

# Emerging Disinfection By-Products of Toxicological Interest: Results of a Nationwide Occurrence Study

Susan D. Richardson, Alfred D. Thruston, Jr., Stuart W. Krasner, Howard S. Weinberg, Russell Chinri, Michael J. Scrimshaw, Salvador Pastor, and Gretchen D. Onstad<sup>3</sup>  
<sup>1</sup>Office of Research and Development, National Exposure Research Laboratory, Ecosystems Research Division, Athens, Georgia  
<sup>2</sup>Metropolitan Water District of Southern California, La Verne, California  
<sup>3</sup>Department of Environmental Sciences & Engineering, University of North Carolina, Chapel Hill, North Carolina

Year of Water:  
Thirty Years of Progress  
Through Partnering

## Environmental Issue

The Safe Drinking Water Act and Amendments require that EPA address disinfection by-products (DBPs) in drinking water. DBPs are formed when a disinfectant (such as chlorine) reacts with organic matter and/or bromide naturally present in source waters. Drinking water disinfection by-products (DBPs) are of concern because epidemiologic studies indicate that some may be associated with cancer and adverse reproductive/developmental effects in human populations, and other studies have shown that certain DBPs cause cancer and adverse reproductive/developmental effects in laboratory animals. A few DBPs are regulated; however, most DBPs have not been tested for adverse health effects due to high costs involved. In order to prioritize new DBPs for health effects testing, we initiated a Nationwide Occurrence Study to quantify 'high priority' DBPs that were selected from an extensive prioritization effort of all DBPs that had ever been reported. DBPs were prioritized according to predicted adverse health effects (cancer) by a multidisciplinary group of experts, including toxicologists, structure-activity specialists, and chemists (1). The fate and transport of these DBPs in the distribution system was also studied, and new DBPs were identified. Scientists from the University of North Carolina and the Metropolitan Water District of Southern California collaborated with NERL scientists on this effort.

## Scientific Approach

Drinking waters were chosen across the United States in locations to provide waters with low and high bromide, different pH conditions, and different organic matter levels. Regulated and Information Collection Rule DBPs were also measured for comparison purposes. Samples were collected quarterly from the 12 treatment plants studied. Analytical methods developed for quantifying the high priority DBPs in drinking water included, and mass spectrometry methods were used to identify new DBPs. Analytical techniques that were used to measure these high priority DBPs include methylation with gas chromatography (GC)-electron capture detection (ECD) for the MX analogs and haloacids, pentafluorobenzylhydroxylamine (PFBHA) derivatization with GC-ECD for carbonyl compounds, liquid-liquid extraction-GC-ECD for haloamides and haloacetates; and liquid-liquid extraction-GC-ECD, solid phase extraction (SPE)-GC/mass spectrometry (MS), and purge-and-trap-GC/MS for halonitromethanes, iodo-TTHMs, other halomethanes, haloaldehydes, halo ketones, and haloacetamides. GC with low and high resolution MS was used to identify new DBPs.

## Results

### Halonitromethanes\*

- Individual halonitromethanes 0.1 to 3 ppb
- Dichloro-, bromochloro-, bromodichloro-, and dibromochloronitromethane most prevalent forms observed
- In some cases, pre-ozonation increased formation of brominated trihalonitromethanes (including tribromonitromethane (bromopicrin))

\*In mammalian cell assays, bromonitromethanes have been shown to be at least an order of magnitude more genotoxic to mammalian cells than MX and have genotoxicities greater than all of the haloacetic acid DBPs regulated in the United States, except for monobromoacetic acid (2).

### Halofuranones (MX analogs)

- Widely observed
- Often >100 ng/L
- Max 310 ng/L (plant with high TOC, chlorine dioxide-chlorine-chloramine treatment; not detected until after chlorine/chloramine treatment)
- BMXs identified (170 & 200 ng/L for BMX-1 and BMX-3 at one location with high bromide)

### Haloacetaldehydes

- Haloacetaldehydes 3rd largest fraction (by wt) of halo-DBPs
- Mostly due to dichloroacetaldehyde (in addition to chloral hydrate)
- Dichloroacetaldehyde detected at low ppb to 16 ppb; highest at plant using chloramines & ozone

### New Iodo-Acids\*

- Iodoacetic acid
- Iodobromoacetic acid
- 3-Iodo-3-bromopropenoic acid (2 isomers)
- 2-Iodo-3-methylbutenedioic acid

\*In toxicologic studies of iodoacetic acid (mammalian cell and *in vivo* developmental assays), iodoacetic acid caused developmental effects in mouse embryos (neutral tube closures) at low M levels (1) Iodoacetic acid was also approximately 3-fold more genotoxic and cytotoxic than bromoacetic acid to mammalian cells (4).

### Iodo-Trihalomethanes (TTHMs)

- Individual iodo-TTHMs 0.2 to 15 ppb
- Total iodo-TTHMs (when detected): 2-81% of TTHM4 (sum of 4 regulated TTHMs)
- Highest at plant using chloramines (and no pre-chlorine)
- Dichloroiodomethane most common

### Haloacids

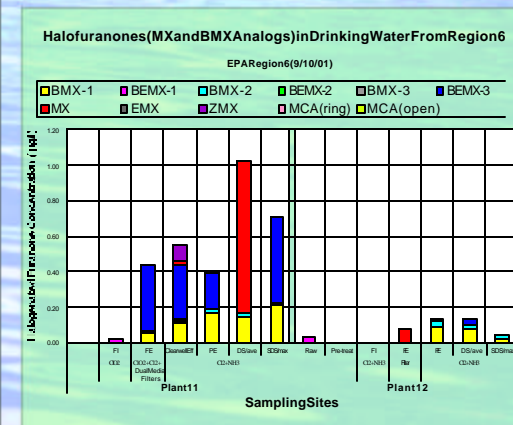
- 3,3-Dichloropropenoic acid high priority DBP measured
- Found in drinking water from most plants sampled
- 0.1 to 5.5 ppb
- Therefore, higher chain halo-acids (not just haloacetic acids) likely significant in drinking water

### Haloamides

- New class not previously measured
- Found in most plants sampled
- Individual haloamides ranged from 0.1 to 9.4 ppb
- Level similar to other commonly measured DBPs

### Distribution System (DS) & Simulated DS Tests

- Chloramines: Most DBPs stable (except for high pH waters)
- Free chlorine: TTHMs, haloacetic acids (HAAs) increase
- Haloacetamides, halonitromethanes, haloacetaldehydes: Typically stable (none of these systems with free chlorine were at high pH levels)
- Halo ketones: Some degraded
- BMXs: Mostly stable
- MXs: Sometimes stable, sometimes degraded - but never completely degraded



## Important Findings

- Many of the high priority DBPs found in drinking waters across the U.S.
- MX levels exceeded previous levels in limited studies conducted to date (300-400 ng/L found); BMXs reached 100-200 ng/L
- Although the use of alternative disinfectants minimized the formation of TTHM4, certain other DBPs formed at significant concentrations:
  - Iodo-TTHMs highest at plant using chloramines
  - Bromo-trihalonitromethanes highest at plant using pre-ozonation
  - Dichloroacetaldehyde highest at plants using chloramines & ozone
  - MX and BMXs highest at plants using chlorinated dioxide (followed by chlorine-chloramines) that treated waters high in NOM and bromide; chlorinated dioxide did not remove MX precursors
- The presence of bromide resulted in a shift in speciation for TTHMs, HAAs, and other classes of DBPs
- Iodo-acids identified for the first time (also new bromo-acids)

## Impact of this Study

This research expands our knowledge on the occurrence of DBPs beyond those that are currently regulated, will help to prioritize future DBP health effects research, and will allow EPA's Office of Water to make improved decisions regarding the safety of drinking water and ultimately minimize any that are found to be hazardous.

## References

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Table 1. High priority DBPs included in Nationwide DBP Occurrence Study

MX and MX-Analogs:	Haloaldehydes:	Haloacetates:	Haloamides:	Non-Halogenated Aldehydes and Ketones:	Volatiles organic compounds (VOCs) and Miscellaneous DBPs:
3-Chloro-4-(dichloromethyl)-5-hydroxy-2(5H)-furanone (MX)	Chloroacetaldehyde	Bromochloromethylacetate	Monobromoacetamide <sup>e</sup>	2-Hexenal	1,1,1,2-Tetrabromo-2-chloroethane
3-Chloro-4-(dichloromethyl)-2-(5H)-furanone (red-MX)	Dichloroacetaldehyde	Bromochloromethylacetate	Dibromoacetamide <sup>e</sup>	5-Keto-1-hexanal	1,1,2,2-Tetrabromo-2-chloroethane
(E)-2-Chloro-3-(dichloromethyl)-butenedioic acid (ox-MX)	Bromochloroacetaldehyde <sup>a</sup>	Bromochloromethylacetate	Trichloroacetamide <sup>e</sup>	Cyanoformaldehyde	Methyl-tert-butylether
(E)-2-Chloro-3-(dichloromethyl)-4-oxobutenoic acid (EMX)	Trichloroacetaldehyde <sup>a</sup>	Bromochloromethylacetate		Methylglyoxal (2,3-butanedione)	Benzylchloride
2,3-Dichloro-4-oxobutenoic acid (Mucrochloric acid)					
3-Chloro-4-(bromochloromethyl)-5-hydroxy-2(5H)-furanone (BMX-1)					
3-Chloro-4-(dibromomethyl)-5-hydroxy-2(5H)-furanone (BMX-2)					
3-Bromo-4-(dibromomethyl)-5-hydroxy-2(5H)-furanone (BMX-3)					
(E)-2-Chloro-3-(bromochloromethyl)-4-oxobutenoic acid (BEMX-1)					
(E)-2-Chloro-3-(dibromomethyl)-4-oxobutenoic acid (BEMX-2)					
(E)-2-Bromo-3-(dibromomethyl)-4-oxobutenoic acid (BEMX-3)					
<b>Haloacids:</b>					
3,3-Dichloropropenoic acid					
<b>Haloacetamides:</b>					
Chloroacetamide					
Bromomethane (methyl bromide) <sup>d</sup>					
Dibromomethane					
Bromochloromethane					
Bromochloroiodomethane					
Dichloroiodomethane					
Dibromoiodomethane <sup>e</sup>					
Chloroiodomethane <sup>e</sup>					
Bromiodomethane <sup>e</sup>					
Iodoform <sup>e</sup>					
Chlorobromomethane					
Carbon tetrachloride					
<b>Halonitromethanes:</b>					
Bromonitromethane					
Chloronitromethane <sup>e</sup>					
Dibromonitromethane					
Dichloronitromethane <sup>e</sup>					
Bromochloronitromethane <sup>e</sup>					
Bromodichloronitromethane <sup>e</sup>					
Dibromodichloronitromethane <sup>e</sup>					
Tribromonitromethane (bromopicrin) <sup>f</sup>					
<b>Haloacetamides:</b>					
Bromoacetamide					
Chloroacetamide					
Tribromoacetamide					
Bromodichloroacetamide					
Dibromochloroacetamide					
<b>Halo ketones:</b>					
Chloropropanone					
1,3-Dichloropropanone					
1,1-Dibromopropanone 1,1,3-Trichloropropanone					
1-Bromo-1,1-dichloropropanone					
1,1,1,3-Tetrachloropropanone					
1,1,3,3-Tetrachloropropanone					
1,1,3,3-Tetrabromopropanone <sup>e</sup>					
1,1,1,3,3-Pentachloropropanone					
Hexachloropropanone					

## Sampling Survey:

- 12 plants sampled quarterly
- 2 plants - same watershed, different treatment/disinfection
- Plant sampled in EPA Regions 3, 4, 5, 6, 7, and 9

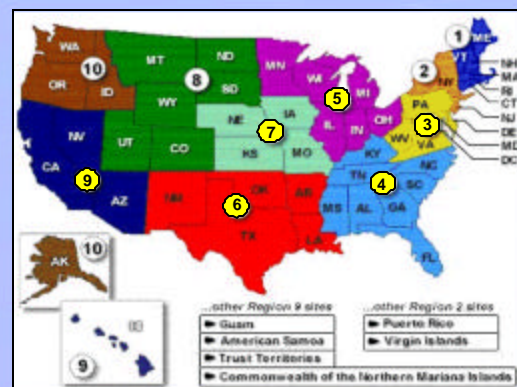


Table 2. Drinking Water Utilities Sampled

Utility (EPA Region)	Disinfection Used
Plant 1 (EPA Region 3)	Ozone-chlorine-chloramines
Plant 2 (EPA Region 3)	Chlorine-chloramines
Plant 3 (EPA Region 6)	Chlorine dioxide-chloramines
Plant 4 (EPA Region 6)	Chlorine dioxide-chloramines
Plant 5 (EPA Region 4)	Chlorine-chloramines
Plant 6 (EPA Region 4)	Chloramines-ozone
Plant 7 (EPA Region 4)	Chlorine dioxide-chloramines
Plant 8 (EPA Region 4)	Ozone-chlorine
Plant 9 (EPA Region 3)	Chlorine-chloramines
Plant 10 (EPA Region 3)	Chlorine
Plant 11 (EPA Region 5)	Chlorine-chloramines
Plant 12 (EPA Region 7)	Chlorine-chloramines

## Structures of MX and Analogs

