in Fig. 1. The potential-energy curves shown in the figure are for the electronic states of Cl_2 which dissociate into the lowest atomic limit $[Cl(^2P_u)+Cl(^2P_u)]$. In Fig. 2 are also shown the potential-energy curves for the lowest electronic states of Cl_2^+ with various asymptotic limits and a potential-energy curve for the ground state of Cl_2^- for the asymptotic limit $Cl(^2P_u)+Cl^-(^1S_g)$. As will be seen from subsequent discussions in this paper, the potential-energy curves for Cl_2, Cl_2^-, and Cl_2^+ in Fig. 1 are most helpful in understanding the low-energy electron interaction processes with the Cl_2 molecule. Peyerimhoff and Buenker calculated for the dissociation energy D_0, the vertical ionization energy, and the electron affinity of Cl_2, respectively, the values 2.455, 11.48, and 2.38 V. These values are in good agreement with experimental values discussed later in this paper. The estimated electronic-state energies (MRD-CI values) are listed in Table 2.

There have been many photoabsorption and photoionization studies,[33,45,47–66] as well as a number of photoelectron studies[67–73] of Cl_2. The data on the total photoabsorption cross section, $\sigma_{pa,t}(\lambda)$, of Cl_2 have been discussed and summarized by a number of groups (e.g., Gallagher *et al.*,[64] Maric *et al.*,[33] and Hubinger and Nee[66]). In Fig. 2 are plotted the measurements of $\sigma_{pa,t}(\lambda)$ of a number of investigators[33,49,51,57,59,62,63,65,66] in the wavelength range 15.5–550 nm. Between 250 and ~500 nm the agreement among the various measurements is good. A least squares fit to the data in this wavelength region is shown in Fig. 2 by the solid line. Data taken off this line are given in Table 3 as our recommended values for the $\sigma_{pa,t}(\lambda)$ of Cl_2 in this wavelength range. The extensive measurements of Samson and Angel[63] cover the wavelength range 15.5–103.8 nm and are recommended for this energy range (Table 3). The measurements of Samson and Angel of the total photoionization cross section $\sigma_{pi,t}(\lambda)$ show that $\sigma_{pi,t}(\lambda)$ is equal to the total photoabsorption cross section $\sigma_{pa,t}(\lambda)$ except in the wavelength range 82.5–77.0 nm where it is up to 10% lower, depending on the value of the wavelength. The decrease of the photoionization efficiency to values below 1.0 in this wavelength range has been attributed to photoabsorption processes leading to the production of neutrals.[63]

Measurements have also been made by Samson and Angel[63] of the partial photoionization cross section, $\sigma_{pi,partial}(\lambda)$, for the production of Cl_2^+ and Cl^+ by photon impact on Cl_2. These are shown in Fig. 3. In Fig. 3 are also plotted the results of Samson and Angel for the production of Cl_2^{++} by photon impact on Cl_2. The data in Fig. 3 show that for photon wavelengths down to ~80.0 nm, the cross section for the production of the Cl_2^+ ion is about equal to the total, that is, it far exceeds the cross section for the production of the Cl^+ ion (dissociative photoionization has a much lower probability than nondissociative photoionization). At progressively shorter wavelengths, dissociative photoionization becomes more probable. The cross section for double electron ejection is negligible down to about 40 nm. Carlson *et al.*[71] and Gallagher *et al.*[64] have published measurements of the production of Cl_2^+ in the ionic states $(2\pi_g^{-1})X\,^2\Pi_g$, $(2\pi_u^{-1})A\,^2\Pi_u$, and $(5\sigma_g^{-1})B\,^2\Sigma_g$ by photon impact on Cl_2.

Data on photoionization energetics are given in Table 4 where they are compared with data obtained using other methods. For further spectroscopic investigations of the electronic structure of the chlorine molecule see Lee and Walsh,[54] Iczkowski *et al.*,[55] Douglas *et al.*,[56] Frost *et al.*,[67] Bondybey and Fletcher,[81] Huber and Herzberg,[45] Douglas,[61] Moeller *et al.*,[82] Burkholder and Bair,[62] McLoughlin *et al.*,[83] Lonkhuyzen and de Lange,[72] and Frost *et al.*[77]

There have been three major electron-impact studies of the electronic structure of Cl_2: the threshold-electron excitation study of Jureta *et al.*[43] which covered the excitation energy range up to 11.5 eV, the electron energy-loss study of Spence *et al.*[44] which covered the energy-loss range 5.5–14.5 eV, and the electron energy-loss study of Stubbs *et al.*[84]

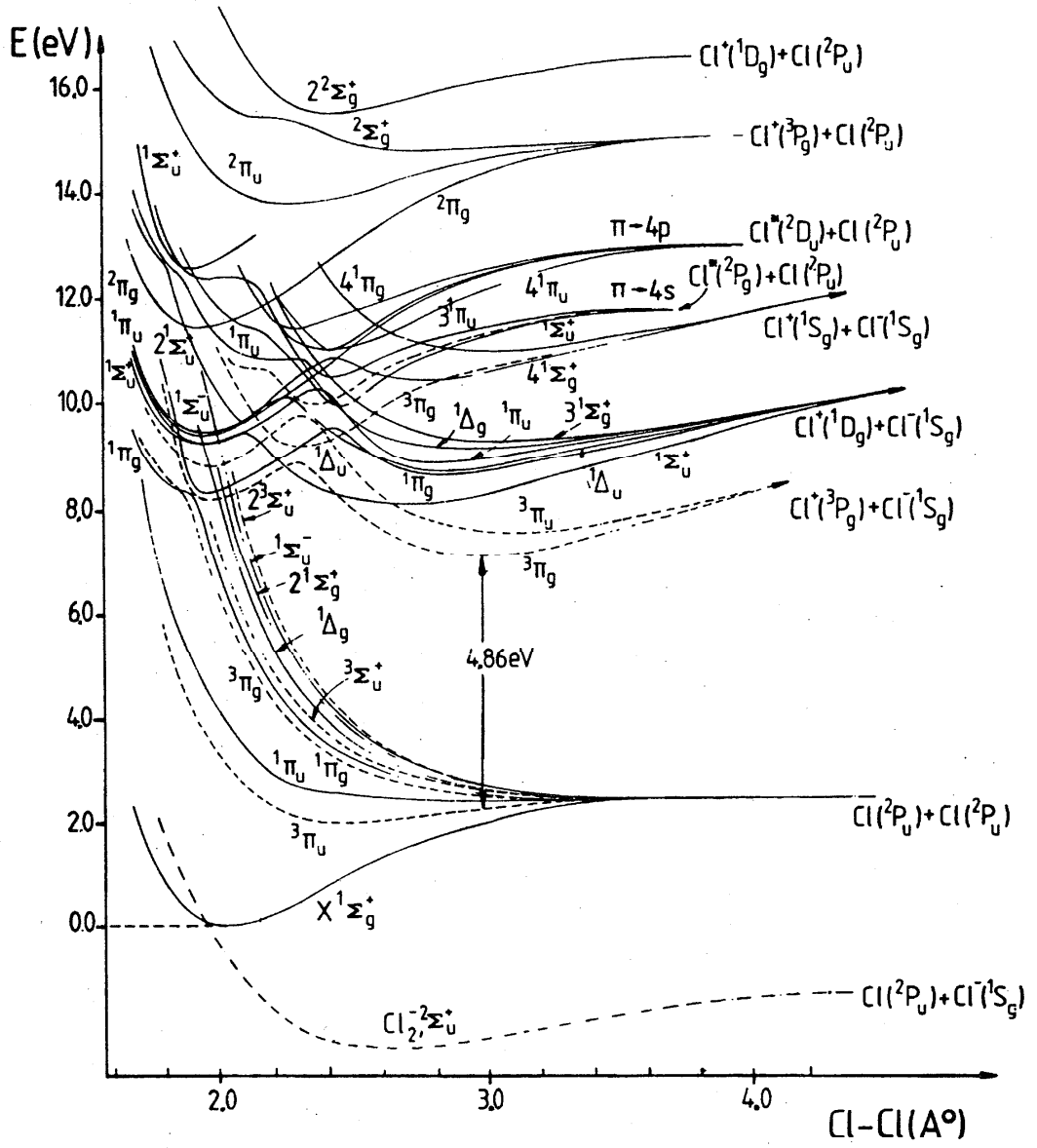


FIG. 1. Composite potential-energy diagrams for most of the electronic states of the Cl_2 molecule as calculated by Peyerimhoff and Buenker (Ref. 46).

which covered the electron energy-loss range up to 14.252 eV. Threshold-electron excitation methods are best suited for locating optically forbidden states, while electron energy-loss spectra using sufficiently energetic electrons give spectra similar to photoabsorption. Figure 4(a) shows the threshold-electron excitation spectrum of Cl_2 obtained by Jureta *et al.*[43] in the region of Rydberg excitation between about 7.5 and 11.5 eV, and Fig. 4(b) shows an electron energy-loss (in the range 5.5–11.5 eV) spectrum of Cl_2 obtained by Spence *et al.*[44] at a scattering angle of 3° using 200 eV incident energy electrons. The threshold-electron excitation technique of Jureta *et al.* had an energy resolution [full-width-at-half-maximum (FWHM)] of about 35 meV and the energy-loss experiment of Spence *et al.* had an energy resolution of 50–60 meV. As expected, the spectra they obtained at small scattering angles [Fig. 4(b)] differ from the threshold-electron excitation spectra and correspond closely to the photoabsorption spectra. The most prominent features of the energy-loss spectra arise from excitation of optically allowed Rydberg states. Larger-angle scattering data showed additional structures due to excitation of optically forbidden states. The spectra also showed the presence of hot bands. Such observation of hot bands in electron scattering spectra is unusual, but because the ground-state vibrational spacing of Cl_2 is small (0.0694 eV) and the Franck–Condon overlapings are particularly favorable, such structures become relatively strong for some electronic transitions. Stubbs *et al.*[84] had a better electron beam energy resolution (FWHM $\cong$18 meV) than the other two studies. This allowed them to obtain highly resolved electron energy-loss spectra up to

TABLE 2. Vertical electronic energies (MRD-CI values) from the ground state of Cl_2 to various excited states as calculated by Peyerimhoff and Buenker (Ref. 46)

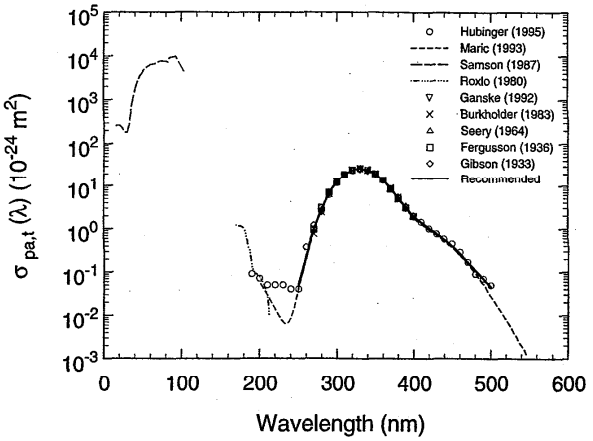
State/Excitation	Vertical energy (eV)
$X\,^1\Sigma_g^+$	0.00
$1\,^3\Pi_u\ \pi_g \rightarrow \sigma_u$	3.24
$1\,^1\Pi_u\ \pi_g \rightarrow \sigma_u$	4.04
$1\,^3\Pi_g\ \pi_u \rightarrow \sigma_u$	6.23
$1\,^1\Pi_g\ \pi_u \rightarrow \sigma_u$	6.86
$1\,^3\Sigma_u^+\ \sigma_g \rightarrow \sigma_u$	6.80
$^1\Delta_g\ \pi_g^2 \rightarrow \sigma_u^2$	8.25
$2\,^3\Pi_g\ \pi_g \rightarrow 4s$	8.34
$2\,^1\Pi_g\ \pi_g \rightarrow 4s$	8.38
$^1\Delta_g\ \pi_g^2 \rightarrow \sigma_u^2$	8.25
$2\,^1\Sigma_g^+\ \pi_g^2 \rightarrow \sigma_u^2$	8.35
$2\,^3\Pi_u\ \pi_g \rightarrow 4p\sigma$	8.80
$2\,^1\Pi_u\ \pi_g \rightarrow 4p\sigma$	9.22
$1\,^1\Sigma_u^+\ \pi_g \rightarrow 4p\pi$	9.32
$1\,^1\Sigma_u^-\ \pi_g \rightarrow 4p\pi$	9.58
$1\,^1\Delta_u\ \pi_g \rightarrow 4p\pi$	9.62
$2\,^1\Sigma_u^-\ \pi_u\pi_g \rightarrow \sigma_u^2$	9.67
$2\,^3\Sigma_u^+\ \pi_u\pi_g \rightarrow \sigma_u^2$	9.75
$^1\Delta_g\ \pi_g \rightarrow 4d\pi$	9.92
$^1\Pi_g\ \pi_g \rightarrow 4d\sigma$	10.01
$^1\Pi_g\ \pi_g \rightarrow 4d\delta$	10.10
$2\,^1\Sigma_u^+\ \sigma_g \rightarrow \sigma_u;\ \pi_g \rightarrow 4p\pi$	10.34
$3\,^3\Pi_u\ \pi_u \rightarrow 4s$	11.33
$3\,^1\Pi_u\ \pi_u \rightarrow 4s$	11.51
$^1\Pi_u\ \pi_u \rightarrow 4d\delta$	12.74

TABLE 3. Recommended total photoabsorption cross section, $\sigma_{pa,t}(\lambda)$, for Cl_2. Data of Samson and Angel (Ref. 63) in the wavelength range of 16–103.8 nm, and data taken off the solid line in Fig. 2 between 250 and 500 nm

Wavelength (nm)	$\sigma_{pa,t}(\lambda)$ (10^{-24} m^2)	Wavelength (nm)	$\sigma_{pa,t}(\lambda)$ (10^{-24} m^2)
16.0	255	320	23.7
20.0	260	330	25.6
30.0	186	340	23.6
40.0	1480	350	19.1
50.0	4220	360	13.3
60.0	6300	370	8.41
70.0	7180	380	5.10
80.0	7480	390	3.06
85.0	9040	400	1.92
92.4	9803	410	1.39
103.8	4384	420	1.02
		430	0.77
250	0.050	440	0.56
260	0.23	450	0.39
270	0.91	460	0.26
280	2.66	470	0.17
290	6.44	480	0.11
300	11.9	490	0.071
310	18.4	500	0.046

.4.252 eV, using electrons with incident energy between 10 ınd 120 eV. These incident energies are lower than that (200 :V) used by Spence *et al.*

The strongest structures in the energy-loss spectrum of :hlorine lie between 9 and 10 eV under all scattering condiions. They primarily consist of two vibrational series com-›rising transitions that are allowed by electric-dipole selecion rules and have been previously reported in both photoabsorption and electron impact studies. In Table 5 are listed the energy positions of the energy-loss peaks observed in the electron impact studies of Spence *et al.*[44] and Jureta *et al.*[43] For comparison, photoabsorption data from Moeller *et al.*[82] are also shown along with possible identification of the states responsible for the observed energy losses. A comparison of the values of the energies of the various states as determined from the energy-loss spectra and from the threshold-electron excitation spectra shows excellent agreement (Table 5). The higher-energy resolution in the study of Stubbs *et al.*[84] allowed detection of more transitions than in the other two studies. These are listed in Table 6.



IG. 2. Total photoabsorption cross section as a function of photon wavength, $\sigma_{pa,t}(\lambda)$, for Cl_2: (– –) Ref. 63; (–···–) Ref. 59; (○) Ref. 66; (- - -) ef. 33; (▽) Ref. 65; (×) Ref. 62; (△) Ref. 57; (□) Ref. 51; (◇) Ref. 49; –) recommended.

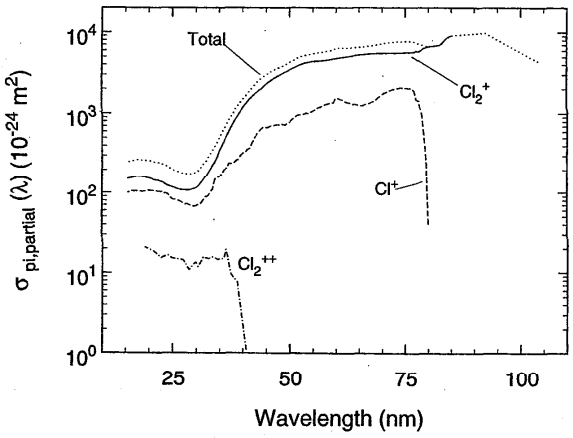


FIG. 3. Partial photoionization cross sections as a function of photon wavelength, $\sigma_{pi,partial}(\lambda)$, for the production of the positive ions Cl_2^+, Cl^+, and Cl_2^{++} from Cl_2 (data of Samson and Angel, Ref. 63): (—) Cl_2^+; (– –) Cl^+; (–·–) Cl_2^{++}; (···) total.

TABLE 4. Dissociation energy, vibrational energy, equilibrium internuclear separation, spin-orbit splitting, electron affinity, energy position of negative ion states, ionization threshold energy, dissociative ionization threshold energy, energy threshold for double ionization, and energy threshold for ion-pair formation of Cl_2

Physical quantity	Value/Method/Reference
Dissociation energy Cl_2 ($X\,^1\Sigma_g^+$) (eV)	$D_0=2.4793$, (45) $D_0=2.4794$, (56) $D_e=2.5139$, (56)
Vibrational energy (eV)	0.0694, (45), (56)
Equilibrium internuclear separation (Å)	1.9879, (45) 1.9878, (56)
Spin-orbit splitting (eV)	0.080 ± 0.002 [for the $(1\pi_g)^{-1}$ state], (68)
Electron affinity (eV)	2.45,[a] (74)
Energy position of negative ion states (eV)	See Table 15 in Sect. 6
Ionization threshold energy (eV)	
Cl_2^+ ($X\,^2\Pi_{g,3/2}$)	*Adiabatic* 11.50 (photoelectron), (45) 11.48 ± 0.01 (photoionization), (53) 11.49[b] (photoelectron) ($^2\Pi_g$), (69) 11.48 (photoelectron), (72) 11.480 ± 0.005 eV, (75) 11.50 (photoelectron) ($^2\Pi_g$), (67) 11.5 (photoelectron), (70) 11.51 ± 0.01 (photoelectron), (68) *Vertical* 11.48 ± 0.01 (electron impact), (76) 11.559 (photoelectron), (72) 11.59 ± 0.01 (photoelectron), (68)
$Cl_2^+(^2\Pi_{g,1/2})$	11.56, (75) 11.63 (vertical, electron momentum spectroscopy, $2\pi_g$), (77) $\sim$11.6 (electron impact), (76)
$Cl_2^+(^2\Pi_{u,3/2})$	$\sim$11.8 (electron impact), (76) 11.80 (electron impact), (78) 11.80 ± 0.1 (electron impact), (79)
$Cl_2^+(^2\Pi_{u,1/2})$	$\sim$11.9 (electron impact), (76)
$Cl_2^+(^2\Pi_u)$	*Adiabatic* 13.96 ± 0.02 (photoelectron), (68) 14.0 (photoelectron), (70) 14.0[c] (photoelectron), (69) 14.04 (photoelectron, $^2\Pi_{u,3/2}$), (72) 14.11 (photoelectron), (67) *Vertical* 14.33 (photoelectron), (70) 14.39 (photoelectron) ($^2\Pi_{u,3/2}$), (72) 14.40 ± 0.02 (photoelectron), (68) 14.43[d] (photoelectron), (69) 14.41 (electron momentum spectroscopy, $2\pi_u$), (77)
$Cl_2^+(^2\Sigma_g^+)$	*Adiabatic* 15.72 ± 0.02 (photoelectron), (68) 15.70[c] (photoelectron), (72) 15.8 (photoelectron), (70) 15.8[c] (photoelectron), (69) *Vertical* 14.09 ± 0.03 (electron impact), (76) 15.94 (photoelectron), (67) 16.082 (photoelectron), (72) 16.08 ± 0.02 (photoelectron), (68) 16 (photoelectron), (70) 16.10[d] (photoelectron), (69) 16.18 (electron momentum spectroscopy, $5\sigma_g$), (77)

TABLE 4. Dissociation energy, vibrational energy, equilibrium internuclear separation, spin-orbit splitting, electron affinity, energy position of negative ion states, ionization threshold energy, dissociative ionization threshold energy, energy threshold for double ionization, and energy threshold for ion-pair formation of Cl_2—Continued

Physical quantity	Value/Method/Reference
$Cl_2^+(^2\Sigma_u^+)$	*Vertical* 20.61±0.06 (electron impact), (76) 21.8, 24.0 (electron momentum spectroscopy, $4\sigma_u$), (77) 27.3 (electron momentum spectroscopy, $4\sigma_g$), (77)
Dissociative ionization ($Cl_2+e\rightarrow Cl^++Cl+2e$) threshold energy (eV)	15.45 (adiabatic), (70) 15.7±0.3 (electron impact), (79)
Energy threshold for double ionization (eV)	30.5 (photoionization), (63) 31.13 [Cl_2^{++} ($X\,^3\Sigma_g^-$, $v=0$), threshold photoelectron spectroscopy], (73)
Energy threshold for ion-pair ($Cl_2+e\rightarrow Cl^++Cl^-+e$) formation (eV)	11.9±0.2, (79), (80)

[a]Thirteen values are listed by Christodoulides *et al.* (Ref. 74). If we exclude the lowest three as too low and the highest one as too high, the average of the other nine values is 2.45 eV which is within the combined quoted uncertainty of the averaged values.
[b]0–0 band.
[c]Onset.
[d]Band maximum.

2.2. Cl_2^-

The Cl_2^- negative ion consisting of Cl^- ($^1S_0$) and Cl ($^2P_{3/2,1/2}$) has four electronic states. These states can be arranged[80,85–90] in order of increasing energy as: $^2\Sigma_u^+$, $^2\Pi_g$, $^2\Pi_u$, and $^2\Sigma_g^+$. In Fig. 5(a) are shown the potential-energy curves for these states as calculated by Gilbert and Wahl[85] in the molecular-orbital self-consistent-field (SCF) approximation. In Fig. 5(b) similar curves are shown for the negative ion states $^2\Sigma_u^+$, $^2\Pi_{g,1/2}$, and $^2\Sigma_g^+$ as they have been determined by Lee *et al.*[89] using their photodissociation cross section measurements for Cl_2^- and adjusted potential-energy curves for Cl_2^- based on those calculated by Gilbert and Wahl.[85] The numerical values shown in Fig. 5(b) for the various quantities are those used by Lee *et al.*, and the designations a_1 and a_2 refer, respectively, to the Cl ($^2P_{3/2}$) and Cl ($^2P_{1/2}$) asymptotic limits. Data for a number of physical parameters of the Cl_2^- ion are given in Table 7.

2.3. Cl_2^+

Photoelectron studies have shown[70] that the known states of the Cl_2^+ ion correspond to the ejection of one electron from one of the occupied orbitals of the outer orbital structure of the Cl_2 molecule [$(\sigma_g)^2$, $(\pi_u)^4$, $(\pi_g)^4$]. The $^2\Sigma_g^+$ state of Cl_2^+ lies above the first dissociation limit (Ref. 70; Table 4). Optical emission from the excited $A\,^2\Pi_u$ state of Cl_2^+ to the ground state $X\,^2\Pi_g$ of Cl_2^+ is known,[70] but emission from the $^2\Sigma_g^+$ state to the $A\,^2\Pi_u$ state, although allowed by the selection rules, has not been observed, possibly because the $^2\Sigma_g^+$ state is entirely predissociated.[70] The photodissociation spectra[83] of Cl_2^+ obtained in the range 1.80–2.55 eV showed vibrational structure indicating that the

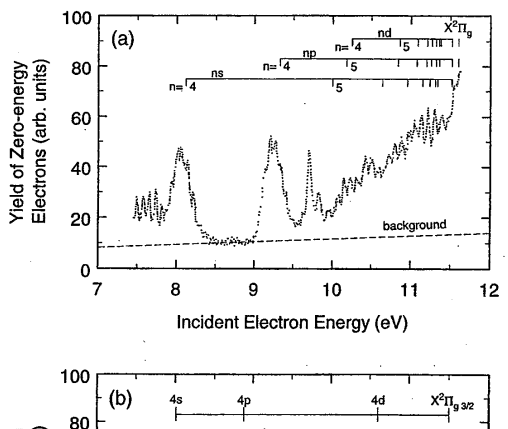


FIG. 4. (a) Threshold-electron excitation spectrum of Cl_2. Here the electron energy is varied and the excitation spectrum reflects the relative probability of electrons having energy from about 7.5 eV to about 11.5 eV to lose all their energy in a collision with a Cl_2 molecule. The experiment detects only the "zero-energy electrons" (data of Jureta *et al.*, Ref. 43). (b) Electron energy-loss spectrum from Ref. 44 of Cl_2 between 7.5 and 11.5 eV using 200 eV incident electrons and a scattering angle of 3°. The scale above the spectrum shows the expected locations of Rydberg states due to excitations of a $\pi_g 3p$ electron to $4s$, $4p$, and $4d$ orbitals.

TABLE 5. Comparison of the energies of the $4s\sigma_g\,^3\Pi_g$, $4s\sigma_g\,^1\Pi_g$, $2\,^3\Pi(1u)$, $2\,^1\Pi_u$, $2\,^1\Sigma_u^+$, and $^1\Pi_g(?)$ states of Cl_2. (The electronic configuration and term symbol are as given by Spence *et al.* in Ref. 44.)

Name/ Assignment	Vibrational level	Energy (eV) Energy-loss experiment (Ref. 44)	Energy (eV) Threshold-electron excitation experiment (Ref. 43)	Energy (eV) Photoabsorption experiments[a] (Ref. 82)
$4s\sigma g\,^3\Pi_g$	hot band		(7.83)[b]	
	$v=0$		7.91	
	$v=1$		7.99	
	$v=2$		(8.07)	
	$v=3$		(8.15)	
$4s\sigma g\,^1\Pi_g$	hot band		(7.87)	
	$v=0$	7.939	(7.95)	
	$v=1$	8.019	8.03	
	$v=2$	8.101	8.11	
	$v=3$	8.186	8.19	
	$v=4$	8.270	8.27	
	$v=5$	8.354	(8.35)	
$2\,^3\Pi(1u)$	$v=0$		9.130	9.116
	$v=1$		9.190	9.193
$2\,^1\Pi_u$	$v=0$	9.225	9.250	9.230
	$v=1$	9.305	9.320	9.307
	$v=2$	9.381	9.395	9.384
	$v=3$	9.455	9.465	9.459
	$v=4$		9.530	9.534
$2\,^1\Sigma_u^+$	hot band		(9.620)	
	$v=0$	9.682	9.695[c]	9.688
	$v=1$	9.815	9.815[c]	9.807
	$v=2$	9.938	9.930[c]	9.928
	$v=3$	10.046		10.028
	$v=4$	10.141		
$^1\Pi g(?)$		9.900		
		9.966		
		10.025		

[a]Other photoabsorption data can be found in Refs. 54, 55, and 61.
[b]Numbers in parentheses represent unresolved components.
[c]May be due to the presence of a nearby triplet state.

dissociation of these ions involves a predissociation mechanism. Data on low-lying ionic states of Cl_2^+ derived from optical emission and photoelectron spectroscopy investigations are listed in Table 8.

3. Electron Scattering for Cl_2

3.1. Total Electron Scattering Cross Section, $\sigma_{sc,t}(\varepsilon)$

Up until very recently, the only data on the total electron scattering cross section, $\sigma_{sc,t}(\varepsilon)$, for Cl_2 were the 1937 measurements of Fisk[91] which are very large (Fig. 6). The absence of reliable experimental data on $\sigma_{sc,t}(\varepsilon)$, coupled with the lack of calculations of this quantity, led to two very recent measurements[92,93] of $\sigma_{sc,t}(\varepsilon)$ for Cl_2. The measurements of Gulley *et al.*[92] covered the energy range 0.02–9.5 eV and those of Cooper *et al.*[93] the energy range 0.3–23 eV. They are plotted in Fig. 6 and are seen to be very much smaller than the old measurements of Fisk. The uncertainties are estimated to be ±20% in the measurements of Cooper *et al.* and ±8% in the measurements of Gulley *et al.* While the shape of $\sigma_{sc,t}(\varepsilon)$ as determined by the two recent measurements is similar, the magnitude of $\sigma_{sc,t}(\varepsilon)$ as measured by Cooper *et al.* is systematically lower than that measured by Gulley *et al.* at all but the lowest energies (below ~0.7 eV). The magnitude of the data of Gulley *et al.* is consistent with the total rotational excitation cross sections (see Sec. 3.2). Cooper *et al.*[93] pointed out that the lower values of their $\sigma_{sc,t}(\varepsilon)$ measurements may in part result from the fact that electrons scattered into small angles ($\leq 2°$) with little energy loss are detected as "unscattered" in their apparatus, and since the measurements of Gote and Ehrhardt[94] on rotational scattering from Cl_2 showed (see Sec. 3.5.1) that forward scattering is appreciable, this may be a significant cause of error in determining the value of $\sigma_{sc,t}(\varepsilon)$.

In the energy range covered by the two recent experimental studies, the $\sigma_{sc,t}(\varepsilon)$ for Cl_2 has two distinct features: It shows a minimum around 0.4 eV and structure that can be attributed to resonance-enhanced electron scattering. In connection with the latter, the peaks at low energies in the Gulley *et al.*[92] data and the bump (or unresolved peak) in the Cooper *et al.*[93] data at 2.5 eV correspond to the energy po-

LE 6. Transitions observed by Stubbs *et al.* in Ref. 84 in a high-lution energy-loss experiment below the second ionization $Cl_2^+(\Pi_u)$ et

ne/Assignment	Vibrational level (v)	Excitation energy (eV)
$(5s\sigma_g)^1\Pi_g$	0	9.803[a]
	1	9.886
	2	9.962
	3	10.037
	4	10.121
F	–	9.162[b]
	–	9.602[b]
	–	9.743[b]
	–	10.693[b]
G$(8p\sigma_u)^1\Pi_u$	1	11.272
	2	11.356
	3	11.435
	4	11.513
	5	11.593
	6	11.670
a	0	10.937[c]
	1	11.029
	2	11.105
	3	11.193
	4	11.275
	5	11.358
	6	11.437
	7	11.500
	8	11.581
H	–	10.162[b]
	–	10.230
)	–	10.196[c]
	–	10.278
[	–	10.764[b]
	–	11.157[b]
	–	11.251[b]
:	–	10.711[c]
I$(4s\sigma_g)^1\Sigma_u^+$	0	12.565[b]
	1	12.605
	2	12.655
	3	12.699
	4	12.742
	5	12.785
	6	12.834
	7	12.880
	8	12.926
	9	12.971
K	0	12.953[b]
	1	13.005
	2	13.063
	3	13.112
	4	13.170
	5	13.224
$A(5s\sigma_g)^1\Sigma_u^+$	0	13.290[b]
	1	13.335
	2	13.380
	3	13.429
	4	13.466
	5	13.516
	6	13.559
	7	13.599

TABLE 6. Transitions observed by Stubbs *et al.* in Ref. 84 in a high-resolution energy-loss experiment below the second ionization $Cl_2^+(\Pi_u)$ onset—Continued

Name/Assignment	Vibrational level (v)	Excitation energy (eV)
M$(6s\sigma_g)^1\Sigma_u^+$	0	13.631[b]
	1	13.674
	2	13.715
	3	13.757
	4	13.803
	5	13.844
	6	13.884
	7	13.934
	8	13.977
	9	14.023
	10	14.066
d	0	11.835[c]
	1	11.915
	2	11.984
	3	12.055
	4	12.127
e	0	12.113[c]
	1	12.167
	2	12.215
	3	12.264
	4	12.319
f	–	12.353[c]
	–	12.402
	–	12.459
	–	12.488

[a]Symmetry forbidden.
[b]Allowed.
[c]Spin forbidden.

sitions of the negative ion states identified in electron attachment studies near 0 and 2.5 eV (see Sec. 6.1 and Refs. 87, 95, and 96). Similarly, the strong peak near 7 eV (Fig. 6) corresponds to the negative ion state (see Sec. 6.1 and Refs. 87, 95, 96) at 5.5 eV overlapping with the lowest electron-excited Feshbach resonance of Cl_2 at 7.5 eV which has been identified by Spence[97] in an electron transmission experiment. The spacing of the peaks and inflections in the data of Gulley *et al.* at 0.09, 0.14, and 0.2 eV may be associated with indirect (resonance enhanced) vibrational excitation of Cl_2 via the near 0 eV negative ion state of Cl_2^- (see discussion in subsequent sections).

In view of the fact that the data of Gulley *et al.*[92] exhibit lower uncertainty, superior electron energy resolution, a more extensive energy range, and consistency with rotational excitation cross section data,[94] we performed a least squares fit to the measurements of Gulley *et al.* which we extended to 23 eV using the shape of the Cooper *et al.* cross section between 9.5 and 23 eV. The solid line in Fig. 6 is a plot of this least-squares fit, and represents our recommended $\sigma_{sc,t}(\varepsilon)$ for Cl_2. Values taken from this curve are listed in Table 9.

3.2. Total Rotational Electron Scattering Cross Section, $\sigma_{rot,t}(\varepsilon)$

Recently Gote and Ehrhardt[94] measured the absolute differential cross sections for electron-impact rotational excita-

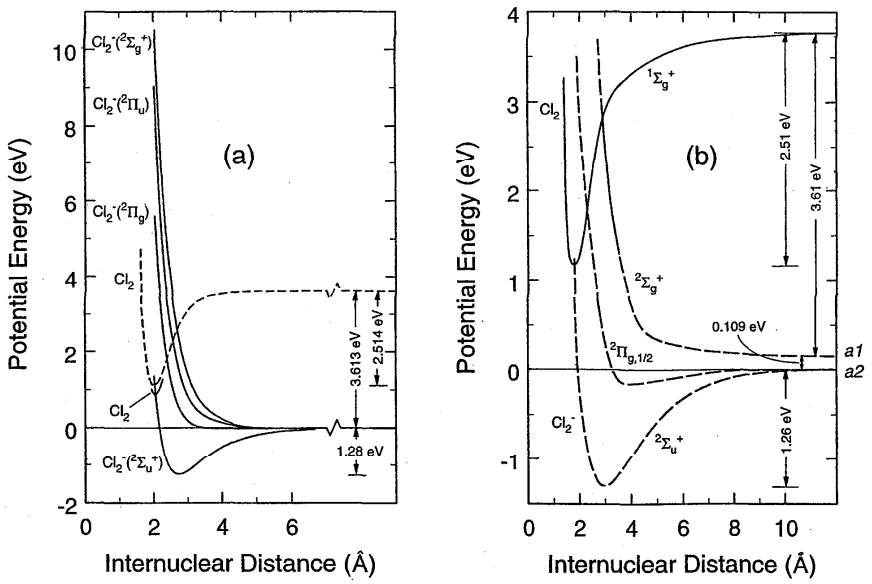


FIG. 5. (a) Potential-energy curves for the lowest four negative ion states ($^2\Sigma_u^+$, $^2\Pi_g$, $^2\Pi_u$, $^2\Sigma_g^+$) of Cl_2^- as calculated by Gilbert and Wahl in Ref. 85 using the molecular orbital self-consistent-field approximation. (The broken and solid curves for Cl_2 are two different determinations by Gilbert and Wahl.) (b) Potential-energy curves for the states $^2\Sigma_u^+$, $^2\Pi_{g,1/2}$, and $^2\Sigma_g^+$ of Cl_2^- determined by Lee *et al.* in Ref. 89 using their photodissociation cross section measurements for Cl_2^- and the potential-energy curves of Gilbert and Wahl in (a). The designations $a1$ and $a2$ represent, respectively, the asymptotic limits Cl ($^2P_{3/2}$)+Cl^- and Cl ($^2P_{1/2}$)+Cl^-.

tion of Cl_2. The measurements of Gote and Ehrhardt are listed in Table 10. These data allowed determination of the total cross section for rotational scattering (rotational elastic plus rotational inelastic) as a function of electron-impact energy, $\sigma_{rot,t}(\varepsilon)$. In Fig. 7 is plotted (open circles) the cross section $\sigma_{rot,t}(\varepsilon)$ as determined (summed over all j values and all scattering angles) by Kutz and Meyer[98] from the data of Gote and Ehrhardt (Table 10). Also plotted in Fig. 7 are the full-potential calculation results of Kutz and Meyer (solid circles) which extend over a much larger energy range. There is good agreement between theory and experiment in the overlapping energy range. It is interesting to observe that both the experimental measurements (Table 10) and the calculation[98,99] show a "rotational rainbow," i.e., a maxi-

TABLE 7. *Some physical parameters for* Cl_2^-

Quantity	Value	Method/Reference
Dissociation energy, D_e	1.28 eV	Calculation, (85)
	1.24 eV	Calculation, (86)
Dissociation energy, D_0	1.26 eV	(45)
Equilibrium internuclear distance, R_e	2.65 Å	Calculation, (85)
	2.71 Å	Calculation, (86)
Fundamental vibrational frequency	0.0322 eV	Calculation, (85)
	0.0320 eV	Calculation, (86)
Transition energies[a]	3.46 eV ($^2\Sigma_g^+ \rightarrow {}^2\Sigma_u^+$)	Calculation, (86)
	2.89 eV ($^2\Pi_u \rightarrow {}^2\Sigma_u^+$)	
	1.78 eV ($^2\Pi_g \rightarrow {}^2\Sigma_u^+$)	
Ionization energy of Cl_2^-[b]	2.39 eV	(45)
Electron affinity of Cl_2[b]	2.45 eV[c]	(74)
Negative ion states	~0.0 eV ($^2\Sigma_u^+$)	From Table 15, Sec. 6
	2.5 eV ($^2\Pi_g$)	
	5.5 eV ($^2\Pi_u$)	
	7.5 eV ($X\,^2\Pi_g$)($4s\sigma$)^2[$^2\Pi_{1/2,3/2}$]	

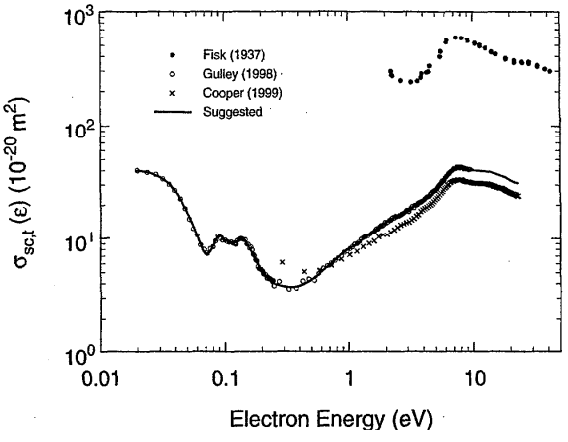
[a]At the ground state equilibrium bond length.
[b]These two quantities should be the same and have both adiabatic and vertical values. The vertical values normally exceed the adiabatic.
[c]Thirteen values are listed by Christodoulides *et al.* in Ref. 74. If we exclude the lowest three values listed in this reference as too low and the highest one as too high, the average of the remaining nine values is 2.45 eV. This value is within the combined quoted uncertainty of the values used in the averaging.

TABLE 8. Some parameters for Cl_2^+

Parameter	Value	Reference
Equilibrium separation (Å)	1.890 ($^2\Pi_{g,3/2}$)	72
Dissociation energy (eV)	3.99 ($^2\Pi_g$)	67
	3.966 ($^2\Pi_{g,3/2}$)	72
	3.876 ($^2\Pi_{g,1/2}$)	72
	1.38 ($^2\Pi_u$)	67
	1.41 ($^2\Pi_{u,3/2}$)	72
	1.32 ($^2\Pi_{u,1/2}$)	72
D_0	3.95	45
Energy of fundamental vibration (eV)	0.08004 ($^2\Pi_{g,3/2}$)	45
	~0.080 ($^2\Pi_{g,3/2}$)	72
	0.07994 ($^2\Pi_{g,1/2}$)	45
	~0.0459 ($^2\Pi_{u,3/2}$)	72
	~0.0347 ($^2\Sigma_g^+$)	72
Energies to various ionic states (adiabatic/vertical) (eV)	See Table 4	

num in the rotational excitation cross section at a relatively ıigh Δj. The experimental and calculated values of $\sigma_{rot,t}(\varepsilon)$ re compared in Fig. 7 with the suggested value of $\sigma_{sc,t}(\varepsilon)$ solid line in Fig. 6). From the figure it can be seen that $\sigma_{rot,t}(\varepsilon)$ exceeds the total scattering cross section near 2 eV. This is physically impossible, but the discrepancy is well within the combined uncertainties of the two measurements. t is interesting to note the deep minimum shown by the alculated $\sigma_{rot,t}(\varepsilon)$ that is also present in the measured $\sigma_{sc,t}(\varepsilon)$. Below this minimum the calculated values for $\sigma_{rot,t}(\varepsilon)$ exceed the measured $\sigma_{sc,t}(\varepsilon)$.

In Fig. 8 are shown the various contributions to $\sigma_{rot,t}(\varepsilon)$, ıat is, the integrated (over angle) excitation cross sections, $\sigma_{rot,j\leftarrow 0}(\varepsilon)$, for $j=0$, 2, 4, and 6. Clearly the rotationally lastic electron scattering channel ($j=0$) dominates over all nergies, especially below the minimum.



G. 6. Total electron scattering cross section, $\sigma_{sc,t}(\varepsilon)$, for Cl_2: (●) Ref. 91;)) Ref. 92; (×) Ref. 93; (—) recommended values.

TABLE 9. Recommended total electron scattering cross section, $\sigma_{sc,t}(\varepsilon)$, for Cl_2

Electron energy (eV)	$\sigma_{sc,t}(\varepsilon)$ (10^{-20} m^2)	Electron energy (eV)	$\sigma_{sc,t}(\varepsilon)$ (10^{-20} m^2)
0.02	40.0	0.80	6.55
0.03	35.2	0.90	7.36
0.04	26.8	1.00	7.97
0.05	17.0	1.20	9.06
0.06	10.7	1.50	11.1
0.07	7.36	2.00	13.9
0.08	8.50	2.50	16.0
0.09	10.6	3.00	17.9
0.10	9.68	3.50	19.9
0.11	9.26	4.00	21.9
0.12	9.06	4.50	24.2
0.13	9.76	5.00	26.8
0.14	9.89	6.00	34.5
0.15	8.90	7.00	41.2
0.17	7.19	8.00	42.8
0.20	5.09	9.00	41.0
0.22	4.44	10.0	40.3
0.25	4.00	12.0	39.7
0.30	3.75	14.0	38.6
0.35	3.70	16.0	36.7
0.40	3.80	18.0	35.1
0.50	4.32	20.0	33.0
0.60	5.00	22.0	31.5
0.70	5.83	23.0	31.0

3.3. Total Elastic Electron Scattering Cross Section, $\sigma_{e,t}(\varepsilon)$

There are no measurements of the total elastic electron scattering cross section, $\sigma_{e,t}(\varepsilon)$, for Cl_2. There have been, however, two calculations of this cross section, the old phase-shift calculation of Fisk,[91] and the more recent close-coupling calculation of Rescigno.[100] These results are shown in Fig. 9. The Fisk result is clearly unacceptable. We have also plotted in Fig. 9 the total rotational scattering cross section $\sigma_{rot,t}(\varepsilon)$ as calculated by Kutz and Meyer.[98] Similarly, we have plotted the $\sigma_{rot,t}(\varepsilon)$ determined by Kutz and Meyer from the measurements of Gote and Ehrhardt.[94] From an experimental perspective, $\sigma_{rot,t}(\varepsilon)$ may be considered equivalent to $\sigma_{e,t}(\varepsilon)$ since $\sigma_{rot,t}(\varepsilon)$ contains a large elastic component and the energy loss of rotational excitations is small ($<10^{-4}$ eV).[98] Clearly, the $\sigma_{rot,t}(\varepsilon)$ based on the experimental data of Gote and Ehrhardt and the $\sigma_{e,t}(\varepsilon)$ calculated by Rescigno are similar in shape and comparable in magnitude over a large energy range. The solid line in Fig. 9 represents a fit to these two data sets, and values obtained from this fit are given in Table 11 as our suggested set of data for $\sigma_{e,t}(\varepsilon)$ for molecular chlorine.

3.4. Momentum Transfer Cross Section, $\sigma_m(\varepsilon)$

There are no measurements of the momentum transfer cross section, $\sigma_m(\varepsilon)$, for Cl_2. The results of two Boltzmann-code analyses[1,40] are questionable, in part because they were hindered by the lack of accurate electron transport coefficient measurements. The two Boltzmann analyses used the early

TABLE 10. Differential rotational excitation cross sections for electron scattering from Cl_2 from Ref. 94. The rotationally summed cross sections, $\sigma_{rot,sum}(\varepsilon)$, (in units of 10^{-16} cm^2 sr^{-1}) are also listed. The partial cross sections are listed as the percentage of their relative contribution to $\sigma_{rot,sum}(\varepsilon)$

Scattering angle	10°	20°	30°	40°	50°	60°	70°	80°	90°	100°	110°	120°	130°	140°	150°	160°
2 eV																
$j_t=0$	53.1	63.7	97.0	98.2	100	100	100	90.7	86.5	77.9	49.6	31.5	17.3	18.6	34.6	43.7
$j_t=2$	40.7	31.0	<1	<1	<1	<1	<1	4.1	12.1	17.6	50.4	67.6	82.7	70.8	65.3	54.9
$j_t=4$	6.1	3.4	<1	<1	<1	<1	<1	1.6	<1	4.2	<1	<1	<1	6.3	<1	<1
$j_t=6$	<1	<1	<1	<1	<1	<1	<1	3.5	<1	<1	<1	<1	<1	4.2	<1	<1
$j_t=8$	<1	<1	<1	1.8	<1	<1	<1	<1	1.4	<1	<1	<1	<1	<1	<1	<1
$\sigma_{rot,sum}(\varepsilon)$	1.59	1.13	0.86	1.02	1.32	1.53	1.62	1.58	1.45	1.29	1.03	0.72	0.54	0.51	0.60	0.74
5 eV																
$j_t=0$	100	90.1	73.1	67.2	79.6	74.1	68.4	33.1	19.6	4.2	32.3	28.4	42.4	55.5	56.5	52.5
$j_t=2$	<1	6.1	24.6	32.8	14.3	14.4	17.0	55.0	68.5	74.7	64.2	70.3	40.9	36.3	24.4	31.6
$j_t=4$	<1	3.4	2.3	<1	3.7	9.0	7.9	11.4	7.7	21.1	<1	<1	13.2	6.7	16.8	15.8
$j_t=6$	<1	<1	<1	<1	<1	2.5	4.8	<1	4.1	<1	1.1	<1	2.1	1.5	2.4	<1
$j_t=8$	<1	<1	<1	<1	<1	<1	2.0	<1	<1	<1	1.1	<1	1.4	<1	<1	<1
$\sigma_{rot,sum}(\varepsilon)$	5.82	4.58	3.35	2.98	2.54	2.02	1.72	1.51	1.38	1.18	1.12	1.20	1.15	1.16	1.23	1.29
10 eV																
$j_t=0$	97.1	100	97.3	76.3	72.2	43.7	31.8	29.2	27.3	18.3	<1	18.0	16.4	11.2	11.7	1.3
$j_t=2$	2.8	<1	<1	10.6	27.3	45.0	45.3	55.6	40.4	57.7	73.3	62.5	44.7	65.5	47.9	55.5
$j_t=4$	<1	<1	1.5	6.3	<1	<1	12.1	5.3	26.4	18.5	21.6	7.6	28.6	20.4	39.2	41.0
$j_t=6$	<1	<1	<1	6.1	<1	3.8	9.1	8.4	3.4	1.4	<1	7.2	6.7	<1	<1	<1
$j_t=8$	<1	<1	<1	<1	<1	5.8	<1	1.2	<1	1.4	<1	3.2	<1	1.1	<1	<1
$\sigma_{rot,sum}(\varepsilon)$	21.28	14.39	7.49	4.62	1.74	1.41	1.08	0.96	0.84	0.86	0.87	0.98	1.08	1.57	2.06	3.12
20 eV																
$j_t=0$	100	94.9	79.2	83.9	58.1	39.0	80.4	62.6	28.8	14.5	24.0	15.0	18.9	11.0	12.6	<1
$j_t=2$	<1	<1	20.8	5.5	35.2	35.9	8.6	17.1	45.1	43.0	35.2	19.6	24.9	31.8	12.8	17.1
$j_t=4$	<1	1.4	<1	7.8	3.7	16.4	7.0	13.4	26.1	33.5	38.5	55.7	39.5	57.2	51.0	62.3
$j_t=6$	<1	2.3	<1	2.7	3.0	5.8	4.0	2.7	<1	5.2	2.2	9.7	13.6	<1	17.5	16.3
$j_t=8$	<1	1.2	<1	<1	<1	3.1	<1	4.2	<1	3.8	<1	<1	3.1	<1	6.1	4.3
$\sigma_{rot,sum}(\varepsilon)$	31.61	19.32	9.41	3.97	1.83	1.32	0.96	0.75	0.80	0.89	0.89	0.83	0.65	0.55	0.57	0.89
30 eV																
$j_t=0$	99.8	87.3	68.1	30.6	25.2	<1	<1	<1	<1	<1	<1	<1	2.1	<1	6.6	2.8
$j_t=2$	<1	12.2	24.1	69.4	51.5	83.0	68.7	54.3	32.1	26.1	17.1	12.4	16.2	11.1	15.5	4.5
$j_t=4$	<1	<1	6.1	<1	18.2	10.6	25.8	41.1	52.5	69.1	58.1	58.2	48.5	51.5	51.7	70.9
$j_t=6$	<1	<1	1.7	<1	<1	6.4	3.2	<1	13.9	3.9	23.9	28.3	29.6	34.4	22.6	21.8
$j_t=8$	<1	<1	<1	<1	<1	<1	<1	<1	1.4	<1	<1	<1	<1	<1	<1	<1
$\sigma_{rot,sum}(\varepsilon)$	31.84	15.90	5.86	2.26	1.18	0.69	0.41	0.43	0.62	0.71	0.59	0.39	0.22	0.11	0.11	0.20
50 eV																
$j_t=0$	95.2	84.6	50.8	3.6	16.7	13.5	<1	8.7	<1	3.6	2.5	5.8	<1	10.5	8.6	4.9
$j_t=2$	3.3	11.9	31.4	81.4	69.7	69.3	46.9	21.8	15.8	10.0	15.3	13.3	17.3	12.0	10.1	8.8
$j_t=4$	1.4	3.5	14.4	12.3	13.6	<1	35.5	54.6	52.8	61.7	53.7	43.7	48.6	31.4	22.9	28.4
$j_t=6$	<1	<1	3.3	2.8	<1	13.2	4.7	11.0	28.1	21.3	27.4	36.2	32.2	29.8	28.2	35.3
$j_t=8$	<1	<1	<1	<1	<1	<1	1.3	<1	3.1	3.4	<1	<1	<1	15.4	21.5	22.5
$\sigma_{rot,sum}(\varepsilon)$	28.72	8.61	2.41	1.01	0.51	0.20	0.16	0.30	0.44	0.49	0.44	0.29	0.18	0.15	0.33	0.69
100 eV																
$j_t=0$	99.6	76.3	17.5	15.7	20.5	26.5	<1	8.1	6.3	7.0	<1	13.9	6.0	<1	<1	<1
$j_t=2$	<1	23.5	72.1	55.4	50.4	31.8	19.1	23.8	15.3	14.3	7.1	18.4	6.7	5.3	<1	<1
$j_t=4$	<1	<1	7.0	23.8	29.1	22.7	50.6	39.1	45.2	34.9	30.6	15.9	7.6	17.1	7.0	3.0
$j_t=6$	<1	<1	3.4	5.1	<1	10.3	24.4	24.0	30.2	29.9	37.7	12.3	23.5	31.3	34.5	30.6
$j_t=8$	<1	<1	<1	<1	<1	6.5	5.7	3.2	1.9	11.5	20.1	19.8	30.2	33.0	42.0	41.9
$j_t=10$	<1	<1	<1	<1	<1	1.7	<1	<1	<1	<1	4.2	18.1	20.6	12.9	16.2	24.0
$j_t=12$	<1	<1	<1	<1	<1	<1	<1	<1	<1	<1	<1	1.7	4.4	<1	<1	<1
$\sigma_{rot,sum}(\varepsilon)$	15.70	3.50	0.91	0.42	0.19	0.13	0.15	0.20	0.19	0.15	0.08	0.04	0.07	0.15	0.59	0.91

TABLE 10. Differential rotational excitation cross sections for electron scattering from Cl_2 from Ref. 94. The rotationally summed cross sections, $\sigma_{rot,sum}(\varepsilon)$, (in units of 10^{-16} cm^2 sr^{-1}) are also listed. The partial cross sections are listed as the percentage of their relative contribution to $\sigma_{rot,sum}(\varepsilon)$ —Continued

Scattering angle	10°	20°	30°	40°	50°	60°	70°	80°	90°	100°	110°	120°	130°	140°	150°	160°
200 eV																
$j_t=0$	9.4	14.9	24.3	13.7	<1	3.6	<1	6.3	<1	<1	<1	<1	<1	<1	<1	<1
$j_t=2$	5.6	85.1	52.0	30.9	19.0	12.2	2.0	6.7	<1	13.0	<1	<1	<1	<1	<1	<1
$j_t=4$	<1	<1	21.6	42.5	39.3	25.5	23.9	28.5	33.1	6.2	<1	1.1	<1	<1	<1	<1
$j_t=6$	<1	<1	1.1	8.9	34.2	39.7	42.5	34.3	47.6	11.3	32.5	23.7	5.3	<1	2.6	<1
$j_t=8$	<1	<1	<1	1.7	6.9	17.6	28.6	14.2	13.3	23.9	33.6	31.1	21.0	20.1	17.5	15.7
$j_t=10$	<1	<1	<1	<1	<1	<1	1.4	7.2	3.6	24.8	20.4	30.9	29.7	28.2	25.1	25.0
$j_t=12$	<1	<1	<1	<1	<1	<1	<1	2.4	<1	13.0	13.2	13.2	30.7	40.1	35.3	42.8
$j_t=14$	<1	<1	<1	<1	<1	<1	<1	<1	<1	7.6	<1	<1	7.0	9.5	15.4	15.3
$\sigma_{rot,sum}(\varepsilon)$	12.92	2.52	0.83	0.36	0.23	0.19	0.13	0.09	0.06	0.04	0.03	0.05	0.11	0.24	0.34	0.58

data of Bailey and Healey[101] for the electron drift velocity and the characteristic energy for a 20%Cl_2:80%He mixture by volume, and the ionization and attachment coefficients of Božin and Goodyear.[102] In Fig. 10 the Boltzmann-calculation results are compared with the close-coupling calculation result of Rescigno.[100] These cross sections differ substantially, especially at low energies, stressing the need for a direct measurement of $\sigma_m(\varepsilon)$. They also indicate the need for measurements of electron transport coefficients that would allow a more reliable Boltzmann-code analysis. The need for such measurements is made more apparent because the cross sections of Rogoff *et al.*[1] have been used commonly in various discharge models. Of the available values for $\sigma_m(\varepsilon)$, the *ab initio* calculations of Rescigno[100] are preferred because they are not model dependent and because of the agreement between Rescigno's calculations and measured values of $\sigma_{e,t}(\varepsilon)$ and $\sigma_{diss,neut,t}(\varepsilon)$ (see Secs. 3.3 and 5, respectively).

3.5. Inelastic Electron Scattering Cross Section, $\sigma_{inel}(\varepsilon)$

3.5.1. Rotational Excitation Cross Section, $\sigma_{rot}(\varepsilon)$

Rotational excitation of Cl_2 by electron impact can be either direct or indirect via the formation of short-lived negative ion states. The experimental measurements of Gote and Ehrhardt[94] on the absolute differential cross sections for rotational excitation of Cl_2 by electron impact at energies between 2 and 200 eV and in the angular range 10°–160°, clearly show (Table 10) that rotational excitation of the Cl_2 molecule in its vibrational and electronic ground states by slow electrons is an efficient electron scattering process. Cross sections exceeding 10^{-16} cm^2 have been measured. As discussed earlier in this section, Gote and Ehrhardt reported rotationally summed cross sections and partial rotational excitation cross sections (i.e., cross sections for excitation to various rotational levels) as the percentage of their relative

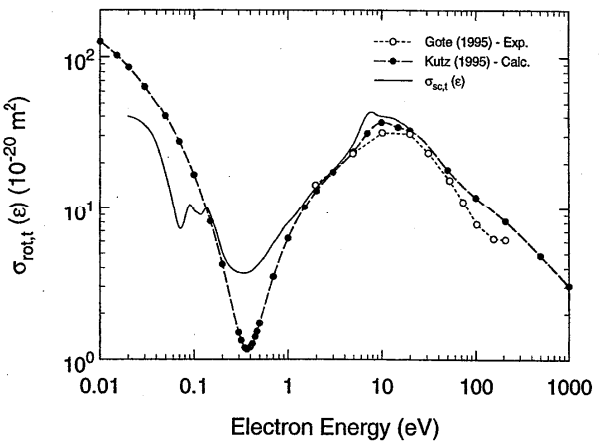


FIG. 7. Total cross section for rotational scattering, $\sigma_{rot,t}(\varepsilon)$, for Cl_2 as reported by Kutz and Meyer (Ref. 98): (○) values calculated by Kutz and Meyer from the measurements of Gote and Ehrhardt (Ref. 94); (●) *ab initio* calculations (Ref. 98). For comparison the suggested $\sigma_{sc,t}(\varepsilon)$ from Table 9 (solid line in Fig. 6) is also plotted.

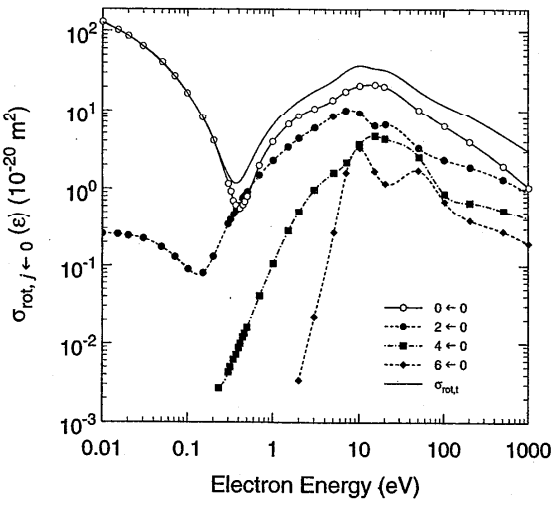


FIG. 8. Integrated (over angle) excitation cross sections, $\sigma_{rot,j\leftarrow 0}(\varepsilon)$, for Cl_2 from Ref. 98 for the rotational excitation channels (○) 0←0; (●) 2←0; (■) 4←0; (♦) 6←0 of Cl_2. Also shown for comparison is $\sigma_{rot,t}(\varepsilon)$.

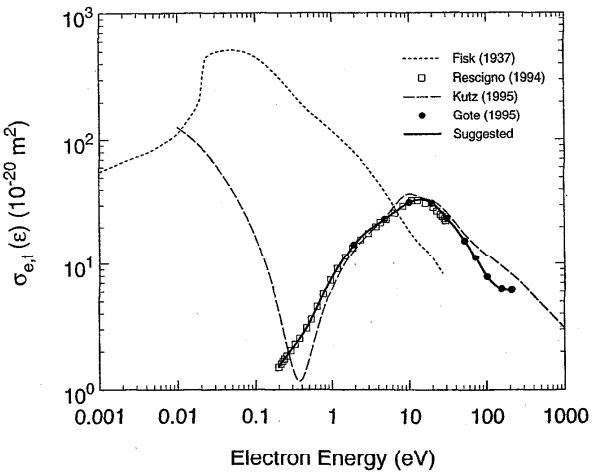


FIG. 9. Total elastic electron scattering cross section, $\sigma_{e,t}(\varepsilon)$, for Cl_2: (- - -) calculated total elastic electron scattering cross section, $\sigma_{e,t}(\varepsilon)$ from Ref. 91; (●) measured $\sigma_{rot,t}(\varepsilon)$ [data of Ref. 98 based on measurements by Ref. 94]; (– –) calculated total rotational electron scattering cross section, $\sigma_{rot,t}(\varepsilon)$ from Ref. 98; (□) calculated total elastic electron scattering cross section, $\sigma_{e,t}(\varepsilon)$ from Ref. 100; (—) suggested values.

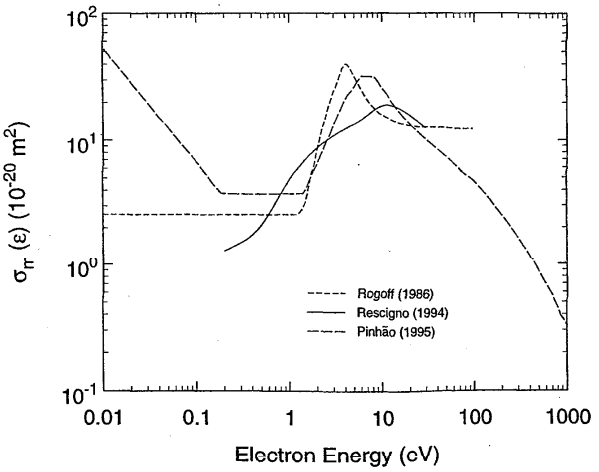


FIG. 10. Calculated momentum transfer cross sections, $\sigma_m(\varepsilon)$, for Cl_2: (- - -) Ref. 1; (—) Ref. 100; (– –) Ref. 40.

contribution to the rotationally summed cross sections. The cross sections in the forward direction belong mostly to rotationally elastic scattering. Above a scattering angle of about 30°, the scattering is dominated by rotationally inelastic processes. Kutz and Meyer's[98] close-coupling calculation of the rotational excitation of Cl_2 by electron impact over the energy range of 0.01–1000 eV, neglecting vibrational, resonant, and electronic excitation, shows two different excitation mechanisms, the importance of each depends on electron energy. At low electron energies only a few rotational quanta are exchanged and the differential cross section decreases exponentially with Δj. At high electron energies the excitation spectrum shows a rotational rainbow, i.e., the differential cross section has a maximum at a relatively high Δj. The location of the maximum depends on electron en-

TABLE 11. Suggested total elastic electron scattering cross section, $\sigma_{e,t}(\varepsilon)$, for Cl_2

Electron energy (eV)	$\sigma_{e,t}(\varepsilon)$ (10^{-20} m^2)	Electron energy (eV)	$\sigma_{e,t}(\varepsilon)$ (10^{-20} m^2)
0.20	1.50	7.00	27.1
0.22	1.64	8.00	28.8
0.25	1.82	9.00	30.2
0.30	2.11	10.0	31.3
0.35	2.38	12.0	32.7
0.40	2.66	14.0	33.1
0.50	3.30	16.0	32.9
0.60	4.10	18.0	32.1
0.70	4.98	20.0	30.9
0.80	5.99	22.0	29.5
0.90	6.89	23.0	28.8
1.00	7.77	25.0	27.3
1.20	9.34	30.0	24.0
1.50	11.4	40.0	19.4
2.00	14.6	50.0	16.1
2.50	16.9	60.0	13.6
3.00	18.6	70.0	11.6
3.50	19.9	80.0	10.1
4.00	21.1	90.0	8.87
4.50	22.1	100.0	7.99
5.00	23.2	150.0	6.31
6.00	25.2	200.0	6.16

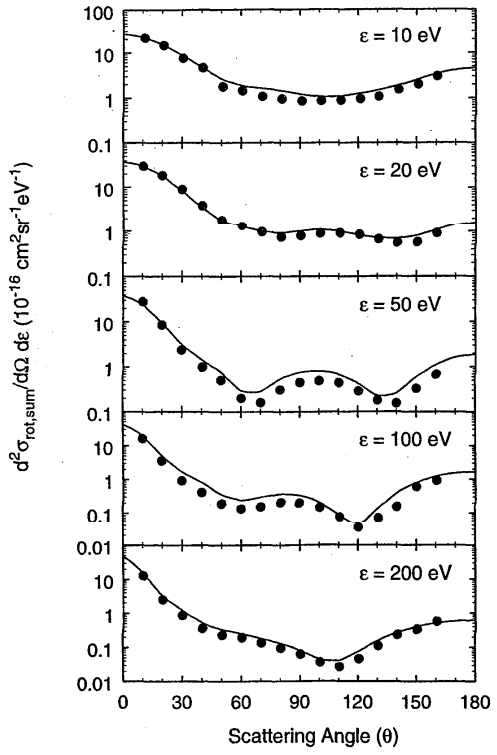


FIG. 11. Comparison of experimental and calculated rotationally summed differential electron scattering cross sections $d^2\sigma_{rot,sum}/d\Omega\, d\varepsilon$, for Cl_2 at incident electron energies of 10, 20, 50, 100, and 200 eV from Gote and Ehrhardt (Ref. 94): (●) experimental data from Ref. 94; (—) close-coupling calculation results from Ref. 98.

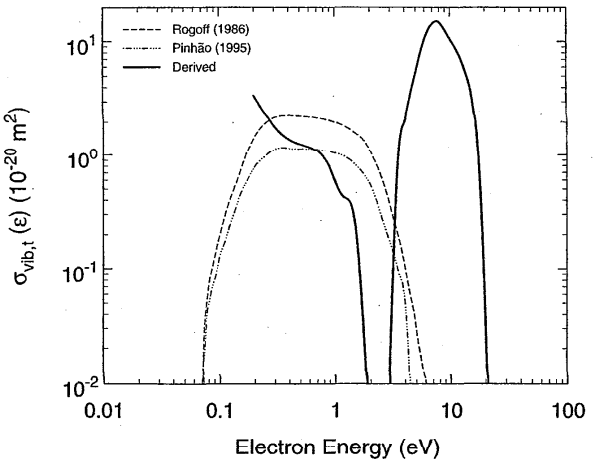


FIG. 12. Total vibrational excitation cross section, $\sigma_{vib,t}(\varepsilon)$, for Cl_2. Results of Boltzmann-code analyses: (- - -) Ref. 1; (–···–) Ref. 40. Estimate of $\sigma_{vib,t}(\varepsilon)$ derived from present analysis described in text (—).

ergy and scattering angle. For the observation of a rotational rainbow not only high electron energies, but also high scattering angles are needed. The scattering angle can only be large when, classically speaking, the impact parameter is small, i.e., when the impacting electron penetrates the electron cloud and comes near the core of the molecule. At low incident energy, the electron essentially interacts with the long-range parts of the potential of the target. For homonuclear molecules these are the quadrupole and polarization potentials.[98] In their calculation Kutz and Meyer[98] used the polarizabilities $\alpha_0 = 24.42$ a.u. and $\alpha_2 = 16.293$ a.u. (1 a.u. $= 0.1482 \times 10^{-24}$ cm^3).

In Fig. 11 are compared the close-coupling rotationally summed differential electron scattering cross sections calculated by Kutz and Meyer[98] (solid line) with the experimental values of Gote and Ehrhardt[94] for various incident electron energies. The agreement is good adding credence to the calculation and the underlying assumptions.

The full-model potential calculation results of Kutz and Meyer for the integrated excitation cross section and for the first four rotational excitation channels are shown in Fig. 8. The total scattering cross section (for all scattering channels) has a minimum at about 0.5 eV which was found to be very sensitive to small changes of the potential. The integrated cross section decreases with the final rotational state j. At low scattering angles and electron energies only a few rotational quanta are transferred ("normal" excitation mechanism), whereas at high scattering angles and electron energies many rotational quanta can be exchanged (rotational rainbow mechanism).

Another calculation of rotational excitation of Cl_2 was performed by Ernesti *et al.*[99] within the two-center Coulomb-scattering approximation. This study predicted a rainbow scattering pattern which is consistent both with the close-coupling result and with the experimental data.

3.5.2. Total Vibrational Excitation Cross Section, $\sigma_{vib,t}(\varepsilon)$

There are no experimentally determined total vibrational excitation cross sections, $\sigma_{vib,t}(\varepsilon)$, for Cl_2. There are only the results of two Boltzmann-code calculations,[1,40] based upon limited experimental data. These results are compared in Fig. 12. Their assumed energy dependence is similar (although there is no experimental evidence to support such a shape), and their magnitudes differ. Thus, there is a need for a direct measurement of the vibrational excitation cross section for this molecule and there is also a need for more and better electron transport data to enhance the usefulness of the $\sigma_{vib,t}(\varepsilon)$ calculated from Boltzmann codes.

Vibrational excitation cross sections are important in efforts to model plasma reactors due to their large effect on the electron energy distribution function (see, for example, Refs. 103 and 104). For this reason, we have attempted to deduce a rough estimate of $\sigma_{vib,t}(\varepsilon)$ from the available cross sections for other processes. We assumed the suggested values for $\sigma_{sc,t}(\varepsilon)$ (Sec. 3.1, Fig. 6), $\sigma_{e,t}(\varepsilon)$ (Sec. 3.3, Fig. 9), $\sigma_{i,t}(\varepsilon)$ (to be discussed in Sec. 4.1, Fig. 14), $\sigma_{diss,neut,t}(\varepsilon)$ (to be discussed in Sec. 5, Fig. 16), and $\sigma_{da,t}(\varepsilon)$ (to be discussed in Sec. 6.1, Fig. 17), and took the difference

$$\sigma_{sc,t}(\varepsilon) - [\sigma_{e,t}(\varepsilon) + \sigma_{i,t}(\varepsilon) + \sigma_{diss,neut,t}(\varepsilon) + \sigma_{da,t}(\varepsilon)] \approx \sigma_{vib,t}(\varepsilon) \approx \sigma_{vib,indir}(\varepsilon) \quad (1)$$

to be a measure of $\sigma_{vib,t}(\varepsilon)$. Since direct vibrational excitation for a homopolar molecule such as Cl_2 is expected to be small,[105,106] $\sigma_{vib,t}(\varepsilon)$ may be taken, in this case, to be the cross section for indirect (resonance enhanced) vibrational excitation, $\sigma_{vib,indir}(\varepsilon)$, of the Cl_2 molecule via its temporary negative ion states. Values of $\sigma_{vib,indir}(\varepsilon)$ derived in this way are shown in Fig. 12 (solid line), where the two Boltzmann computed values of $\sigma_{vib,t}(\varepsilon)$ are also shown. The $\sigma_{vib,indir}(\varepsilon)$ deduced in this study bares no similarity to the computed $\sigma_{vib,t}(\varepsilon)$. In spite of the large uncertainty involved in the derivation of $\sigma_{vib,indir}(\varepsilon)$, this deduced cross section shows that the indirect vibrational excitation cross section of Cl_2 is very large. In the absence of any direct measurements of $\sigma_{vib,t}(\varepsilon)$, the present derived cross section $\sigma_{vib,indir}(\varepsilon)$ is preferred to those provided by the Boltzmann codes.

3.5.3. Electronic Excitation Cross Sections, $\sigma_{elec}(\varepsilon)$

There have been no measurements of the cross sections for electron-impact excitation of any of the electronic states of Cl_2. However, there have been three calculations of cross sections for some of the lowest excited electronic states of Cl_2. Rogoff *et al.*[1] report cross sections for electron impact excitation of the electronic states $^3\Pi_u$, $^1\Pi$, and the sum $2\,^1\Pi_u + 2\,^1\Sigma_u^+$ that are derived from a Boltzmann-code analysis. Another Boltzmann-code calculation by Pinhão and Chouki[40] report cross sections for electronic excitation of $^3\Pi_u + {}^1\Pi_u$, $^3\Sigma_u + {}^3\Pi_g + {}^1\Pi_g$, and $2\,^1\Pi_u + 2\,^1\Sigma_u^+$. Also, Rescigno[100] performed close-coupling calculations using the complex Kohn variational method and reported excitation cross sections for $^3\Pi_u$, $^1\Pi_u$, $^3\Pi_g$, $^1\Pi_g$, and $^3\Sigma_u^+$. Rescigno refers to the cross sections he calculated for these five states

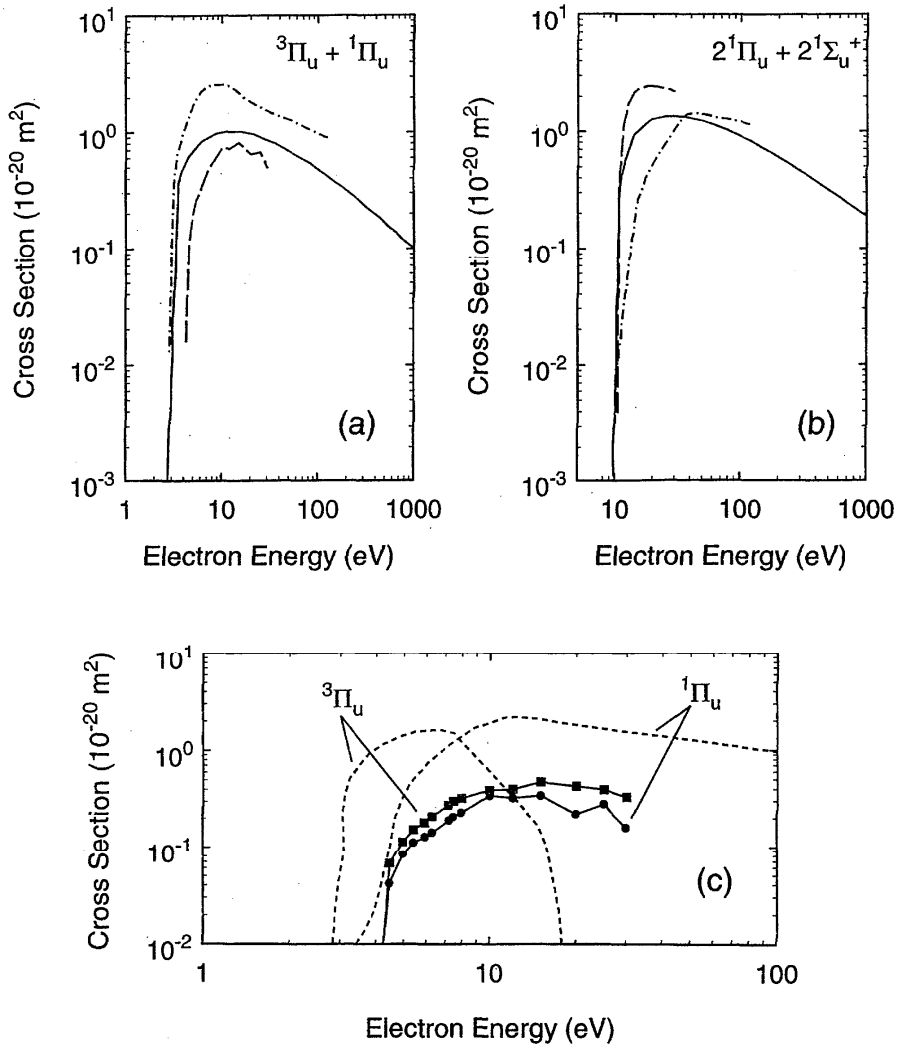


FIG. 13. Comparison of calculated cross sections for electronic excitation, $\sigma_{elec}(\varepsilon)$, of Cl_2. (a) Excitation of $^3\Pi_u+{}^1\Pi_u$: (–·–) Ref. 1; (—) Ref. 40; (– –) Ref. 100. (b) Excitation of $2\,{}^1\Pi_u+2\,{}^1\Sigma_u^+$: (–·–) Ref. 1; (—) Ref. 40; (– –) Ref. 100. (c) Excitation of $^3\Pi_u$ and $^1\Pi_u$: (- - -) Ref. 1; (—) Ref. 100.

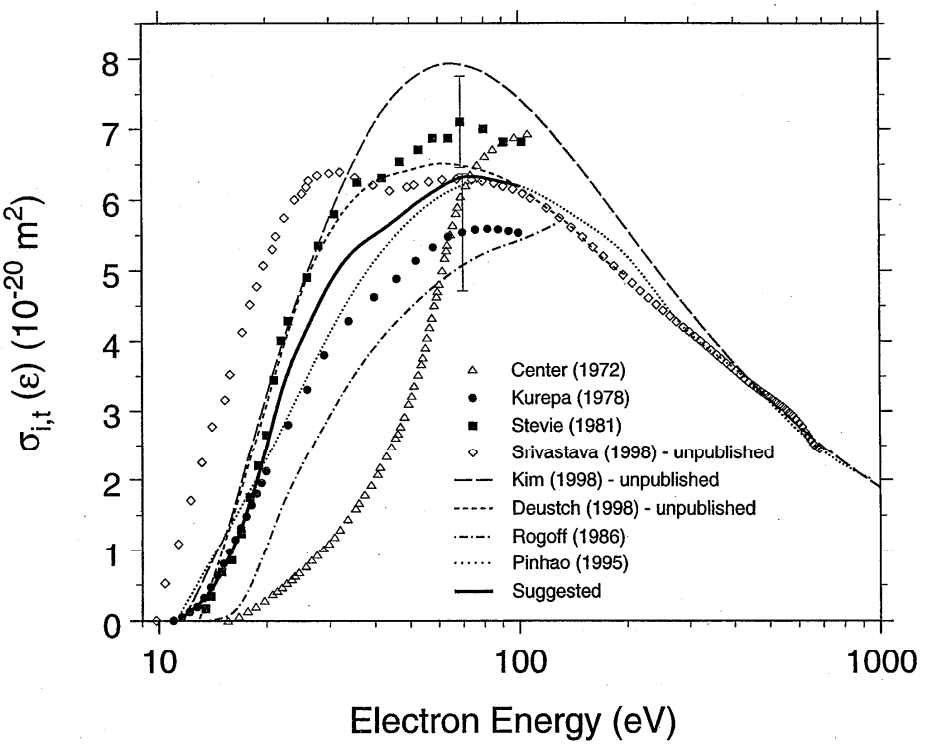
indiscriminately as cross sections for excitation or as cross sections for dissociation, the implication being that all excitations to these states lead to dissociation. This would be consistent with the potential energy curves for the excited states calculated by Peyerimhoff and Buenker[46] (Fig. 1). He also calculated the total cross sections for electron-impact excitation of the $^1\Pi_u$ and $^1\Sigma_u$ Rydberg states of Cl_2 using the Born-dipole approximation and found that the Born-dipole cross sections far exceeded those he calculated using the close-coupling method.

Since the excitation cross sections of Rogoff *et al.*[1] have been used in various plasma models, we compared them with the results of the other two calculations in the few cases where this is possible. Thus, in Fig. 13(a) the cross sections estimated by the three studies for $^3\Pi_u+{}^1\Pi_u$ are compared. In Fig. 13(b) a similar comparison is made for $2\,{}^1\Pi_u+2\,{}^1\Sigma_u^+$. In Fig. 13(c) the cross sections of Rogoff *et al.*[1] and of Rescigno[100] for electron-impact excitation of the electronic states $^3\Pi_u$ and $C\,{}^1\Pi_u$ are compared. The vertical excitation energies of $^3\Pi_u$ and $^1\Pi_u$ are, respectively, 3.31 and ~4.05 eV (see Table 2). The agreement between the Boltzmann-code-deduced electronic excitation cross sections and those of Rescigno is poor. Clearly more work, both experimental and computational, is indicated.

4. Electron Impact Ionization for Cl_2

4.1. Total Ionization Cross Section, $\sigma_{i,t}(\varepsilon)$

In Fig. 14 are compared the available data on the electron-impact total ionization cross section, $\sigma_{i,t}(\varepsilon)$, of Cl_2. These include the measurements by Center and Mandl,[107] Kurepa and Belić,[95] Stevie and Vasile,[108] and Srivastava and Boivin.[109] The Center and Mandl cross section measurements were made using argon as the calibrant gas, and normalizing to the ionization cross section for Ar of Rapp and



G. 14. Electron-impact total ionization cross section, $\sigma_{i,t}(\varepsilon)$, for Cl_2: (△) Ref. 107; (●) Ref. 95; (■) Ref. 108; (◇) Ref. 109; (– –) Ref. 110; (- - -) Ref. .1; (–·–) Ref. 1; (···) Ref. 40; (—) suggested.

nglander-Golden.[112] The stated uncertainty of these mea- ırements is ±15%. Kurepa and Belić's cross section mea- ırements are absolute. They were made in the energy range f 10–100 eV and have a reported relative error of ±20%. elow ~50 eV they are higher than the values obtained by enter and Mandl. The third set of measurements were made y Stevie and Vasile[108] in the energy range 12–100 eV using mass spectrometer and a modulated molecular beam. These eterminations of $\sigma_{i,t}(\varepsilon)$ were made relative to those of the ıree calibrant gases Ar, O_2, and Kr for which they used the spective data of Rapp and Englander-Golden.[112] The val- es plotted in the figure are the averages of the data using ıch of the three calibrant gases. The authors indicated an ror bar in their data for 70 eV as shown in Fig. 14. Their ıcertainties are approximately ±20%. Their measurements gree with those of Kurepa and Belić[95] near the threshold, ıt they are considerably higher for energies greater than 15 eV. Clearly these three sets of data differ not only in ıagnitude, but also in the measured energy dependence of $_{i,t}(\varepsilon)$. The more recent unpublished relative measurements f Srivastava[109] are also shown in Fig. 14. These cover a oader energy range, from threshold to 700 eV, and were bitrarily normalized to the 70 eV point of the "suggested" ırve discussed later and shown by the solid line in Fig. 14. ıterestingly, the cross section of Srivastava shows structure near 25 eV which, although not as evident, is nonetheless indicated by some of the other measurements, and might be due to autoionization.

In Fig. 14 are also shown the results of two recent unpublished calculations, one by Kim[110] and another by Deutsch *et al.*[111] The results of both of these calculations are in reasonable agreement with the measurements of Kurepa and

TABLE 12. Suggested total ionization cross section, $\sigma_{i,t}(\varepsilon)$, for Cl_2

Electron energy (eV)	$\sigma_{i,t}(\varepsilon)$ (10^{-20} m^2)	Electron energy (eV)	$\sigma_{i,t}(\varepsilon)$ (10^{-20} m^2)
11.5	0.03	35	5.26
12	0.11	40	5.49
13	0.25	45	5.68
14	0.43	50	5.87
15	0.69	55	6.03
16	0.99	60	6.15
17	1.32	65	6.25
18	1.67	70	6.32
19	2.06	75	6.33
20	2.47	80	6.31
22	3.25	85	6.28
24	3.79	90	6.25
26	4.17	95	6.22
28	4.51	100	6.19
30	4.80		

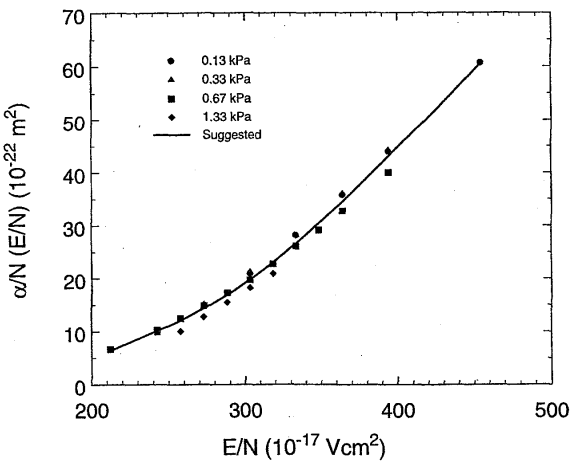


FIG. 15. Density-reduced electron-impact ionization coefficient, $\alpha/N(E/N)$, for Cl_2 at various gas pressures. Data from Ref. 102. The solid line is a least-squares fit to all the data points and represents the suggested values for $\alpha/N(E/N)$.

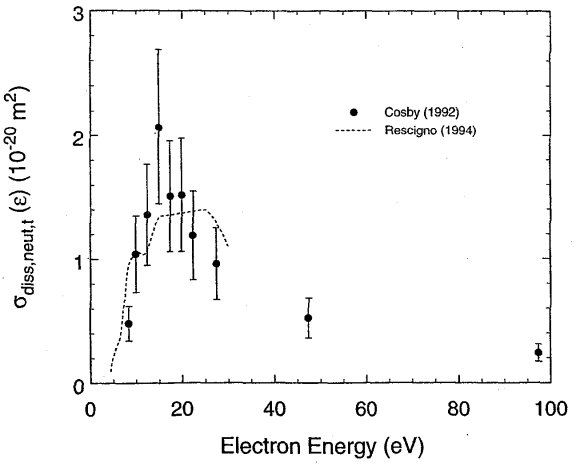


FIG. 16. Total cross section for electron-impact dissociation into neutral fragments, $\sigma_{dis,neut,t}(\varepsilon)$, for Cl_2: (●) measurements by Cosby and Helm from Ref. 114; (- - -) calculations by Rescigno from Ref. 100 [sum of the cross sections for electronic excitation of the lowest five electronic states ($^3\Pi_u$, $^1\Pi_u$, $^3\Pi_g$, $^1\Pi_g$, $^3\Sigma_u^+$) of Cl_2].

Belić[95] and Stevie and Vasile.[108] The calculation of Kim includes multiple ionization but not autoionization.

At the present time we have averaged the measured values of Kurepa and Belić[95] and Stevie and Vasile,[108] even though the differences in their magnitudes exceed their combined uncertainties, and take this to be our suggested value for the $\sigma_{i,t}(\varepsilon)$ of Cl_2. We have not included the values of Center and Mandl[107] due to the obviously inconsistent shape of their cross section when compared to the other[95,108] measured values. These average values are shown by the bold line in Fig. 14 (Table 12).

The model-dependent total ionization cross section of Rogoff *et al.*,[1] and Pinhão and Chouki[40] deduced from modeling of chlorine discharges are also plotted in Fig. 14. While the Pinhão and Chouki cross section is in general agreement with the most reliable measurements, that of Rogoff *et al.* is not. However, such a comparison is biased by the input cross section assumed by each calculation.

Threshold ionization energies leaving the Cl_2^+ ion in various states of excitation have been given in Table 4. Also listed in Table 4 are the values for the threshold energy for dissociative ionization ($Cl_2+e\rightarrow Cl^++Cl+2e$) and for double ionization.

There seem to be no cross section data on either the partial ionization, or the cross sections for multiple ionization of Cl_2 by electron impact. Therefore, the relative production of Cl_2^+ and Cl^+ by electron impact is not known. Photoabsorption measurements, however, show that the production of Cl_2^+ far exceeds the production of Cl^+ for dissociative photoionization (see Fig. 3).

4.2. Density-Reduced Electron-Impact Ionization Coefficient, $\alpha/N(E/N)$

The only measurement of the density-reduced electron-impact ionization coefficient, $\alpha/N(E/N)$, of Cl_2 is that of Božin and Goodyear[102] shown in Fig. 15. These measurements were made at $T=293$ K for Cl_2 pressures of 0.13, 0.33, 0.67, and 1.33 kPa. From a least-square fit to the data in Fig. 15, we obtained the values listed in Table 13 which represent our suggested values for the $\alpha/N(E/N)$ of Cl_2.

TABLE 13. Suggested density-reduced electron-impact ionization coefficient, $\alpha/N(E/N)$, for Cl_2 (based on measurements of Božin and Goodyear from Ref. 102)

E/N (10^{-17} V cm^2)	$\alpha/N(E/N)$ (10^{-22} m^2)	E/N (10^{-17} V cm^2)	$\alpha/N(E/N)$ (10^{-22} m^2)
213	6.45	340	28.2
220	7.34	360	33.4
240	9.82	380	39.0
260	12.4	400	44.7
280	15.5	420	50.4
300	19.2	440	56.2
320	23.4	450	59.1

TABLE 14. Total cross section for electron-impact dissociation into neutral fragments, $\sigma_{diss,neut,t}(\varepsilon)$, for Cl_2 (data of Cosby and Helm from Ref. 114)

Electron energy (eV)	$\sigma_{diss,neut,t}(\varepsilon)$ (10^{-20} m^2)
8.4	0.48±0.14
9.9	1.04±0.31
12.4	1.36±0.41
14.9	2.07±0.62
17.4	1.51±0.45
19.9	1.52±0.46
22.4	1.19±0.36
27.4	0.96±0.29
47.4	0.52±0.16
97.4	0.24±0.07

TABLE 15. Negative ion states of Cl_2

Energy (eV)	Assigned symmetry of corresponding negative ion state	Method/Reference
0.03 ± 0.03 2.5 ± 0.15 5.5 ± 0.15	$^2\Pi_g$ [a] $^2\Pi_u$ [a] $^2\Sigma_g^+$ [a]	Maxima in the dissociative electron attachment cross section measured in an electron-impact mass-spectrometric study (87)
0.0 2.5 ± 0.05 5.75 ± 0.05 9.7[b]	$^2\Sigma_u^+$ $^2\Pi_g$ $^2\Pi_u$ $^2\Sigma_g^+$	Maxima in the dissociative electron attachment cross section measured in an electron-impact mass-spectrometric study (80, 95)
0 2.5 5.5	$^2\Sigma_u^+$ $^2\Pi_g$ $^2\Pi_u$	Dissociative electron attachment using a crossed-beam electron impact spectrometer. Assignments based on angular distribution analysis of the Cl^- ions (96)[c]
0.07	$^2\Sigma_u^+$	Electron swarm (117)
2.4 ± 0.1		Electron-impact mass spectrometry (76)
7.46 ($v=0$) and subsequent peaks separated by 0.08 eV corresponding to $v=1-5$	Electron-excited Feshbach resonances formed by addition of two $ns\sigma$ electrons to the $X\,^2\Pi_g$ positive-ion core	Electron transmission (97)
~2[d] ~5[d]	$^2\Pi_g$ $^2\Pi_u$	Studies of Cl^- ions produced by dissociative electron attachment from *condensed* Cl_2 (118)

se assignments are incorrect, see text.
a *et al.* (Ref. 96) did not observe the 9.7 eV resonance indicated by the data of Kurepa and Belić (Ref. 95).
ording to Azria *et al.* (Ref. 96), there may be a small contribution of the $^2\Sigma_g^+$ state of Cl_2^- to the Cl^- formation at the low-energy side of the resonance 5 eV.
se values are about 0.5 eV lower than the corresponding gaseous data. This may be due to the effect of the polarization energy of condensed Cl_2 on the itive-ion states of the isolated Cl_2 molecule (Refs. 118 and 119).

Total Cross Section for Electron-Impact Dissociation Into Neutral Fragments, $\sigma_{diss,neut,t}(\varepsilon)$ for Cl_2

here has been one measurement[113,114] of the total cross ion for electron impact dissociation into neutrals, $_{s,neut,t}(\varepsilon)$, for Cl_2, and these data of Cosby and n[113,114] are shown in Fig. 16. In Fig. 16 is also shown the of the cross sections calculated by Rescigno[100] for the est five excited electronic states ($^3\Pi_u$, $^1\Pi_u$, $^3\Pi_g$, $^1\Pi_g$,) of Cl_2 which are reached by promoting an occupied nce electron into the antibonding ($5\sigma_u$) orbital. The cal- tion by Rescigno showed that the total dissociation cross ion is the largest for the $^3\Pi_u$ state up to the highest gy (30 eV) he investigated. The agreement between Re- no's calculations and the experimental data is good, sup- ing the premise that all electronic excitations result in ociation. The experimental data of Cosby and Helm[113,114] listed in Table 14 as the presently suggested values for $_{s,neut,t}(\varepsilon)$ for the Cl_2 molecule.

n an earlier study, Wells and Zipf[115] observed dissocia- excitation of Cl_2 and identified the fragments as, in part, nic chlorine in long-lived high-Rydberg excited states * (HR)] produced through

$$e+Cl_2\rightarrow e+Cl+Cl^*(HR) \quad (2)$$

and

$$e+Cl_2\rightarrow 2e+Cl^++Cl^*(HR). \quad (3)$$

They associated an energy threshold for reactions (2) and (3), respectively, equal to 14.8 ± 1 eV and 29.2 ± 5 eV.

Another process for neutral fragment production is dissociative recombination ($e+Cl_2^+\rightarrow Cl+Cl$). No data exit on this process (see Mitchel[116] for data on this process for other species).

6. Electron Attachment to Cl_2

As we have discussed in Sec. 2.2, the Cl_2^- negative ion consisting of Cl^- ($^1S_0$) and Cl ($^2P_{3/2,1/2}$) has four electronic states whose order of increasing energy is: $^2\Sigma_u^+$, $^2\Pi_g$, $^2\Pi_u$, $^2\Sigma_g^+$ (Fig. 5). The participation of these states in dissociative electron attachment of Cl_2 depends on the way their potential-energy curves cross the ground-state potential energy curve $X\,^1\Sigma_g^+$ of the neutral Cl_2 molecule. On the basis of the Cl_2^- potential-energy curves in Fig. 5, one would expect formation of the parent anion Cl_2^- at zero energy, and

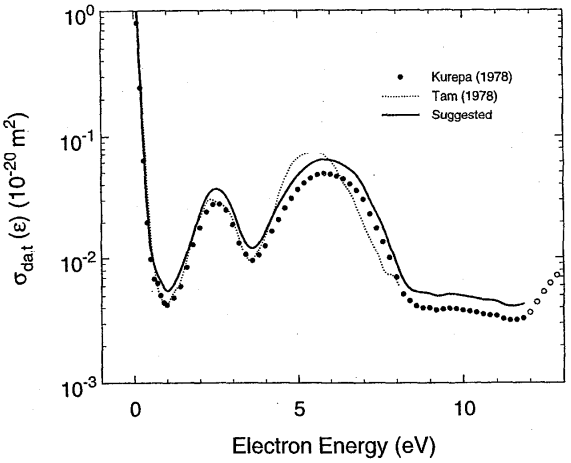


FIG. 17. Total dissociative electron attachment cross section, $\sigma_{da,t}(\varepsilon)$, for Cl_2: (●) measurements of Kurepa and Belić from Ref. 95; (···) relative cross section for the production of Cl^- from Cl_2 measured by Tam and Wong in Ref. 87 normalized to the Kurepa and Belić cross section at 2.5 eV; (—) cross section of Kurepa and Belić (Ref. 95) adjusted upwards by 30%. The open symbols represent the contribution to the measured cross section attributed to ion-pair production (see Sec. 6.5.).

the formation of Cl^- at near-zero energy and in three higher-energy ranges below 10 eV. For the fragment negative ion Cl^-, the dissociative attachment reactions

$$Cl_2(X\,^1\Sigma_g^+)+e\rightarrow Cl_2^-{}^*\rightarrow Cl^-(^1S_0)+Cl(^2P_{3/2,1/2}) \quad (4)$$

involve the ground state $X\,^1\Sigma_g^+$ of Cl_2 in the $\nu=0$ and perhaps $\nu=1$ state, and the four $^2\Sigma_u^+$, $^2\Pi_g$, $^2\Pi_u$, $^2\Sigma_g^+$ negative ion states of Cl_2 which are correlated with the dissociation limit $Cl^-(^1S_0)+Cl(^2P_{3/2,1/2})$. This limit lies 1.10 eV [Cl ($^2P_{1/2}$)] and 1.21 eV [Cl ($^2P_{3/2}$)] below the minimum of the potential energy curve for the ground state of Cl_2 [see Fig. 5(b)].

Three electron-beam experimental studies[87,95,96] have shown that the yield of Cl^- from Cl_2 exhibits three peaks: at ~0 eV, at 2.5 eV, and at 5.5 eV (Table 15). These were ascribed[95,96] to the $^2\Sigma_u^+$, $^2\Pi_g$, and $^2\Pi_u$ resonant states of Cl_2^-, respectively. The ground state, $^2\Sigma_u^+$, of Cl_2^- is formed by addition of an extra electron to the lowest unfilled ($\sigma_u 3p$) Cl_2 orbital of the ground-state electron configuration of Cl_2: $[(...)(\sigma_g 3p)^2(\pi_u 3p)^4(\pi_g 3p)^4]$. The core-excited $^2\Pi_g$ and $^2\Pi_u$ states of Cl_2^- are formed by exciting one electron of the $^2\Sigma_u^+$ shape resonance from the $\pi_g 3p$ and $\pi_u 3p$ to the $\sigma_u 3p$ orbital, respectively. An electron-transmission study by Spence[97] located the lowest-lying electron-excited Feshbach resonance in Cl_2 at 7.46 eV. He associated this resonance with Rydberg states having symmetry $(X\,^2\Pi_g)(4s\sigma)^2[^2\Pi_{1/2,3/2}]$. The derivative of the transmitted current in Cl_2 between 7.0 and 9.0 eV showed a progression of six resonances starting at 7.46 eV, with an average spacing between adjacent resonances of 80 meV. A recent high resolution (~60 meV FWHM) electron beam study[120] of dissociative electron attachment to Cl_2 between ~0.0 and 0.7 eV showed two resonances at 0.03 and 0.250 eV. The former peak has been attributed[120] to dissociative electron attachment via the $^2\Sigma_u^+$ state of Cl_2^-. The latter may be due to dissociative electron attachment via one of the excited $^2\Pi$ states of Cl_2^-.[120]

The parent negative ion Cl_2^- is not normally formed in the gas phase. The transient anion in the lowest negative ion state, $Cl_2^-{}^*(^2\Sigma_u^+)$, must be collisionally stabilized before it breaks up by dissociative electron attachment. Since, moreover, dissociative electron attachment occurs at subpicosecond times, collisional stabilization of $Cl_2^-{}^*$ can only take place at high gas densities when the collisional stabilization time becomes comparable to, or shorter than, the dissociative electron attachment time, or in the condensed phase. No parent negative ions have been observed in electron attachment studies in the gas phase. They, however, have been observed in gas-phase negative-ion charge transfer reactions[121,122] and in the condensed phase.[118] With regard to the latter-type investigation, Azria *et al.*[118] studied the production of Cl^- by dissociative electron attachment in electron-stimulated desorption from Cl_2 condensed on a platinum substrate. They found that the energy dependence of the Cl^- signal exhibits two peaks at about 2 and 5 eV which they attributed to the $^2\Pi_g$ and $^2\Pi_u$ Cl_2^- resonant states. Thus, in the condensed phase (in the chlorine lattice on the surface of the substrate) the dissociation dynamics of Cl_2^- are similar to those in the gas phase except possibly with a 0.5 eV downward shift in the resonance energy positions. (See Christophorou[119,123] for a discussion of the effect of the medium and state of matter on the energetics of negative ion states.)

6.1. Total Dissociative Electron Attachment Cross Section, $\sigma_{da,t}(\varepsilon)$

Dissociative electron attachment to Cl_2 is rather simple in its products: only Cl^- is produced directly. Thus, electron beam experiments with mass analysis and total electron attachment experiments without mass analysis should yield the same results. In spite of this, it seems that the only absolute measurement of the total dissociative electron attachment cross section, $\sigma_{da,t}(\varepsilon)$, of Cl_2 is that of Kurepa and Belić.[95] Their data are shown in Fig. 17. They cover the energy range from 0 to 13.0 eV and have an uncertainty of ±20%. They indicate that dissociative electron attachment to Cl_2 principally proceeds via three negative-ion states located at ~0 eV, (2.5±0.05) eV, and (5.75±0.05) eV. A weak process they observed between 9 and 11.5 eV was not observed by others[96] (see Table 15).

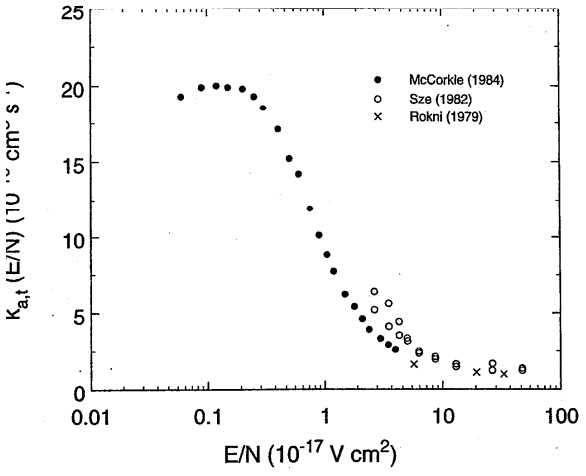
In Fig. 17 is also plotted, for comparison, the relative cross section for the production of Cl^- from Cl_2 as determined in a higher-energy resolution study by Tam and Wong.[87] (Note that the energy scale for Cl^-/Cl_2 in Fig. 2 of the paper of Tam and Wong is not that indicated on the energy axis of the figure in their paper.) Here the data of Tam and Wong have been normalized to the Kurepa and Belić cross section at 2.5 eV. Other than the small differences in the shape and energy position of the resonance at ~5 eV, the overall shapes of the two cross sections are in reasonable agreement. The sharp peak at zero energy is worth noting as it is consistent with

BLE 16. Suggested total dissociative electron attachment cross section, $_{,t}(\varepsilon)$, for Cl_2

ectron energy (eV)	$\sigma_{da,t}(\varepsilon)$ (10^{-20} m^2)	Electron energy (eV)	$\sigma_{da,t}(\varepsilon)$ (10^{-20} m^2)
0.05	1.83	5.2	0.053
0.10	1.04	5.6	0.062
0.20	0.32	6.0	0.062
0.30	0.081	6.2	0.060
0.40	0.026	6.6	0.052
0.50	0.013	7.0	0.039
0.60	0.0088	7.2	0.030
0.80	0.0065	7.6	0.018
1.0	0.0055	8.0	0.0091
1.2	0.0062	8.2	0.0066
1.6	0.011	8.6	0.0053
2.0	0.024	9.0	0.0051
2.2	0.032	9.2	0.0049
2.6	0.036	9.6	0.0051
3.0	0.025	10.	0.0049
3.2	0.018	10.2	0.0048
3.6	0.012	10.6	0.0046
4.0	0.017	11.0	0.0045
4.2	0.022	11.2	0.0042
4.6	0.033	11.6	0.0041
5.0	0.047	11.8	0.0043

: electron swarm data (Sec. 6.2). The energy positions of : negative ion resonances as determined in the study of m and Wong along with the Tam and Wong assignments : compared with other data in Table 15. Comparison with ıer studies indicates that the assignments of Tam and ong are apparently incorrect. The sequence of their assign-:nts is in error because their calculations show the poten-l energy curve for the $^2\Sigma_u^+$ anionic state not crossing the tential energy curve for the $X\,^1\Sigma_g^+$ ground state of the Cl_2)lecule.

Based on the analysis of the total electron attachment rate nstant in Sec. 6.2.2, the values of $\sigma_{da,t}(\varepsilon)$ given by Kurepa



. 18. Total electron attachment rate constant as a function of E/N, (E/N), for Cl_2 ($T\approx298$–300 K); (●) Ref. 117; (○) Ref. 124; (×) f. 88.

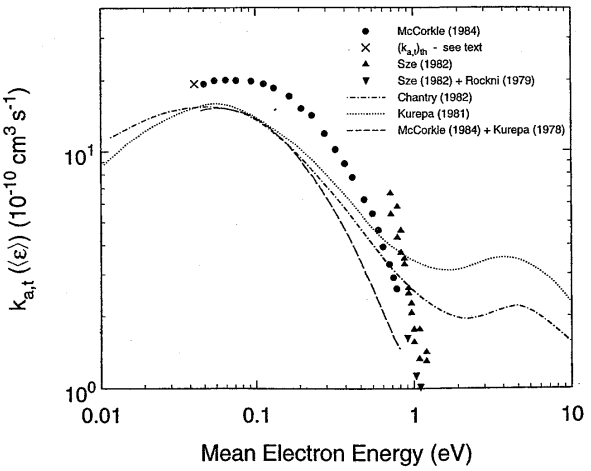


FIG. 19. Total electron attachment rate constant as a function of the mean electron energy $\langle\varepsilon\rangle$, $k_{a,t}(\langle\varepsilon\rangle)$, for Cl_2 ($T\approx298$ K): (●) Ref. 117; (×) $(k_{a,t})_{th}$ determined from the average of the two most recent values of the thermal ($T\approx300$ K) electron attachment rate. see Table 18: (▲) Ref. 124: (▼) Ref. 124 using the rate constants measured by Rockni *et al.* (Ref. 88); (–·–) Ref. 127 using the $\sigma_{a,t}(\varepsilon)$ of Kurepa and Belić (Ref. 95) and a Maxwellian distribution function for the electron energies; (···) Ref. 128 using the $\sigma_{a,t}(\varepsilon)$ of Kurepa and Belić (Ref. 95) and a Maxwellian distribution function for the electron energies; (- - -) Ref. 117 using the $\sigma_{a,t}(\varepsilon)$ of Kurepa and Belić (Ref. 95) and the electron energy distribution functions they calculated for N_2.

and Belić appear to be approximately 30% lower than indicated by the electron swarm measurements. We have, thus, adjusted their cross section upwards by this percentage for our suggested values for $\sigma_{da,t}(\varepsilon)$. This adjusted cross section is shown in the figure by the solid line, and values taken from this curve are listed in Table 16 as our suggested data for the $\sigma_{da,t}(\varepsilon)$ of the Cl_2 molecule.

6.2. Total Electron Attachment Rate Constant as a Function of the Density-Reduced Electric Field E/N, $k_{a,t}(E/N)$, and the Mean Electron Energy $\langle\varepsilon\rangle$, $k_{a,t}(\langle\varepsilon\rangle)$

6.2.1. $k_{a,t}(E/N)$ in N_2

McCorkle *et al.*[117] measured the total electron attachment rate constant, $k_{a,t}(E/N)$, of Cl_2 using mixtures of Cl_2 with N_2. Their measurements covered the E/N range of 6×10^{-19}–4×10^{-17} V cm^2, with a probable uncertainty of ±10%. The measurements were made at room temperature (298 K) and also at other temperatures above and below ambient (Sec. 6.2.4). The total gas number density in their experiments was 6.48×10^{19} molecules/cm^3 and the Cl_2 gas number density was in the range $(0.2$–$2.3)\times10^{15}$ molecules/cm^3. The rate constant was found to be independent of both the total and the attaching gas pressures. The measurements at room temperature are plotted in Fig. 18.

Another measurement of $k_{a,t}(E/N)$ was made by Sze *et al.*[124] using mixtures of Cl_2 with N_2. These measurements were made at 300 K and for only one mixture composition [the Cl_2 gas number density in the mixture was 260 parts per

TABLE 17. Suggested total electron attachment rate constant, $k_{a,t}(\langle\varepsilon\rangle)$ (T = 298 K), for Cl_2 (data from Ref. 117)

$\langle\varepsilon\rangle$ (eV)	$k_{a,t}(\langle\varepsilon\rangle)$ (10^{-10} cm^3 s^{-1})
0.046	19.3 (1.20)[a]
0.054	19.9 (0.9)
0.064	20.0 (0.9)
0.075	19.9 (1.1)
0.094	19.8 (1.2)
0.113	19.3 (1.3)
0.131	18.5 (1.3)
0.165	17.1 (1.5)
0.196	15.2 (1.4)
0.228	14.2 (1.5)
0.275	11.9 (1.5)
0.322	10.1 (1.2)
0.368	8.8 (1.2)
0.411	7.7 (1.4)
0.487	6.2 (1.1)
0.550	5.4 (0.9)
0.599	4.6 (0.7)
0.640	3.9 (0.5)
0.704	3.3 (0.5)
0.745	2.9 (0.5)
0.779	2.6 (0.5)

[a]Values in parentheses are standard deviations as given by the authors.

million (ppm)]. These measurements are also plotted in Fig. 18, along with the limited measurements made by Rokni *et al.*[88] at 300 K. With the exception of the measurements of Sze *et al.* below $\sim 5\times10^{-17}$ V cm^2, these data are not incompatible with those of McCorkle *et al.*

Besides their measurements in mixtures with N_2, Sze *et al.*[124] also reported $k_{a,t}(E/N)$ for one mixture of Cl_2 in Ar. These data are not included in the present paper since the measurements were made for only one mixture concentration (260 ppm) and the effect of Cl_2 on the electron energy distribution function in pure Ar could not be assessed. For the same reason, early measurements by Bradbury[125] on the probability of electron attachment per collision in a Cl_2/Ar mixture are not included.

6.2.2. $k_{a,t}(\langle\varepsilon\rangle)$

McCorkle *et al.*[117] used their measurements of $k_{a,t}(E/N)$, and the electron energy distribution functions for N_2 they calculated at each E/N for which they measured $k_{a,t}$ using a Boltzmann code, and determined the $k_{a,t}(\langle\varepsilon\rangle)$ for Cl_2. These derived data are shown in Fig. 19 for T=298 K. In this figure is plotted also the thermal value, $(k_{a,t})_{th}$, of $k_{a,t}(\langle\varepsilon\rangle)$ as given by the average of the two most recent measurements[117,126] of this quantity (Sec. 6.2.3). In addition, values of $k_{a,t}(\langle\varepsilon\rangle)$ reported by the following four groups of investigators are plotted in the figure:

(1) Values reported by Sze *et al.*[124] determined from their $k_{a,t}(E/N)$ measurements and also from the measurements of Rokni *et al.*[88] These are in fair agreement with the McCorkle *et al.*[117] data.

TABLE 18. Thermal values, $(k_{a,t})_{th}$, of the total electron attachment rate constant for Cl_2 near room temperature

$(k_{a,t})_{th}$ (10^{-10} cm^3 s^{-1})	Temperature (K)	Reference
2.8±0.4	300	129
3.1	293	130
11.0	300	131
18.6±1.2	298	117
20.0±3.0	300	126
37±17	350	132

(2) Values estimated by McCorkle *et al.* using the electron energy distributions in N_2 and the total electron attachment cross section of Kurepa and Belić[95] (Fig. 17). While the Kurepa and Belić-based $k_{a,t}(\langle\varepsilon\rangle)$ have a similar energy dependence to the directly measured rate constants, they are lower in magnitude (at a mean electron energy of 0.08 eV by ~30%) indicating that the Kurepa and Belić cross sections are lower than their true values.

(3) Values of $k_{a,t}(\langle\varepsilon\rangle)$ determined by Chantry[127] and by Kurepa *et al.*[128] using the total electron attachment cross section of Kurepa and Belić[95] (Fig. 17) and a Maxwellian electron energy distribution function. Clearly the assumption of a Maxwellian distribution function for the electron energies is unrealistic at high E/N, as is shown by the large difference between the calculated $k_{a,t}(\langle\varepsilon\rangle)$ and the experimental measurements of $k_{a,t}(\langle\varepsilon\rangle)$. The data of McCorkle *et al.*[117] are listed in Table 17 as our suggested values for the $k_{a,t}(\langle\varepsilon\rangle)$ of Cl_2 at 298 K.

Values of $k_{a,t}(\langle\varepsilon\rangle)$ derived from limited measurements in mixtures of Cl_2 with argon[88,124] are uncertain and are not included in this work.

6.2.3. Thermal Value, $(k_{a,t})_{th}$, of the Total Electron Attachment Rate Constant

In Table 18 are listed the values of the electron attachment rate constant at thermal energies, $(k_{a,t})_{th}$ (T=300 K). These are independent measurements by various groups[117,126,129–132] and they vary significantly. The two most recent measurements[117,126] are consistent with each other and we take their average, 19.3×10^{-10} cm^3 s^{-1}, as the best present estimate of $(k_{a,t})_{th}$ for T=300 K (plotted on Fig. 19 as a × symbol).

6.2.4. Effect of Temperature on the Electron Attachment Rate Constant, $k_{a,t}(\langle\varepsilon\rangle, T)$

There have been two measurements[117,126] of the dependence of the total electron attachment rate constant $k_{a,t}$ of Cl_2 on gas temperature. The measurements of McCorkle *et al.*[117] were made at various mean electron energies from thermal to 0.78 eV, and the measurements of Smith *et al.*[126] were made at only thermal energies. The former results are reproduced in Fig. 20(a), and the latter are compared with the former in Fig. 20(b). All of the data for $(k_{a,t})_{th}$ are tabulated in Table 19.

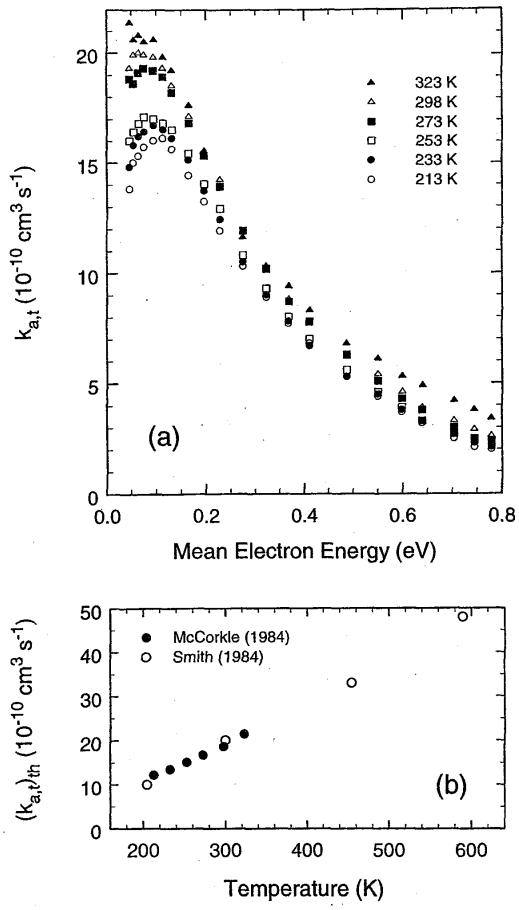


FIG. 20. (a) Variation of $k_{a,t}(\langle\varepsilon\rangle)$ of Cl_2 with temperature from McCorkle *et al.* (Ref. 117). (b) Variation of $(k_{a,t})_{th}$ of Cl_2 with temperature: (●) Ref. 117; (○) Ref. 126.

6.3. Density-Reduced Electron Attachment Coefficient, $\eta/N(E/N)$

The early measurements of $\eta/N(E/N)$ by Bailey and Healey[101] in Cl_2 at 288 K are not in agreement with the more recent measurements of Božin and Goodyear[102] made at 293 K (Fig. 21). Božin and Goodyear indicated an uncertainty of

TABLE 19. Variation of $(k_{a,t})_{th}$ of Cl_2 with temperature

Temperature (K)	$(k_{a,t})_{th}$ (10^{-10} cm^3 s^{-1})	Reference
213	12.2	117
233	13.5	
253	15.1	
273	16.7	
298	18.6	
323	21.4	
203	<10	126
300	20	
455	33	
590	48	

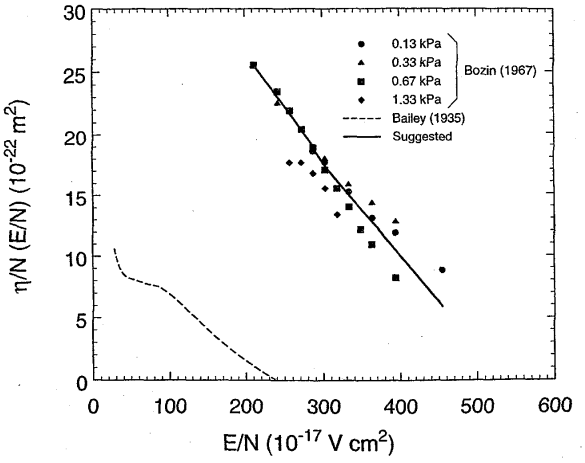


FIG. 21. Density-reduced electron attachment coefficient, $\eta/N(E/N)$, for Cl_2: (●), (▲), (■), (♦) Ref. 102; (- - -) Ref. 101; (—) suggested values.

±10%, but the average uncertainty of their data is more likely twice this value. They also indicated that their measurements for 1.33 kPa may be more uncertain than those at the other three pressures (see Fig. 21). Therefore, the solid line in Fig. 21 is a least-squares fit to the data of Božin and Goodyear at pressures of 0.13, 0.33, and 0.67 kPa, and data taken off this curve are listed in Table 20. In the absence of other measurements, these values are presently suggested for the $\eta/N(E/N)$ of Cl_2, but clearly there is a need for further measurements.

6.4. Density-Reduced Effective Ionization Coefficient, $(\alpha-\eta)/N(E/N)$

Božin and Goodyear[102] reported measurements of the density-reduced effective ionization coefficient $(\alpha-\eta)/N(E/N)$ for pure Cl_2. Their measurements were made at room temperature ($T=293$ K) for gas pressures of 0.13, 0.33, 0.67, and 1.33 kPa. Figure 22 shows their data which have a stated uncertainty of ±10%. The solid curve is a least-squares fit to the data at all pressures, and values taken off this curve are listed in Table 21 as the presently suggested estimates of the $(\alpha-\eta)/N(E/N)$ for pure chlorine.

TABLE 20. Suggested values for the density-reduced electron attachment coefficient, $\eta/N(E/N)$, for Cl_2 (data of Božin and Goodyear from Ref. 102)

E/N (10^{-17} V cm^2)	$\eta/N(E/N)$ (10^{-22} m^2)	E/N (10^{-17} V cm^2)	$\eta/N(E/N)$ (10^{-22} m^2)
215	25.3	350	13.7
225	24.4	375	11.9
250	22.3	400	10.0
275	20.0	425	8.14
300	17.6	450	6.26
325	15.6		

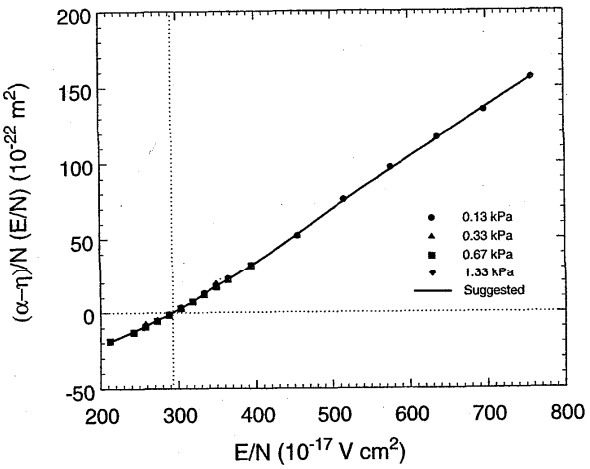


FIG. 22. Density-reduced effective ionization coefficient, $(\alpha-\eta)/N(E/N)$, for Cl_2 from Božin and Goodyear (Ref. 102). The solid line represents the suggested values.

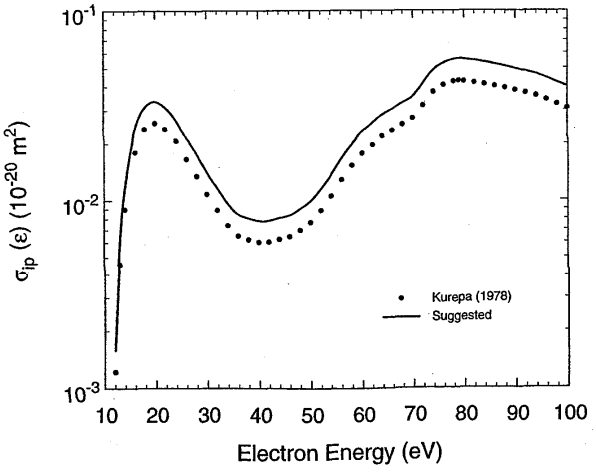


FIG. 23. Cross section for ion-pair formation, $\sigma_{ip}(\varepsilon)$, for Cl_2 of Kurepa and Belić from Ref. 95. The solid curve is the same data adjusted upward by 30% as discussed in Secs. 6.1 and 6.2.

6.5. Cross Section for Ion-Pair Formation, $\sigma_{ip}(\varepsilon)$

Besides the formation of negative ions via the resonant electron attachment processes discussed in the preceding sections, negative ions have been observed to form by electron impact on Cl_2 via the direct process of polar dissociation, that is, via the ion-pair process

$$Cl_2+e\rightarrow Cl^++Cl^-+e. \quad (5)$$

The energy onset for process (5) is 11.9 eV (Table 4). The absolute cross section measurements of Kurepa and Belić[95] for negative-ion formation above 11.9 eV (see Fig. 17) is due largely to the ion-pair process, Eq. (5), although in some energy regions contributions from indirect electron attachment processes are possible. The Kurepa and Belić data (quoted relative uncertainty ±20%) are plotted in Fig. 23. As discussed in Secs. 6.1 and 6.2, these data need to be adjusted upward by 30%, and the so-adjusted data are shown in Fig. 23 by the solid line. Data taken off this line are listed in Table 22 as our suggested values for the $\sigma_{ip}(\varepsilon)$ of Cl_2.

6.6. Negative Ions in Cl_2 Discharges

There have been a number of studies dealing with negative ions in Cl_2 gas discharges. By way of example we refer in this section to two such studies,[133,134] dealing with laser-induced photodetachment of negative ions and its use to infer the density of negative ions in the plasma. Han *et al.*[133] described a technique for sampling negative ions in the hollow cathode and hollow anode of Cl_2/N_2 discharges. The photoelectron transient signals which were induced by laser photodetachment of the negative ions present in the discharge were employed to probe the ion concentration. The observed negative-ion transient signal allowed a study of the kinetics of the three negative ions (Cl^-, Cl_2^-, and Cl_3^-) they observed in the discharge. Interestingly, the authors concluded from their measurements that the Cl_3^- ion is likely to be due to the recombination of Cl^- and Cl_2, and the Cl_2^- ion is likely to be the result of three-body electron attachment to Cl_2. Hebner[134] also employed laser photodetachment spectros-

TABLE 21. Suggested values of the density-reduced effective ionization coefficient, $(\alpha-\eta)/N(E/N)$, for Cl_2 (data of Božin and Goodyear from Ref. 102)

E/N (10^{-17} V cm^2)	$(\alpha-\eta)/N(E/N)$ (10^{-22} m^2)	E/N (10^{-17} V cm^2)	$(\alpha-\eta)/N(E/N)$ (10^{-22} m^2)
215	−18.5	500	69.1
250	−10.7	550	86.8
300	2.13	600	103.7
350	18.0	650	120.3
400	33.7	700	136.9
450	50.9	750	153.4

TABLE 22. Suggested cross section for negative ion–positive ion pair production, $\sigma_{ip}(\varepsilon)$, in Cl_2 between 12 and 100 eV (adjusted data of Kureba and Belić from Ref. 95)

Electron energy (eV)	$\sigma_{ip}(\varepsilon)$ (10^{-20} m^2)	Electron energy (eV)	$\sigma_{ip}(\varepsilon)$ (10^{-20} m^2)
12	0.0016	52	0.0114
13	0.0060	56	0.0166
14	0.0117	60	0.0229
16	0.0234	62	0.0255
18	0.0312	66	0.0304
20	0.0335	70	0.0354
22	0.0312	72	0.0413
26	0.0216	76	0.0528
30	0.0140	80	0.0553
32	0.0116	82	0.0546
36	0.0085	86	0.0525
40	0.0078	90	0.0497
42	0.0078	92	0.0481
46	0.0083	96	0.0442
50	0.0099	100	0.0399

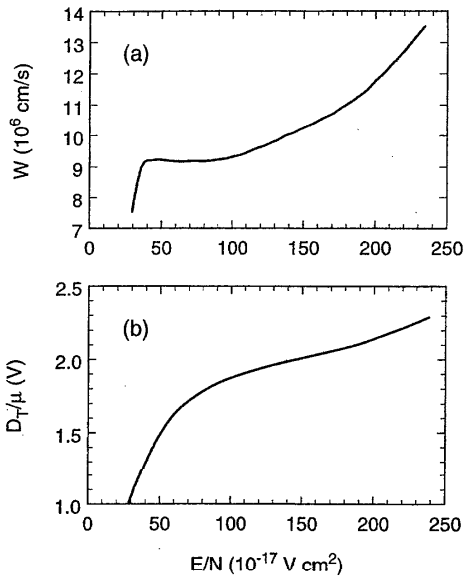


FIG. 24. (a) Electron drift velocity, w, for Cl_2 (T=288 K), from data of Bailey and Healey (Ref. 101). (b) Ratio of the lateral electron diffusion coefficient to electron mobility, D_T/μ, for Cl_2 (T=288 K) derived from data given by Bailey and Healey (Ref. 101). Both data sets are uncertain.

copy to infer the density of chlorine negative ions in low pressure, inductively coupled chlorine plasmas.

7. Electron Transport for Cl_2

7.1. Electron Drift Velocity, w

The only known measurements of electron drift velocity, w, for Cl_2 are those made in 1935 by Bailey and Healey[101] using the volume mixtures 20%Cl_2:80%He, 20%Cl_2:80%CO_2, and 40%Cl_2:60%CO_2. Bailey and Healey also showed a curve for w vs E/N which they identified as the w for Cl_2. This curve is reproduced in Fig. 24(a), but it is considered uncertain because of its indirect determination from the drift velocities they measured in the mixtures just mentioned. Clearly, measurements of $w(E/N)$ for Cl_2 are needed, and efforts are underway to measure this quantity in our laboratory for Cl_2 and its dilute mixtures in argon.[135]

7.2. Lateral Electron Diffusion Coefficient to Electron Mobility Ratio, D_T/μ

There are no measurements of this quantity for Cl_2. Bailey and Healey[101] reported measurements of the quantity k_M (mean energy of agitation of electrons in terms of the mean energy of agitation of the gas molecules at 288 K) which we used to determine the ratio of the lateral electron diffusion coefficient D_T to electron mobility μ, via the relationship $D_T/\mu = k_M(kT/e)$. The values of D_T/μ determined this way are shown in Fig. 24(b). They should be considered uncertain. Measurements of D_T/μ over a wide range of E/N are needed.

8. Optical Emission from Cl_2 Gas Discharges

There have been a number of studies of light emission from chlorine excited in an electrical discharge (see, for instance, Refs. 136–142). It is clear from these investigations that the emission spectrum from a chlorine vapor discharge

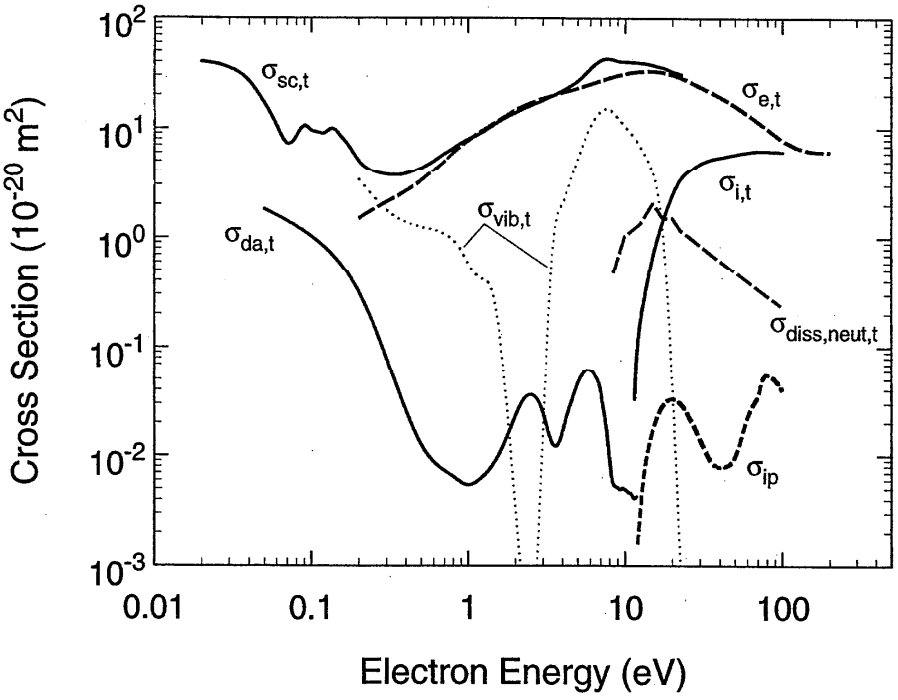


FIG. 25. Recommended and suggested cross sections for Cl_2.

TABLE 23. Ionization energy of Cl ($^2P_{3/2}$) for the production of Cl^+ ($^3P_{2,1,0}$), Cl^+ ($^1D_2$), and Cl^+ ($^1S_0$)

Ionic state	Ionization energy (eV)	References
$^3P_2$	12.97	147[a]
	12.97±0.02	146[a]
	12.967±0.001	148, 149[b]
$^3P_1$	13.06	147
	13.06±0.02	146
	13.053	148, 149
$^3P_0$	13.1	147
	13.090	148, 149
$^1D_2$	14.42	147
	14.41±0.02	146
	14.412	148, 149
$^1S_0$	16.42	147
	16.42±0.02	146
	16.423	148

[a]He I (584 Å) photoelectron spectra data.
[b]Spectroscopic data.

consists, in addition to the atomic line spectrum, of a large number of red-degraded bands extending from about 640 to 340 nm which were generally assigned to Cl_2^+.

9. Suggested Cross Sections and Coefficients for Cl_2

Due to the paucity of confirmed data, only the cross section for total scattering, $\sigma_{sc,t}(\varepsilon)$, (Table 9, Fig. 6) is considered "recommended" at this time. However, a significant amount of data exist which are "suggested" as the best data presently available. These include:

(i) $\sigma_{e,t}(\varepsilon)$ in Table 11 (Fig. 9);
(ii) $\sigma_{i,t}(\varepsilon)$ in Table 12 (Fig. 14);
(iii) $\sigma_{diss,neut,t}(\varepsilon)$ in Table 14 (Fig. 16);
(iv) $\sigma_{da,t}(\varepsilon)$ in Table 16 (Fig. 17); and
(v) $\sigma_{ip}(\varepsilon)$ in Table 22 (Fig. 23).

The cross sections that have been designated as "recommended" or "suggested" in this paper are plotted in Fig. 25. Also shown in Fig. 25 is the derived $\sigma_{vib,indir}(\varepsilon)$ (from Fig. 12) for which we do not provide tabulated data due to the potential for large uncertainties inherent in the derivation method used. It should be observed that the suggested values of $\sigma_{e,t}(\varepsilon)$ exceed those of $\sigma_{sc,t}(\varepsilon)$ near 2 eV. While this is physically impossible, the amount that $\sigma_{e,t}(\varepsilon)$ exceeds $\sigma_{sc,t}(\varepsilon)$ is less than the quoted uncertainties of the two measurements.

The cross section set shown in Fig. 25 is obviously not complete, and should not be used as such. Obvious deficiencies in the set are the lack of a momentum transfer cross section, and the limited energy range of the suggested values. The suggested data in the figure should serve as a basis for the formulation of any complete, self-consistent cross section set for use by modelers.

TABLE 24. Photoionization cross section, $\sigma_{pi,Cl}(\lambda)$, of the Cl atom (measurements of Samson *et al.* from Ref. 151)

Wavelength (nm)	$\sigma_{pi,Cl}(\lambda)$ (10^{-22} m^2)	Wavelength (nm)	$\sigma_{pi,Cl}(\lambda)$ (10^{-22} m^2)
15.8	1.29	47.5	20.2
17.5	1.32	50.0	25.8
20.0	1.30	52.5	32.2
22.5	1.19	55.0	35.7
25.0	1.02	57.5	38.0
27.5	0.90	60.0	39.4
30.0	0.94	62.5	40.6
32.5	1.40	65.0	41.6
35.0	2.50	67.5	42.4
37.5	4.60	70.0	43.0
40.0	7.50	72.5	43.4
42.5	11.0	75.5	43.6
45.0	15.3		

Also suggested are the

(i) rate constant for electron attachment $k_{a,t}(\langle\varepsilon\rangle)$ in Table 17 (Fig. 19);
(ii) density normalized ionization coefficient $\alpha/N(E/N)$ in Table 13 (Fig. 15);
(iii) density-reduced electron attachment coefficient $\eta/N(E/N)$ in Table 20 (Fig. 21); and
(iv) the effective ionization coefficient $(\alpha-\eta)/N(E/N)$ in Table 21 (Fig. 22).

10. Data Needs for Cl_2

Although cross sections have been suggested for total elastic, vibrational excitation, ionization, dissociation into neutrals, dissociative electron attachment, and ion-pair formation, there is a need to improve the accuracy and reliability of all these cross sections. There is a need as well for measurements of the cross sections for momentum transfer, dissociative ionization, vibrational excitation, and electronic excitation, for which no data exist at this time.

With the possible exception of the rate constant for dissociative electron attachment, and the ionization and effective ionization coefficients, there is a need for measurement of all other coefficients (electron drift velocity in pure Cl_2 and in mixtures with rare gases, electron attachment, and electron diffusion).

11. Electron Collision Data for Cl and Cl^+

11.1. Cl

Atomic chlorine is an open-shell atom with a ground-state configuration $1s^22s^22p^63s^23p^5(^2P_{3/2})$. Its electron affinity is well established. Of the 38 values listed by Christodoulides *et al.*,[74] those obtained using the photodetachment method are the most accurate. These are: 3.613±0.003 eV,[143] 3.610±0.002 eV,[144] and 3.616±0.003 eV.[145] A value of 3.613 eV is recommended. Studies of He I photoelectron

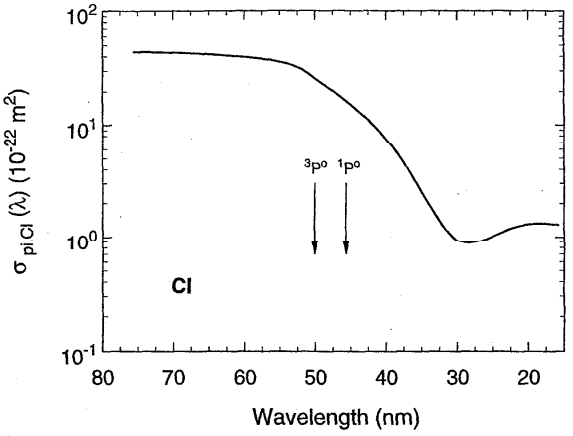


FIG. 26. Photoionization cross section as a function of wavelength, $\sigma_{pi,Cl}(\lambda)$, for atomic chlorine from the measurements of Samson *et al.* (Ref. 151). The vertical lines show the $^3P^0$ and $^1P^0$ limits.

spectra[146,147] of $Cl(^2P_{3/2})$ gave the ionization threshold energies listed in Table 23 for the production of Cl^+ in the ionic states $^3P_{2,1,0}$, $^1D_2$, and $^1S_0$.

De Lange *et al.*[150] used electron modulation spectroscopy and measured the photoionization cross section of Cl at the He I wavelength (584 Å) normalized to that for HCl and HBr at this wavelength. The cross section for ionization of the Cl atom into the ionic states Cl^+ ($^3P_{2,1,0}$), Cl^+ ($^1D_2$), and Cl^+ ($^1S_0$) were measured to be $(19.7\pm2.5)\times10^{-18}$ cm^2, $(11.4\pm1.5)\times10^{-18}$ cm^2, and $(2.16\pm0.28)\times10^{-18}$ cm^2, respectively. The absolute photoionization cross section as a function of photon wavelength, $\sigma_{pi,Cl}(\lambda)$, of the Cl atom was measured by Samson *et al.*[151] from 755 to 158 Å (16.4–75 eV) with an overall estimated uncertainty of ±8%. Their data are listed in Table 24 and are plotted in Fig. 26.

11.1.1. Total Electron Scattering Cross Section, $\sigma_{sc,t,Cl}(\varepsilon)$

There are no measurements or calculations of the total electron scattering cross section, $\sigma_{sc,t,Cl}(\varepsilon)$, of atomic chlorine. However, since we are dealing with an atomic species, below the threshold for electronic excitation of the chlorine atom at 8.90 eV, the total scattering cross section $\sigma_{sc,t,Cl}(\varepsilon)$ is equal to the total elastic electron scattering cross section $\sigma_{e,t,Cl}(\varepsilon)$. Above the ionization onset of the Cl atom at 12.97 eV, $\sigma_{sc,t,Cl}(\varepsilon)=\sigma_{e,t,Cl}(\varepsilon)+\sigma_{exc,t,Cl}(\varepsilon)+\sigma_{i,t,Cl}(\varepsilon)$, where $\sigma_{exc,t,Cl}(\varepsilon)$ and $\sigma_{i,t,Cl}(\varepsilon)$ are, respectively, the total cross sections for electronic excitation and electron-impact ionization of the Cl atom. Between 8.90 and 12.97 eV, $\sigma_{sc,t,Cl}(\varepsilon)=\sigma_{e,t,Cl}(\varepsilon)+\sigma_{exc,t,Cl}(\varepsilon)$. Thus, in principle, the cross section $\sigma_{sc,t,Cl}(\varepsilon)$ for the chlorine atom for the three energy regions mentioned above can be constructed using the expressions indicated for each energy region and data on $\sigma_{e,t,Cl}(\varepsilon)$, $\sigma_{exc,t,Cl}(\varepsilon)$, and $\sigma_{i,t,Cl}(\varepsilon)$. Unfortunately, this exercise is not feasible at the present time since, as will be seen later in this section, only the cross section for single ionization $\sigma_{i,Cl}(\varepsilon)$ is known with reasonable accuracy.

11.1.2. Momentum Transfer Cross Section, $\sigma_{m,Cl}(\varepsilon)$

There have been two calculations of the momentum transfer cross section, $\sigma_{m,Cl}(\varepsilon)$, of the Cl atom, the R matrix calculation of Griffin *et al.*[152] and the multiconfiguration Hartree-Fock calculation of Saha.[153] Figure 27 compares the results of these two calculations. Both calculations show the presence of a Ramsauer–Townsend minimum in $\sigma_{m,Cl}(\varepsilon)$ (at 0.95 eV,[153] at ~0.7 eV[152]). This minimum is similar to the well-known Ramsauer–Townsend minimum in the scattering cross section of the neighboring rare-gas Ar atom.

11.1.3. Total Elastic Electron Scattering Cross Section, $\sigma_{e,t,Cl}(\varepsilon)$

There are four calculations of the total elastic electron scattering cross section, $\sigma_{e,t,Cl}(\varepsilon)$, of the Cl atom,[152–155] but

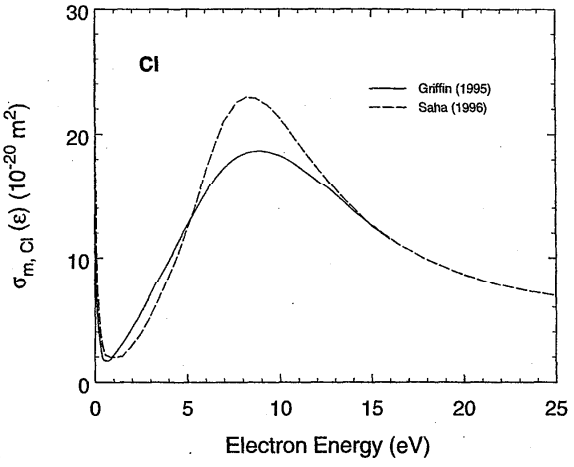


FIG. 27. Momentum transfer cross section, $\sigma_{m,Cl}(\varepsilon)$, for atomic chlorine: (—) R matrix calculation from Ref. 152; (- - -) multiconfiguration Hartree–Fock calculation from Ref. 153.

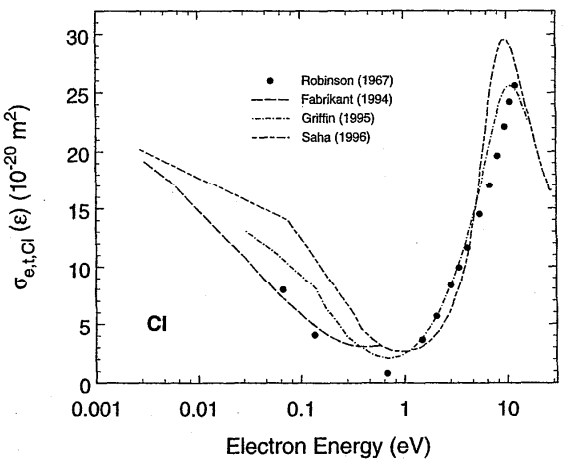


FIG. 28. Calculated total elastic electron scattering cross sections, $\sigma_{e,t,Cl}(\varepsilon)$, for atomic chlorine: (●) Ref. 154; (– –) Ref. 155; (–·–) Ref. 152; (— — —) Ref. 153.

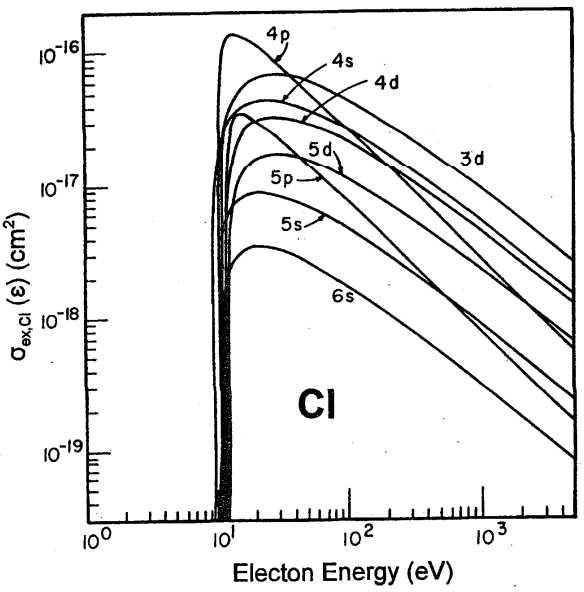


FIG. 29. Calculated cross sections for electron-impact excitation of the $4s$, $5s$, $6s$, $4p$, $5p$, $3d$, $4d$, and $5d$ states of the chlorine atom from the ground state $3p(^2P)$ from Ref. 156.

no measurements. These calculations are compared in Fig. 28. They all show the existence of a Ramsauer–Townsend minimum at ~0.7 eV,[154] at ~0.4 eV,[155] at 0.75 eV,[152] and at 0.95 eV.[153] Robinson and Geltman[154] performed a plane-wave calculation, Fabrikant[155] used the method of extrapolation of potential parameters along the isoelectronic sequence of positive ions to obtain scattering lengths for e-Cl scattering, Griffin *et al.*[152] used the R matrix method, and Saha[153] performed a multiconfiguration Hartree–Fock calculation. The agreement between these calculated cross sections is reasonable.

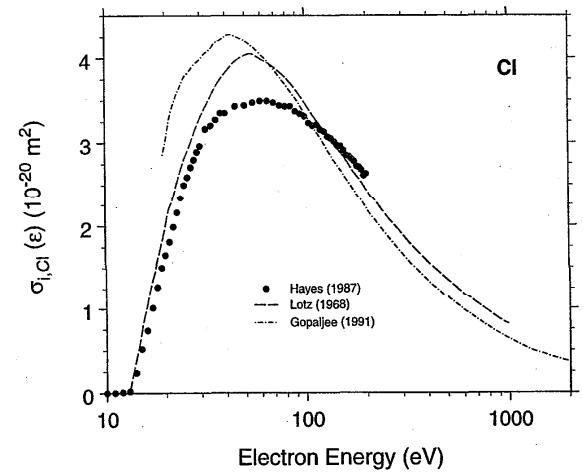


FIG. 30. Electron-impact single-ionization cross section, $\sigma_{i,Cl}(\varepsilon)$, for the Cl atom. (●) measurements from Ref. 157; (— - —) calculations from Ref. 158; (–·–) calculations from Ref. 159.

TABLE 25. Cross section, $\sigma_{i,Cl}(\varepsilon)$, for single ionization of Cl by electron impact (selected data of Hayes *et al.* from Ref. 157)

Electron energy (eV)	$\sigma_{i,Cl}(\varepsilon)$ (10^{-20} m^2)	Electron energy (eV)	$\sigma_{i,Cl}(\varepsilon)$ (10^{-20} m^2)
11	0.00	65	3.49
12	0.01	70	3.47
13	0.02	75	3.44
14	0.24	80	3.43
15	0.52	85	3.43
16	0.74	90	3.37
17	1.01	95	3.34
18	1.27	100	3.31
19	1.50	105	3.23
20	1.65	110	3.20
22	1.99	115	3.21
24	2.34	120	3.15
26	2.59	125	3.13
28	2.80	130	3.07
30	2.96	135	3.05
32	3.16	140	3.01
34	3.20	145	2.97
36	3.27	150	2.96
38	3.35	155	2.91
40	3.35	160	2.85
45	3.43	170	2.81
50	3.44	180	2.72
55	3.47	190	2.68
60	3.49	200	2.63

11.1.4. Electron-Impact Excitation Cross Section, $\sigma_{exc,Cl}(\varepsilon)$

Ganas[156] calculated cross sections, $\sigma_{exc,Cl}(\varepsilon)$, for electron-impact excitation of the $4s$, $5s$, $6s$, $4p$, $5p$, $3d$, $4d$, and $5d$ states of the chlorine atom from its ground state $3p(^2P)$. These are shown in Fig. 29. Similarly, Griffin *et al.*[152] calculated electron-impact excitation cross sections of Cl to the $3p^4 4s^4P_{5/2}$ level using the R matrix method, but the result of their calculation was found to depend on the number of states they considered. For this reason it is not considered here.

11.1.5. Electron-Impact Single-Ionization Cross Section, $\sigma_{i,Cl}(\varepsilon)$

Hayes *et al.*[157] measured the electron-impact single-ionization cross section, $\sigma_{i,Cl}(\varepsilon)$, of the Cl atom from the ionization threshold to 200 eV with an absolute uncertainty of ±14%. Their data are plotted in Fig. 30 and are listed in Table 25 as our suggested data since these are the only experimental measurements with a specified absolute uncertainty. Also shown in Fig. 30 are the $\sigma_{i,Cl}(\varepsilon)$ calculated by Lotz[158] and by Gopaljee *et al.*[159] Lotz calculated $\sigma_{i,Cl}(\varepsilon)$ using an empirical formula and estimated an error of +40%/−30%. Gopaljee *et al.*[159] used the binary encounter approximation. Not included in Fig. 30 are the distorted-wave calculation results of Griffin *et al.*[152] because they were found to vary considerably with the details of the calculation. Lennon *et al.*[160] also reviewed and recommended data for $\sigma_{i,Cl}(\varepsilon)$ and other positive ions of the Cl atom to 16+.

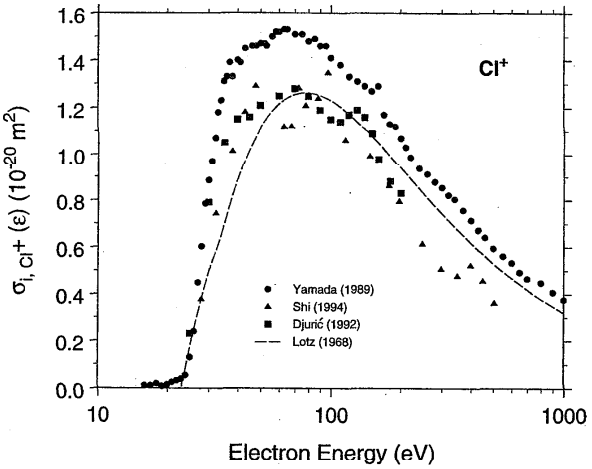


FIG. 31. Cross section, $\sigma_{i,Cl^+}(\varepsilon)$, for single ionization of Cl^+ by electron impact ($Cl^+ + e \rightarrow Cl^{++} + 2e$): (●) measurements from Ref. 165; (▲) measurements from Ref. 166; (■) measurements from Ref. 167; (- - -) semiempirical results using the Lotz formula from Refs. 158 and 165.

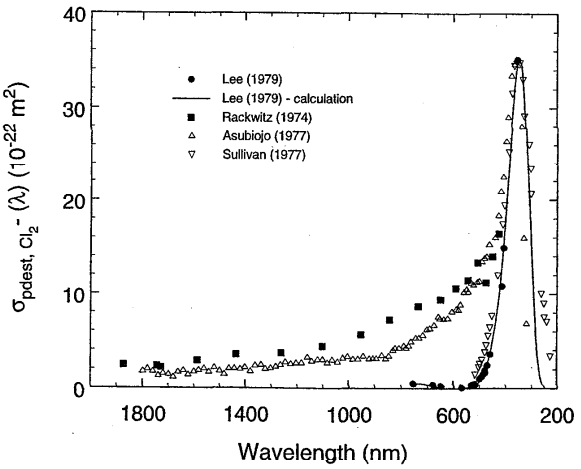


FIG. 32. Cross section for photodestruction of Cl_2^- as a function of photon wavelength, $\sigma_{pdest,Cl_2^-}(\lambda)$; (●) measurements of Lee *et al.* (Ref. 89); (—) calculation by Lee *et al.* (Ref. 89); (■) measurements of Rackwitz *et al.* (Ref. 169); (△) relative measurements of Asubiojo *et al.* (Ref. 170) normalized to the data of Lee *et al.* at 354 nm; (▽) relative measurements of Sullivan *et al.* (Ref. 171) normalized to the data of Lee *et al.* at 354 nm.

11.1.6. Radiative Attachment

Radiative attachment to the Cl atom, viz.,

$$Cl(^3P_{3/2,1/2}) + e \rightarrow Cl^-(^1S_0) + h\nu \qquad (6)$$

and the resulting radiative attachment continuum $h\nu$ (also known as the affinity spectrum) has long been investigated (e.g., see Refs. 144, 145, 161–163). The cross section for process (6) is expected to be very small.[163,164] In reaction (6) the photon energy consists of the electron affinity of the Cl atom and the kinetic energy of the attached electron. Because the kinetic energy of a free electron in, say, a plasma has a continuous range of values, the emission spectrum resulting from process (6) is continuous. From its long-wavelength limit (i.e., for the case where the kinetic energy of the captured electron is zero) the electron affinity (EA) of the Cl atom has been accurately determined. Thus, Pietsch and Rehder[144] obtained $\lambda(P_{3/2}) = (343.4 \pm 0.2)$ nm, corresponding to an EA for Cl ($P_{3/2}$) of (3.610 ± 0.002) eV, and $\lambda(P_{1/2}) = (333.1 \pm 0.4)$ nm, corresponding to an EA for Cl ($P_{1/2}$) of (3.722 ± 0.005) eV. Similarly, the radiative attachment continuum was found by Mück and Popp[145] to begin at 342.8 nm yielding an EA for Cl of 3.616 eV.

11.2. Cl^+

In Fig. 31 are shown the electron-impact ionization cross sections as a function of electron energy for Cl^+, $\sigma_{i,Cl^+}(\varepsilon)$, as measured in three crossed-beam experiments.[165–167] The results of Yamada *et al.*[165] extend from threshold to 1000 eV and have estimated total systematic errors of -8% to $+10\%$. The measurements of Shi *et al.*[166] cover the electron-impact energy range from 30 to 500 eV and have a reported uncertainty of $\pm 13\%$. Similarly, the data of Djurić *et al.*[167] stretch from threshold to 200 eV and have a systematic uncertainty of $\pm 10\%$. The measurements of Yamada *et al.* are consistently $\sim 25\%$ higher than the other two sets of measurements, possibly because of detector efficiency problems.[166,167] Consistent with the measurements of Djurić *et al.* and Shi *et al.* is the prediction of the semiempirical formula of Lotz[158] (see Fig. 31).

For electron-impact ionization cross section data on Cl^{++} see Mueller *et al.*[168] See also, the review by Lennon *et al.*[160] for ionization cross sections and ionization rate coefficients for multiply charged positive ions of Cl.

12. Electron Detachment, Electron Transfer, and Recombination and Diffusion Processes

12.1. Electron Detachment

The large cross section for dissociative electron attachment to the Cl_2 molecule makes the dissociative electron attachment process for chlorine an efficient mechanism to remove slow electrons in chlorine-containing plasma gases. Due to the depletion of free electrons, a higher electric field strength is required to increase the source of ionization[39] and sustain the ionization balance. In the active discharge, although electron detachment processes involve both Cl_2^- and Cl^-, those involving Cl^- are by far more significant in view of the larger abundance of Cl^- (see Sec. 6).

12.1.1. Photodestruction (Photodetachment and Photodissociation) of Cl_2^-

The interaction of light with Cl_2^- may result in either photodetachment

$$Cl_2^- + h\nu \rightarrow Cl_2 + e, \qquad (7)$$

or photodissociation

TABLE 26. Photodestruction cross section, $\sigma_{pdest,\,Cl_2^-}(\lambda)$, for Cl_2^- (data of Lee *et al.* from Ref. 89)

Wavelength (nm)	$\sigma_{pdest,\,Cl_2^-}(\lambda)$ (10^{-22} m^2)
350.7, 356.9	35.1 ±3.0
406.7	14.9 ±1.1
413.1	10.8 ±0.4
457.9	3.55±0.42
468.0	2.41±0.28
476.2	1.64±0.19
476.5	1.85±0.20
482.5	1.46±0.18
488.0	1.19±0.13
496.5	0.99±0.12
514.5	0.43±0.06
520.8	0.39±0.05
530.9	0.28±0.03
568.2	0.11±0.05
647.1	0.25±0.03
676.4	0.37±0.04
752.5	0.51±0.06

$$Cl_2^- + h\nu \rightarrow Cl + Cl^- \qquad (8)$$

of the Cl_2^- ion. These processes can be discussed and understood with reference to the potential energy curves shown in Fig. 5 for the ground state of Cl_2 ($^1\Sigma_g^+$) and Cl_2^- ($^2\Sigma_u^+$) and the excited states of Cl_2^- ($^2\Sigma_g^+$ and $^2\Pi_g$). Photodetachment from Cl_2^- ($^2\Sigma_u^+$, $v=0$) should be observed at a minimum energy corresponding to the EA of Cl_2 (Table 4). The cross section for photodetachment depends on the threshold law for photodetachment and the Franck–Condon factors which describe the overlap of the $v=0$ level of Cl_2^- ($^2\Sigma_u^+$) with the vibrational levels of the Cl_2 ($^1\Sigma_g^+$) ground state. Vibrational excitation in the molecular ion will also have an effect on the probability of photodetachment. Because of the large difference in the bond length of Cl_2^- and Cl_2 (Fig. 5, Tables 4 and 7), photodetachment will occur to high-lying vibrational levels of Cl_2 with low probability. Photodissociation is expected to result from excitation of Cl_2^- ($^2\Sigma_u^+$) into the repulsive excited states of Cl_2^-. The total photodestruction cross section is a combination of the two processes.

About 20 years ago processes (7) and (8) were the subject of a few investigations.[89,169–171] In Fig. 32 are shown the absolute measurements of Lee *et al.*[89] of the cross section, $\sigma_{pdest,Cl_2^-}(\lambda)$, for the photodestruction of the Cl_2^- ion. These were made over the wavelength range 350–760 nm using a drift-tube mass spectrometer–laser apparatus. The solid circles are their experimental measurements (listed in Table 26) and the solid curve is their calculated fit to their data. The strong peak in the photodestruction cross section was attributed[89] to the electronic transition $^2\Sigma_u^+ \rightarrow {}^2\Sigma_g^+$. In the experiments of Lee *et al.*, the Cl_2^- ion was probably produced via a three-body electron attachment process to Cl_2 and was converted to Cl_3^- in collisions with Cl_2. Also plotted in Fig. 32 are the absolute photodestruction cross section measurements of Rackwitz *et al.*[169] made in the photon energy range from 0.5 to 3.0 eV, and the relative photodestruction cross section of Asubiojo *et al.*[170] made between about 400 and 1800 nm and normalized to the data of Lee *et al.*[89] at 354 nm. The results of Rackwitz *et al.* suggest that the Cl_2^- ion formed by electron impact is vibrationally excited and this has a rather significant influence on the photodestruction cross section in the threshold region. This is supported by the work of Sullivan *et al.*[171] who examined photoinduced reactions of Cl_2^- in the gas phase using ion cyclotron resonance techniques. They found that the Cl_2^- ion undergoes photodissociation in preference to photodetachment and that the photodissociation spectrum of Cl_2^- exhibits one broad peak in the wavelength region from 220 to 700 nm with a maximum at (350± 10) nm which they attributed to the $^2\Sigma_u^+ \rightarrow {}^2\Sigma_g$ transition. This cross section has also been plotted in Fig. 32 after it has been normalized to the Lee *et al.* data at 354 nm.

It can thus be concluded[89] from the results of these four investigations that the Cl_2^- ion photodissociates rather than photodetaches, that the cross section for photodestruction depends on the electronic excitation of Cl_2^- upon photon impact, that the cross section threshold shifts to energies lower than the dissociation energy limit (1.26 eV, Table 7) of Cl_2^- into $Cl^- + Cl$ when the anion is vibrationally excited, and that the differences in the band widths between the four studies probably reflect differences in the vibrational temperature in the four experimental methods employed. The measurements of Lee *et al.*[89] with their quoted uncertainty are listed in Table 26 as our recommended values for $\sigma_{pdest,Cl_2^-}(\lambda)$.

12.1.2. Electron-Induced and Collisional Detachment of Cl_2^-

Apparently there are no data on electron-induced detachment, or collisional detachment involving the Cl_2^- ion.

12.1.3. Photodetachment of Cl^-

For Cl^- the most significant reactions and parameters are those involving the removal of the attached electron. These processes have been discussed by many authors (see for instance, Refs. 172 and 173). In this section we discuss briefly data on photodetachment of the Cl^- ion and in Sec. 12.1.4 data on collisional detachment of the Cl^- ion.

The Cl^- has a complete $3p^6$ subshell. Thus, the photodetachment process involves the removal of an electron from the p orbital and can be represented by

$$Cl^-(^1S_0) + h\nu \rightarrow Cl(^2P_{3/2,1/2}) + e. \qquad (9)$$

An early review of the experimental and theoretical data on the cross section, $\sigma_{pd,Cl^-}(\lambda)$, for reaction (9) was given by Popp.[163] In Fig. 33 are compared the experimental[145,161,162,174–177] and the calculated[154,178–180] data on $\sigma_{pd,Cl^-}(\lambda)$ for process (9). Most of these results were obtained over 20 years ago. The uncertainties in the experimental measurements are as follows: the single measurement of Berry *et al.*[174] at 336 nm (15×10^{-18} cm^2) has a quoted

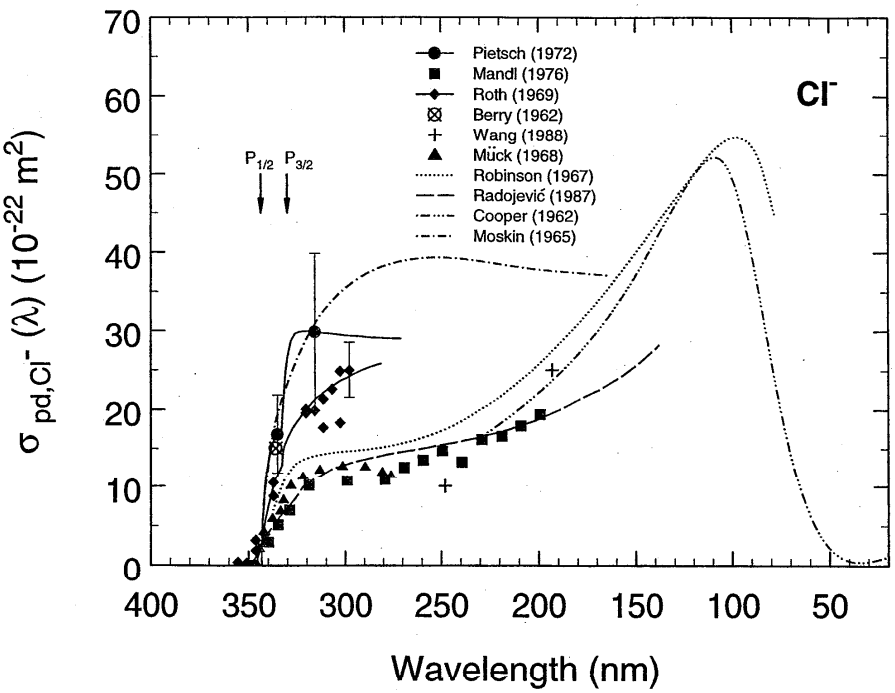


FIG. 33. Photodetachment cross section for Cl^-, $\sigma_{pd,Cl^-}(\lambda)$, as a function of photon wavelength, λ. Measurements: (●) Ref. 162; (■) Ref. 175; (◆) Ref. 161; (⊗) Ref. 174; (+) Ref. 176; (▲) Ref. 145. Calculations: (···) Ref. 154; (– –) Ref. 180; (–···–) Ref. 178; (–·–) Ref. 179. Typical error bars are shown in the figure for only the measurements of Refs. 161 and 162. See the text for the reported uncertainties of the other measurements. The energy positions of the $P_{1/2}$ and $P_{3/2}$ photodetachment thresholds are also shown.

uncertainty of $+12\times10^{-18}$ and $-5\times10^{-18}\,cm^2$; Mück and Popp's,[145] and Mandl's[175] uncertainties were quoted as ±25%; Roth's[161] and Pietsch's[162] uncertainties are as shown by the typical error bars in Fig. 33; Wang and Lee[176,177] reported a photodetachment cross section value for Cl^- equal to 2.5×10^{-17} and $1.0\times10^{-17}\,cm^2$ at 193 and 248 nm, respectively, but gave no uncertainty. On the calculation side, Robinson and Geltman[154] quoted an uncertainty of ±20%. It should be noted that the relativistic random-phase approximation result of Radojević *et al.*[180] extends to 100 eV and that Radojević *et al.* shifted their calculated curve from the theoretical threshold to the experimental value. It is seen from Fig. 33 that the spread in the experimental data is outside of the quoted uncertainties. The limited recent measurements of Wang and Lee[176] are consistent with the earlier measurements of Mandl,[175] and Mück and Popp,[145] but all three measurements are lower (often by a factor of 2 or more) than the data of Rothe,[161] Pietsch,[162] and Berry *et al.*[174] On the theoretical side, the calculated values of $\sigma_{pd,Cl^-}(\lambda)$ by Moskvin[179] differ substantially from the results of the other three calculations.[154,178,180]

12.1.4. Collisional Detachment of Cl^-

Collisional detachment reactions fall into three groups:[172,173] direct detachment, detachment with excitation (of autodetaching levels, or of a neutral product, or via charge transfer to a negative ion state of the target), and detachment with bonding (reactive collision with detachment, or associative detachment). The magnitude and the

TABLE 27. Associative detachment thermal rate constants involving Cl^-

Reactants	Temperature (K)	Associative detachment thermal rate constant (cm^3 $molecule^{-1}$ s^{-1})	Reference
$Cl^-+H\rightarrow HCl+e$	296	9.6×10^{-10}	181
	296	10.0×10^{-10}	182
	296	9.0×10^{-10}	183
$Cl^-+O\rightarrow ClO+e$	300	$<1\times10^{-11}$	184
$Cl^-+N\rightarrow ClN+e$	300	$<1\times10^{-11}$	184
$Cl^-+Cl_2(+He)\rightarrow Cl_3^-$	ambient	0.9×10^{-29} (cm^6 $molecule^{-2}$ s^{-1})	185

TABLE 28. Energy threshold for the detachment of Cl^- in collisions with various target gases as reported by Doverspike *et al.* in Ref. 186

Reactants	Threshold energy (eV)
Cl^-+H_2	5.5±0.1
Cl^-+D_2	5.5±0.1
Cl^-+N_2	7.6±0.1
Cl^-+O_2[a]	4.4±0.2
Cl^-+CO	7.1±0.2
Cl^-+CO_2	7.3±0.2
Cl^-+CH_4	6.2±0.2

[a]In addition to direct detachment there are several other processes which may contribute to the products of this reaction at energies below 4.4 eV, such as the charge-transfer reaction $Cl^-+O_2 \rightarrow Cl+O_2^-$ which is endothermic by ~3.1 eV and the associative detachment reaction $Cl^-+O_2 \rightarrow ClO_2+e$ which is endothermic by 3.4 eV (see Ref. 186).

dependence of the collisional detachment cross section, $\sigma_{cd}(\mathcal{E})$, on the energy, $\mathcal{E}$, of the reactants varies with the type of the detachment process. (Note that $\mathcal{E}$ refers to the energy of reactants, i.e., the projectile ion and the neutral target.) Thus, the rising parts of $\sigma_{cd}(\mathcal{E})$ as the kinetic energy of the reactants increases are principally due to direct collisional detachment, while the rising parts of $\sigma_{cd}(\mathcal{E})$ as the kinetic energy of the reactants decreases toward thermal energy are due to associative detachment. Generally, there is a threshold for the direct collisional detachment process which occurs (when the reactants are in their ground states) when their kinetic energy is equal to the EA of the species carrying the extra electron, although in certain cases such as for the reactions Cl^-+M (where M is a molecule), the $\sigma_{cd}(\mathcal{E})$ increases rapidly from the threshold which itself is considerably greater than the EA of the Cl atom. The associative detachment process besides being responsible for the large cross sections at thermal and near-thermal energies also accounts for maxima often seen in the $\sigma_{cd}(\mathcal{E})$ functions at higher energies due to negative ion resonances. In Table 27 are listed values of the thermal ($T \approx 300$ K) rate constants for the associative detachment reactions

$$Cl^-+X \rightarrow ClX+e, \tag{10}$$

where X=H, O, N, or Cl_2.

The threshold for collisional detachment can be low, and the cross section for collisional detachment can be very large[172,173]—indeed, in many cases, much larger than the cross section for photodetachment. When the associative detachment reactions (10) are exothermic, that is, when the so-called energy defect (the energy difference between the dissociation energy of ClX and the EA of Cl) is positive, and the reactions are not hindered by geometric or other factors, the thermal values of the rate constants are large ($\sim 10^{-9}$ cm^3 s^{-1}) and close to the values of the orbiting Langevin collision rate constants. Collisional detachment, then, especially when it is field assisted, can be a dominant electron release mechanism in electrically stressed gases.

Doverspike *et al.*[186] measured absolute total electron detachment cross sections for collisions of Cl^- with a number of molecular targets X (X=H_2, D_2, O_2, N_2, CO, CO_2, and CH_4) for collision energies below the threshold for detachment to several hundred eV. The reaction studied is

$$Cl^-+X \rightarrow Cl+X+e. \tag{11}$$

In all such collisions the detachment thresholds were found to exceed the electron affinity of the Cl atom. Table 28 lists the threshold values for collisional detachment as reported by Doverspike *et al.*[186] The results of Doverspike *et al.* are shown in Fig. 34(a) for energies near threshold and in Fig. 34(b) for higher energies.

Huq *et al.*[188] measured absolute total cross sections for charge transfer and electron detachment of Cl^- on Cl_2. In Fig. 35 are shown their measurements of the total cross sections for electron detachment and for "slow" ion production (via charge transfer). The quoted uncertainty is about ±10%. In Fig. 35 are also shown the earlier measurements by Hasted and Smith[189] who reported cross sections for electron detachment in collisions of Cl^- with Cl_2 in the energy range from 10 to 2500 eV. According to Huq *et al.*,[188] it appears that, at the lowest energies, the Hasted and Smith study did not fully resolve ions from electrons.

Measurements of the translational energy thresholds for electron transfer reactions for various atomic negative ions to Cl_2 at room temperature[121,122] allowed determination of the electron affinity of the Cl_2 molecule. Thus, from measurements of the energy thresholds for the endothermic electron transfer reactions I^-+Cl_2 and Cl^-+Cl_2, Hughes *et al.*[122] obtained a value of (2.62±0.2) eV for the EA of Cl_2. Similarly, from the room temperature relative cross sections for the reactions of I^-, Br^-, and Cl^- with Cl_2, Chupka *et al.*[121] obtained for the EA of Cl_2 the value of 2.38±0.10 eV.

12.2. Electron Transfer

While the reaction

$$Cl_2^-+Cl_2 \rightarrow Cl_3^-+Cl \tag{12}$$

is endoergic when the reactants are thermalized,[89] Hughes *et al.*[122] found that it becomes exoergic at energies in excess of 0.3 eV with a rate constant at this energy equal to 0.0084×10^{-10} cm^3 molecule^{-1} s^{-1}. Similarly, the reaction

$$Cl_2+Cl^- \rightarrow Cl_3^- \tag{13}$$

was found by Babcock and Streit[185] to have a three-body rate constant (with He as the third body) of 0.9×10^{-29} cm^6 molecule^{-2} s^{-1}.

Measurements of the translational energy thresholds for electron-transfer reactions for various atomic negative ions (e.g., I^- and Cl^-) to Cl_2 allowed determination of the electron affinity of the chlorine molecule. Thus, Hughes *et al.*[122]

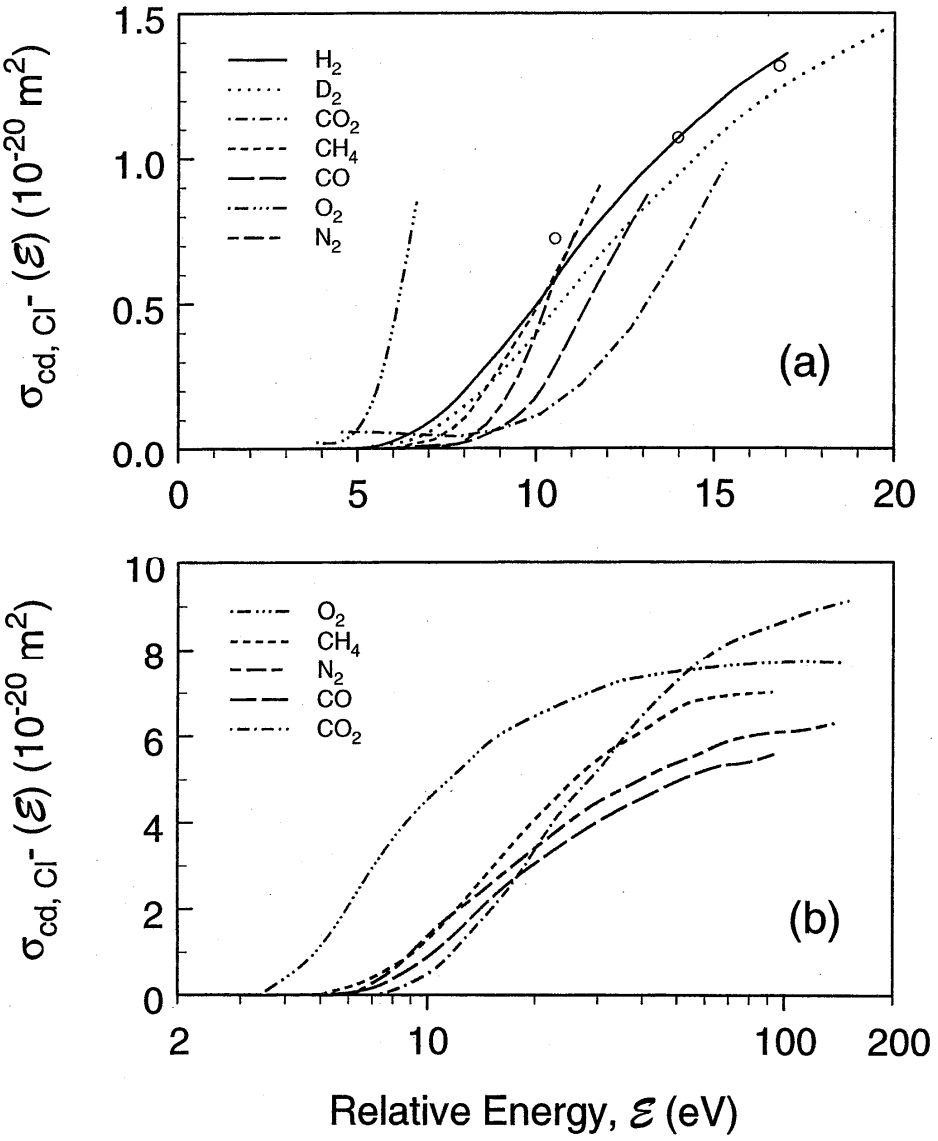


FIG. 34. Collisional detachment cross sections, $\sigma_{cd,Cl^-}(\mathcal{E})$, as a function of the relative energy of the reactants, $\mathcal{E}$, involving Cl^- and various molecular targets (a) near threshold energies (b) over a wider energy range. All data are from Dovespike *et al.* (Ref. 186) except for the three data points (O) for the Cl^-+H_2 reaction which are those of Bydin and Dukel'skii (Ref. 187).

and Chupka *et al.*[121] determined via such reactions the electron affinity of the Cl_2 molecule to be, respectively, (2.32 $\pm$0.1) and (2.38$\pm$0.1) eV.

12.3. Recombination and Diffusion Processes

12.3.1. Recombination of Cl_2^+ and Cl^-

Positive ion–negative ion recombination measurements in flowing afterglow plasmas by Church and Smith[190] gave the value of 5.0×10^{-8} cm^3 molecule^{-1} s^{-1} for the rate constant of the reaction $Cl_2^+ + Cl^- \rightarrow$ products.

12.3.2. Recombination of Cl

Boyd and Burns[191] compared recombination and dissociation rate constants for halogens obtained by a variety of experimental techniques. The Cl–Cl recombination is exothermic ($\Delta H \sim -1.1$ eV) and requires a third body, M, i.e.,

$$Cl + Cl + M \rightarrow Cl_2 + M. \qquad (14)$$

Boyd and Burns observed that the three-body recombination rate constant for reaction (14) decreases with increasing temperature and that the Cl_2 molecules are not efficient third bodies at any temperature. Measurements of atomic chlorine concentration in Cl_2 plasmas using infrared absorption spectroscopy by Richards and Sawin[2] showed that gas-phase re-

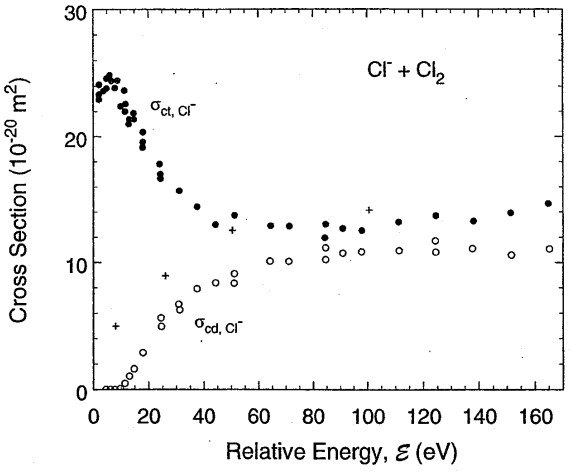


FIG. 35. Cross section, $\sigma_{ct,Cl^-}(\mathcal{E})$, for charge transfer as a function of the relative energy of the reactants, $\mathcal{E}$, in collisions of Cl^- with Cl_2: (●) data of Huq *et al.* from Ref. 188. For comparison the cross section $\sigma_{cd,Cl^-}(\mathcal{E})$ is also shown: (○) data of Huq *et al.* from Ref. 188; (×) data of Hasted and Smith from Ref. 189.

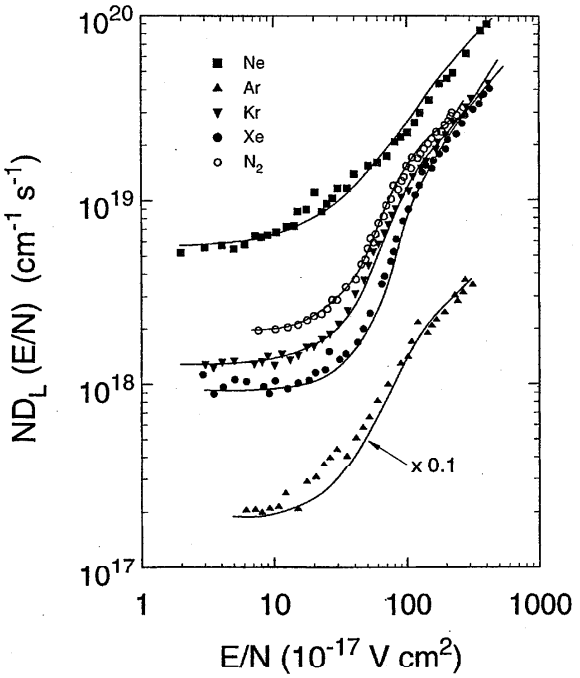


FIG. 36. The product, $D_L N(E/N)$, of the longitudinal diffusion coefficient D_L and the neutral gas number density N as a function of E/N for Cl^- in Ne, Ar, Kr, and Xe (data from Ref. 194). Also shown are measurements made in N_2 from Ref. 198. The points are experimental data and the curves are calculated results using the generalized Einstein relation between D_L and K (see Refs. 194–197). Note that the Ar data have been multiplied by 0.1 for convenience of display.

combination is an insignificant Cl loss mechanism. For the temperature of their experiment (770 K), the rate constant for reaction (14) (M=Cl_2) is $\approx 2.8\times10^{-32}$ cm^6 $molecule^{-2}\,s^{-1}$.[191] Richards and Sawin thus concluded that the major mechanism for Cl loss is likely to be a recombination on the electrode surfaces.

12.3.3. Diffusion of Cl and Cl^- in Gases

Chang *et al.*[192] measured the diffusion coefficient of atomic chlorine in molecular chlorine. They reported a value for the diffusion coefficient of chlorine atoms in chlorine molecules of $(0.149\pm0.025)\ cm^2\,s^{-1}$ at 298 K and 1 atm. Similarly, Hwang *et al.*[193] measured the diffusion coefficients of atomic chlorine in rare gases via radiative recombination reactions. At 296 K and 101.33 kPa (1 atm) of rare-gas pressure, the values of the diffusion constant for Cl in He, Ne, Ar, Kr, and Xe were measured to be, respectively, $(0.75\pm0.12)\ cm^2\,s^{-1}$, $(0.32\pm0.05)\ cm^2\,s^{-1}$, $(0.19\pm0.03)\ cm^2\,s^{-1}$, $(0.14\pm0.02)\ cm^2\,s^{-1}$, and $(0.12\pm0.02)\ cm^2\,s^{-1}$.

Eisele *et al.*[194] measured the longitudinal diffusion coefficients for Cl^- ions in Ne, Ar, Kr, and Xe as a function of E/N. Measurements were made at about 300 K and at gas pressures below 0.067 kPa. They are shown in Fig. 36. As $E/N\rightarrow0$, the ions are in thermal equilibrium with the gas molecules and the diffusion coefficient is isotropic, related to the ionic mobility K by the relation $K=eD/kT$, where e is the ionic charge, k is Boltzmann's constant, and T is the gas temperature. For larger values of E/N, this relation is not valid because the diffusion coefficient has components that refer to the directions parallel and perpendicular to the electric field (for computational techniques allowing the calculation of the diffusion coefficient at any value of E/N from knowledge of the ionic mobility at that E/N see Refs. 195–197).

Finally, Thackston *et al.*[198] reported measurements of the longitudinal diffusion coefficients for Cl^- in N_2. These are also shown in Fig. 36 with an uncertainty of ±7% at all E/N.

13. Summary for Other Species and Processes

With the exception of the limited measurements on electron-impact ionization of Cl and Cl^+, no measurements are known to have been made for other electron collision processes for the species Cl, Cl^-, Cl^+, and Cl_2^-. With regard to data on other important processes in Cl_2 plasmas, data have been summarized in this paper on the photodetachment of Cl^-, charge transfer reactions involving Cl^- and various molecular partners, and diffusion coefficients for Cl^- in rare gases and N_2. Much work is needed on electron collision and other processes involving the main species in Cl_2 plasmas.

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15. References

[1] G. L. Rogoff, J. M. Kramer, and R. B. Piejak, IEEE Trans. Plasma. Sci. **PS-14**, 103 (1986).
[2] A. D. Richards and H. H. Sawin, J. Appl. Phys. **62**, 799 (1987).
[3] D. S. Fischl and D. W. Hess, J. Vac. Sci. Technol. B **6**, 1577 (1988).
[4] R. d'Agostino, F. Cramarossa, F. Fracassi, F. Illuzzi, and M. N. Armenise, J. Vac. Sci. Technol. B **6**, 1584 (1988).
[5] *Plasma Etching*, edited by D. M. Manos and D. L. Flamm (Academic, Boston, 1989).
[6] L. E. Kline and M. J. Kushner, Crit. Rev. Solid State Mater. Sci. **16**, 1 (1989).
[7] T. Matsuura, H. Uetake, T. Ohmi, J. Murota, K. Fukuda, N. Mikoshiba, T. Kawashima, and Y. Yamashita, Appl. Phys. Lett. **56**, 1339 (1990).
[8] V. V. Boiko, A. T. Rakhimov, and N. V. Suetin, Sov. Phys. Tech. Phys. **35**, 1268 (1990).
[9] L. Peters, Semicond. Int., May, 66 (1992).
[10] E. S. Aydil and D. J. Economou, J. Electrochem. Soc. **139**, 1396 (1992).
[11] N. L. Bassett and D. J. Economou, J. Appl. Phys. **75**, 1931 (1994).
[12] S. C. Deshmukh and D. J. Economou, J. Appl. Phys. **72**, 4597 (1992).
[13] P. L. G. Ventzek, M. Grapperhaus, and M. J. Kushner, J. Vac. Sci. Technol. B **12**, 3118 (1994).
[14] C. C. Cheng, K. V. Guinn, V. M. Donnelly, and I. P. Herman, J. Vac. Sci. Technol. A **12**, 2630 (1994).
[15] D. P. Lymberopoulos and D. J. Economou, IEEE Trans. Plasma Science **23**, 573 (1995).
[16] C. Lee and M. A. Lieberman, J. Vac. Sci. Technol. A **13**, 368 (1995).
[17] M. J. Kushner, J. Appl. Phys. **82**, 5312 (1997).
[18] Y. T. Lee, M. A. Lieberman, A. J. Lichtenberg, F. Bose, H. Baltes, and R. Patrick, J. Vac. Sci. Technol. A **15**, 113 (1997).
[19] G. I. Font and I. D. Boyd, J. Vac. Sci. Technol. A **15**, 313 (1997).
[20] J. P. Chang, A. P. Mahorowala, and H. H. Sawin, J. Vac. Sci. Technol. A **16**, 217 (1998).
[21] H. Doshita, K. Ohtani, and A. Namiki, J. Vac. Sci. Technol. A **16**, 265 (1998).
[22] G. P. Kota, J. W. Coburn, and D. B. Graves, J. Vac. Sci. Technol. A **16**, 270 (1998).
[23] M. V. Malyshev, V. M. Donnelly, A. Kornblit, and N. A. Ciampa, J. Appl. Phys. **84**, 137 (1998).
[24] M. V. Malyshev, V. M. Donnelly, and S. Samukawa, J. Appl. Phys. **84**, 1222 (1998).
[25] A. K. Hays, Opt. Commun. **28**, 209 (1979).
[26] M. C. Castex, J. Le Calvé, D. Haaks, B. Jordan, and G. Zimmerer, Chem. Phys. Lett. **70**, 106 (1980).
[27] M. Rokni and J. H. Jacob, in *Applied Atomic Collision Physics*, edited by H. S. W. Massey, E. W. McDaniel, and B. Bederson, Gas Lasers, Vol. 3 (Academic, New York, 1982), p. 273.
[28] W. L. Nighan, in *Applied Atomic Collision Physics*, edited by H. S. W. Massey, E. W. McDaniel, and B. Bederson, Gas Lasers, Vol. 3 (Academic, New York, 1982), p. 319.
[29] M. R. Flannery, in *Applied Atomic Collision Physics*, edited by H. S. W. Massey, E. W. McDaniel, and B. Bederson, Gas Lasers, Vol. 3 (Academic, New York, 1982), p. 141.
[30] D. R. Bates, Adv. At. Mol. Phys. **20**, 1 (1985).
[31] W. L. Morgan, Plasma Chem. Plasma Process. **12**, 449 (1992).
[32] T. E. Graedel and P. J. Crutzen, *Chemie der Atmosphäre* (Spektrum Akademischer, Heidelberg, 1994).
[33] D. Maric, J. P. Burrows, R. Meller, and G. K. Moortgat, J. Photochem. Photobiol. A: Chem. **70**, 205 (1993).
[34] L. G. Christophorou, J. K. Olthoff, and M. V. V. S. Rao, J. Phys. Chem. Ref. Data **25**, 1341 (1996).
[35] L. G. Christophorou, J. K. Olthoff, and M. V. V. S. Rao, J. Phys. Chem. Ref. Data **26**, 1 (1997).
[36] L. G. Christophorou, J. K. Olthoff, and Y. Wang, J. Phys. Chem. Ref. Data **26**, 1205 (1997).
[37] L. G. Christophorou and J. K. Olthoff, J. Phys. Chem. Ref. Data **27**, 1 (1998).
[38] L. G. Christophorou and J. K. Olthoff, J. Phys. Chem. Ref. Data **27**, 889 (1998).
[39] R. Nagpal and A. Garscaddden, Contrib. Plasma Phys. **35**, 301 (1995).
[40] N. Pinhão and A. Chouki, Proceedings XXII International Conference on Phenomena in Ionized Gases, Hoboken, NJ, July 31–August 4, 1995, edited by K. H. Becker, W. E. Carr, and E. E. Kunhardt, Contributed Paper 2, p. 5.
[41] A. M. Efremov, V. I. Svettsov, and V. P. Mikhalkin, High Energy Chem. **29**, 433 (1995).
[42] G. Herzberg, *Molecular Spectra and Molecular Structure I. Spectra of Diatomic Molecules*, 2nd ed. (Van Nostrand Reinhold, New York, 1950).
[43] J. Jureta, S. Cvejanović, M. Kurepa, and D. Cvejanović, Z. Phys. A **304**, 143 (1982).
[44] D. Spence, R. H. Huebner, H. Tanaka, M. A. Dillon, and R.-G. Wang, J. Chem. Phys. **80**, 2989 (1984).
[45] K. P. Huber and G. Herzberg, *Molecular Spectra and Molecular Structure. IV. Constants of Diatomic Molecules* (Van Nostrand Reinhold, New York, 1979).
[46] S. D. Peyerimhoff and R. J. Buenker, Chem. Phys. **57**, 279 (1981).
[47] H. von Halban and K. Siedentopf, Z. Phys. Chem. **103**, 71 (1922a).
[48] H. von Halban and K. Siedentopf, Z. Electrochem. **28**, 496 (1922b).
[49] G. E. Gibson and N. S. Bayliss, Phys. Rev. **44**, 188 (1933).
[50] F. W. Jones and W. Spooner, Trans. Faraday Soc. **31**, 811 (1935).
[51] W. C. Fergusson, L. Slotkin, and W. G. Style, Trans. Faraday Soc. **32**, 956 (1936).
[52] R. G. Aickin and N. S. Bayliss, Trans. Faraday Soc. **33**, 1333 (1937).
[53] K. Watanabe, J. Chem. Phys. **26**, 542 (1957).
[54] J. Lee and A. D. Walsh, Trans. Faraday Soc. **55**, 1281 (1959).
[55] R. P. Iczkowski, J. L. Margrave, and J. W. Green, J. Chem. Phys. **33**, 1261 (1960).
[56] A. E. Douglas, C. K. Moeller, and B. P. Stoicheff, Can. J. Phys. **41**, 1174 (1963).
[57] D. J. Seery and D. Britton, J. Phys. Chem. **68**, 2263 (1964).
[58] H. Okabe, *Photochemistry of Small Molecules* (Wiley-Interscience, New York, 1978).
[59] C. Roxlo and A. Mandl, J. Appl. Phys. **51**, 2969 (1980).
[60] J. A. Coxon, J. Mol. Spectrosc. **82**, 264 (1980).
[61] A. E. Douglas, Can. J. Phys. **59**, 835 (1981).
[62] J. B. Burkholder and E. J. Bair, J. Phys. Chem. **87**, 1859 (1983).
[63] J. A. R. Samson and G. C. Angel, J. Chem. Phys. **86**, 1814 (1987).
[64] J. W. Gallagher, C. E. Brion, J. A. R. Samson, and P. W. Langhoff, J. Phys. Chem. Ref. Data **17**, 9 (1988).
[65] J. A. Ganske, H. N. Berko, and B. J. Finlayson-Pitts, J. Geophys. Res. **97**, 7651 (1992).
[66] S. Hubinger and J. B. Nee, J. Photochem. Photobiol., A **86**, 1 (1995).
[67] D. C. Frost, C. A. McDowell, and D. A. Vroom, J. Chem. Phys. **46**, 4255 (1967).
[68] A. W. Potts and W. C. Price, J. Chem. Soc. Faraday Trans. 2 **67**, 1242 (1971).
[69] A. B. Cornford, D. C. Frost, C. A. McDowell, J. L. Ragle, and I. A. Stenhouse, J. Chem. Phys. **54**, 2651 (1971).
[70] J. H. D. Eland, in *Electron Spectroscopy: Theory, Techniques, and Applications*, edited by C. R. Brundle and A. D. Baker (Academic, New York, 1979), Vol. 3, Chap. 5.
[71] T. A. Carlson, M. O. Krause, F. A. Grimm, and T. A. Whitley, J. Chem. Phys. **78**, 638 (1983).
[72] H. Van Lonkhuyzen and C. A. de Lange, Chem. Phys. **89**, 313 (1984).
[73] A. G. McConkey, G. Dawber, L. Avaldi, M. A. MacDonald, G. C. King, and R. I. Hall, J. Phys. B **27**, 271 (1994).
[74] A. A. Christodoulides, D. L. McCorkle, and L. G. Christophorou, in *Electron Molecule Interactions and Their Applications*, edited by L. G. Christophorou (Academic, New York, 1984), Vol. 2, Chap. 6.
[75] S. G. Lias, J. E. Bartmess, J. F. Liebman, J. L. Holmes, R. D. Levin, and W. G. Mallard, J. Phys. Chem. Ref. Data **17**, 592 (1988).
[76] D. C. Frost and C. A. McDowell, Can. J. Chem. **38**, 407 (1960).
[77] L. Frost, A. M. Grisogono, I. E. McCarthy, E. Weigold, C. E. Brion, A. O. Bawagan, P. K. Mukherjee, W. Von Niessen, M. Rosi, and A. Sgamellotti, Chem. Phys. **113**, 1 (1987).
[78] J. D. Morrison and A. J. C. Nicholson, J. Chem. Phys. **20**, 1021 (1952).
[79] R. Thorburn, Proc. Phys. Soc. (London) **123**, 122 (1959).
[80] M. V. Kurepa and D. S. Belić, Chem. Phys. Lett. **49**, 608 (1977).
[81] V. E. Bondybey and C. Fletcher, J. Chem. Phys. **64**, 3615 (1976).
[82] T. Moeller, B. Jordan, P. Gürtler, G. Zimmerer, D. Haaks, J. Le Calve, and M.-C. Castex, Chem. Phys. **76**, 295 (1983).
[83] R. G. McLoughlin, J. D. Morrison, and D. L. Smith, Int. J. Mass Spectrom. Ion Proc. **58**, 201 (1984).

[84] R. J. Stubbs, T. A. York, and J. Comer, J. Phys. B **18**, 3229 (1985).
[85] T. L. Gilbert and A. C. Wahl, J. Chem. Phys. **55**, 5247 (1971).
[86] P. W. Tasker, G. G. Balint-Kurti, and R. N. Dixon, Mol. Phys. **32**, 1651 (1976).
[87] W.-C. Tam and S. F. Wong, J. Chem. Phys. **68**, 5626 (1978).
[88] M. Rokni, J. H. Jacob, and J. A. Mangano, Appl. Phys. Lett. **34**, 187 (1979).
[89] L. C. Lee, G. P. Smith, J. T. Moseley, P. C. Cosby, and J. A. Guest, J. Chem. Phys. **70**, 3237 (1979).
[90] J. G. Dojahn, E. C. M. Chen, and W. E. Wentworth, J. Phys. Chem. **100**, 9649 (1996).
[91] J. B. Fisk, Phys. Rev. **51**, 25 (1937).
[92] R. J. Gulley, T. A. Field, W. A. Steer, N. J. Mason, S. L. Lunt, J.-P. Ziesel, and D. Field, J. Phys. B **31**, 2971 (1998).
[93] G. D. Cooper, J. E. Sanabia, J. H. Moore, J. K. Olthoff, and L. G. Christophorou, J. Chem. Phys. **110**, 682 (1999).
[94] M. Gote and H. Ehrhardt, J. Phys. B **28**, 3957 (1995).
[95] M. V. Kurepa and D. S. Belić, J. Phys. B **11**, 3719 (1978).
[96] R. Azria, R. Abouaf, and D. Teillet-Billy, J. Phys. B **15**, L569 (1982).
[97] D. Spence, Phys. Rev. A **10**, 1045 (1974).
[98] H. Kutz and H.-D. Meyer, Phys. Rev. A **51**, 3819 (1995).
[99] A. Ernesti, M. Gote, and H. J. Kotsch, Phys. Rev. A **52**, 1266 (1995).
[100] T. N. Rescigno, Phys. Rev. A **50**, 1382 (1994).
[101] V. A. Bailey and R. H. Healey, Philos. Mag. **19**, 725 (1935).
[102] S. E. Božin and C. C. Goodyear, Br. J. Appl. Phys. **18**, 49 (1967).
[103] M.-C. Bordage, P. Ségur, and A. Chouki, J. Appl. Phys. **80**, 1325 (1996).
[104] M.-C. Bordage, P. Ségur, L. G. Christophorou, and J. K. Olthoff, J. Appl. Phys. (submitted).
[105] L. G. Christophorou, *Atomic and Molecular Radiation Physics* (Wiley-Interscience, New York, 1971), p. 328.
[106] H. Ehrhardt, L. Langhans, F. Linder, and H. S. Taylor, Phys. Rev. **173**, 222 (1968).
[107] R. E. Center and A. Mandl, J. Chem. Phys. **57**, 4104 (1972).
[108] F. A. Stevie and M. J. Vasile, J. Chem. Phys. **74**, 5106 (1981).
[109] S. K. Srivastava and R. Boivin, Bull. Am. Phys. Soc. **42**, 1738 (1997); S. K. Srivastava (private communication, January 1998).
[110] Y.-K. Kim (private communication, January 1998).
[111] H. Deutsch, K. Becker, and T. D. Märk (private communication, February 1998).
[112] D. Rapp and P. Englander-Golden, J. Chem. Phys. **43**, 1464 (1965).
[113] P. C. Cosby, Bull. Am. Phys. Soc. **35**, 1822 (1990).
[114] P. C. Cosby and H. Helm, Wright Laboratory Report No. WL-TR-93-2004, Wright Patterson AFB, OH 45433-7650 (1992); (private communication, June 1998).
[115] W. C. Wells and E. C. Zipf, J. Chem. Phys. **66**, 5828 (1977).
[116] J. B. A. Mitchell, Phys. Rep. **186**, 216 (1990).
[117] D. L. McCorkle, A. A. Christodoulides, and L. G. Christophorou, Chem. Phys. Lett. **109**, 276 (1984).
[118] R. Azria, L. Parenteau, and L. Sanche, J. Chem. Phys. **87**, 2292 (1987).
[119] L. G. Christophorou, in *Linking the Gaseous and the Condensed Phases of Matter, the Behavior of Slow Electrons*, edited by L. G. Christophorou, E. Illenberger, and W. F. Schmidt (Plenum, New York, 1994), p. 3.
[120] D. Muigg, G. Denifl, A. Stamatović, E. Illenberger, I. Walker, and T. D. Märk, Chem. Phys. Lett. (submitted); (private communication, October 1998).
[121] W. A. Chupka, J. Berkowitz, and D. Gutman, J. Chem. Phys. **55**, 2724 (1971).
[122] B. M. Hughes, C. Lifshitz, and T. O. Tiernan, J. Chem. Phys. **59**, 3162 (1973).
[123] L. G. Christophorou, Z. Phys. Chem. (Munich) **195**, 195 (1996).
[124] R. C. Sze, A. E. Greene, and C. A. Brau, J. Appl. Phys. **53**, 1312 (1982).
[125] N. E. Bradbury, J. Chem. Phys. **2**, 827 (1934).
[126] D. Smith, N. G. Adams, and E. Alge, J. Phys. B **17**, 461 (1984).
[127] P. J. Chantry, in *Applied Atomic Collision Physics*, edited by H. S. W. Massey, E. W. McDaniel, and B. Bederson, Gas Lasers, Vol. 3 (Academic, New York, 1982), p. 35.
[128] M. V. Kurepa, D. S. Babić, and D. S. Balić, Chem. Phys. **59**, 125 (1981).
[129] J. A. Ayala, W. E. Wentworth, and E. C. M. Chen, J. Phys. Chem. **85**, 768 (1981).
[130] A. A. Christodoulides, R. Schumacher, and R. N. Schindler, J. Phys. Chem. **79**, 1904 (1975).
[131] E. Schultes, A. A. Christodoulides, and R. N. Schindler, Chem. Phys. **8**, 354 (1975).
[132] G. D. Sides, T. O. Tiernan, and R. J. Hanrahan, J. Chem. Phys. **65**, 1966 (1976).
[133] J. C. Han, M. Suto, J. C. Lee, and Z. Lj. Petrović, J. Appl. Phys. **68**, 2649 (1990).
[134] G. A. Hebner, J. Vac. Sci. Technol. A **14**, 2158 (1996).
[135] R. Siegel, Y. Wang, L. G. Christophorou, and J. K. Olthoff (to be published).
[136] Y. Ota and Y. Uchida, Jpn. J. Phys. **5**, 53 (1928).
[137] A. Elliott and W. H. B. Cameron, Proc. R. Soc. London, Ser. A **158**, 681 (1937).
[138] A. Elliott and W. H. B. Cameron, Proc. R. Soc. London, Ser. A **164**, 531 (1938).
[139] H. G. Howell, Proc. Phys. Soc. London, Ser. A **66**, 759 (1953).
[140] V. V. Rao and P. T. Rao, Can. J. Phys. **36**, 1557 (1958).
[141] P. B. V. Haranath and P. T. Rao, Indian J. Phys. **32**, 401 (1958).
[142] F. P. Huberman, J. Mol. Spectrosc. **20**, 29 (1966).
[143] R. S. Berry and C. W. Reimann, J. Chem. Phys. **38**, 1540 (1963).
[144] G. Pietsch and L. Kehder, Z. Naturforsch. A **22a**, 2127 (1967).
[145] G. Mück and H.-P. Popp, Z. Naturforsch. A **23a**, 1213 (1968).
[146] D. M. de Leeuw, R. Mooyman, and C. A. de Lange, Chem. Phys. Lett. **54**, 231 (1978).
[147] K. Kimura, T. Yamazaki, and Y. Achiba, Chem. Phys. Lett. **58**, 104 (1978).
[148] R. E. Huffman, J. C. Larrabee, and Y. Tanaka, J. Chem. Phys. **47**, 856 (1967).
[149] R. E. Huffman, J. C. Larrabee, and Y. Tanaka, J. Chem. Phys. **48**, 3835 (1968).
[150] C. A. de Lange, P. Van der Meulen, and W. J. Van der Meer, J. Mol. Struct. **173**, 215 (1988).
[151] J. A. R. Samson, Y. Shefer, and G. C. Angel, Phys. Rev. Lett. **56**, 2020 (1986).
[152] D. C. Griffin, M. S. Pindzola, T. W. Gorczyca, and N. R. Badnell, Phys Rev. A **51**, 2265 (1995).
[153] H. P. Saha, Phys. Rev. A **53**, 1553 (1996).
[154] E. J. Robinson and S. Geltman, Phys. Rev. **153**, 4 (1967).
[155] I. I. Fabrikant, J. Phys. B **27**, 4545 (1994).
[156] P. S. Ganas, J. Appl. Phys. **63**, 277 (1988).
[157] T. R. Hayes, R. C. Wetzel, and R. S. Freund, Phys. Rev. A **35**, 578 (1987).
[158] W. Lotz, Z. Phys. **216**, 241 (1968).
[159] Gopaljee, S. N. Chatterjee, and B. N. Roy, Pramara, J. Phys. **36**, 325 (1991).
[160] M. A. Lennon, K. L. Bell, H. B. Gilbody, J. G. Hughes, A. E. Kingston, M. J. Murray, and F. J. Smith, J. Phys. Chem. Ref. Data **17**, 1285 (1988).
[161] D. E. Rothe, Phys. Rev. **177**, 93 (1969).
[162] G. Pietsch, Z. Naturforsch. A **27a**, 989 (1972).
[163] H.-P. Popp, Phys. Rep. **16**, 169 (1975).
[164] L. G. Christophorou, *Atomic and Molecular Radiation Physics* (Wiley-Interscience, New York, 1971), Chap. 7.
[165] I. Yamada, A. Danjo, T. Hirayama, A. Matsumoto, S. Ohtani, H. Suzuki, T. Takayanaki, H. Tawara, K. Wakiya, and M. Yoshino, J. Phys. Soc. Jpn. **58**, 3151 (1989).
[166] W. Shi, D. Fang, F. Lu, H. Gao, J. Gu, S. Wu, W. Wu, J. Tang, and F. Yang, Chin. Phys. Lett. **11**, 73 (1994).
[167] N. Djurić, E. W. Bell, E. Daniel, and G. H. Dunn, Phys. Rev. A **46**, 270 (1992).
[168] D. W. Mueller, T. J. Morgan, G. H. Dunn, D. C. Gregory, and D. H. Crandall, Phys. Rev. **31**, 2905 (1985).
[169] R. Rackwitz, D. Feldmann, E. Heinicke, and H. J. Kaiser, Z. Naturforsch. A **29a**, 1797 (1974).
[170] O. I. Asubiojo, H. L. McPeters, W. N. Olmstead, and J. I. Brauman, Chem. Phys. Lett. **48**, 127 (1977).
[171] S. A. Sullivan, B. S. Freiser, and J. L. Beauchamp, Chem. Phys. Lett. **48**, 294 (1977).
[172] R. L. Champion and L. D. Doverspike, in *Electron-Molecule Interactions and Their Applications*, edited by L. G. Christophorou (Academic, New York, 1984), Vol. 1, p. 619.
[173] L. G. Christophorou, Contrib. Plasma Phys. **27**, 237 (1987).
[174] R. S. Berry, C. W. Reimann, and G. N. Spokes, J. Chem. Phys. **37**, 2278 (1962).

[175]A. Mandl, Phys. Rev. A **14**, 345 (1976).
[176]W. C. Wang and L. C. Lee, J. Phys. D **21**, 675 (1988).
[177]W. C. Wang and L. C. Lee, IEEE Trans. Plasma Science **PS-15**, 460 (1987).
[178]J. W. Cooper and J. B. Martin, Phys. Rev. **126**, 1482 (1962).
[179]Yu. V. Moskvin, High Temp. **3**, 765 (1965).
[80]V. Radojević, H. P. Kelly, and W. R. Johnson, Phys. Rev. A **35**, 2117 (1987).
[81]C. J. Howard, F. C. Fehsenfeld, and M. McFarland, J. Chem. Phys. **60**, 5086 (1974).
[82]F. C. Fehsenfeld, C. J. Howard, and E. E. Ferguson, J. Chem. Phys. **58**, 5841 (1973).
[83]F. C. Fehsenfeld, J. Chem. Phys. **54**, 438 (1971).
[84]F. C. Fehsenfeld, E. E. Ferguson, and A. L. Schmeltekopf, J. Chem. Phys. **45**, 1844 (1966).
[35]L. M. Babcock and G. E. Streit, J. Chem. Phys. **76**, 2407 (1982).
[36]L. D. Doverspike, B. T. Smith, and R. L. Champion, Phys. Rev. A **22**, 393 (1980).
[187]Iu. F. Bydin and V. M. Dukel'skii, Sov. Phys.-JETP **4**, 474 (1957).
[188]M. S. Huq, D. Scott, N. R. White, R. L. Champion, and L. D. Doverspike, J. Chem. Phys. **80**, 3651 (1984).
[189]J. B. Hasted and R. A. Smith, Proc. R. Soc. London Ser. A **235**, 349 (1956).
[190]M. J. Church and D. Smith, J. Phys. D **11**, 2199 (1978)
[191]R. K. Boyd and G. Burns, J. Phys. Chem. **83**, 88 (1979).
[192]Y.-T. Chang, C.-J. Hwang, and T.-M. Su, Chem. Phys. Lett. **114**, 92 (1985).
[193]C.-J. Hwang, R.-C. Jiang, and T.-M. Su, J. Chem. Phys. **84**, 5095 (1986).
[194]F. L. Eisele, M. G. Thackston, W. M. Pope, H. W. Ellis, and E. W. McDaniel, J. Chem. Phys. **70**, 5918 (1979).
[195]L. A. Vieland and E. A. Mason, J. Chem. Phys. **63**, 2913 (1975).
[196]H. R. Skullerud, J. Phys. B **9**, 535 (1976).
[197]E. A. Mason and E. W. McDaniel, *Transport Properties of Ions in Gases* (Wiley, New York, 1988).
[198]M. G. Thackston, M. S. Byers, F. B. Holleman, R. D. Chelf, J. R. Twist, and E. W. McDaniel, J. Chem. Phys. **78**, 4781 (1983).