Rate Data for Inelastic Collision Processes in the Diatomic Halogen Molecules. 1986 Supplement.

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The previously published compilation of rate data for inelastic collision processes involving the homonuclear and heteronuclear diatomic halogen molecules [J. Phys. Chem. Ref. Data 13, 445 (1984)] has been updated through June, 1986. Additional data on collision processes involving the interhalogens, and on processes at very low kinetic temperatures, are presented; in addition, several previously accepted rate data have been corrected.

Key words: energy transfer; halogens; inelastic collisions; quenching; radiative lifetimes; rotational relaxation; vibrational relaxation.

List of Tables

(Numbering follows that of Ref. 1)

1. Introduction

In 1984, we published a survey[1] of rate data for inelastic collision processes in diatomic halogen molecules, including both homonuclear ($X_2 = Br_2, Cl_2, F_2, I_2$) and heteronuclear (XY = BrCl, BrF, BrI, ClF, ClI, and FI) species. Processes reviewed in the survey included electronic quenching, electronic $\leftrightarrow$ vibrational energy transfer, vibrational relaxation, rotational relaxation, dephasing, depolarization, line broadening, and radiative decay. Theoretical treatments of these processes were also noted. The survey was based on literature published through April, 1983.

During the past several years, sufficient additional data have been published to warrant this supplement (see Tables 1.1 and 1.4–1.10). In particular, much more data are available on the interhalogens, particularly FI, and on collision processes at low relative kinetic energies in supersonic molecular beams. In addition, some previously reported data have been corrected, such as the I^*-Cl_2 reaction rate and the BrI radiative lifetime. Other conclusions based on the original survey, particularly the applicability of angular-momentum based scaling laws,[2] have been borne out by additional measurements.

The supplementary literature references are based on material sent to us by scientists active in the field who have seen the original survey, and on searches of the Molecular Spectroscopy Newsletter published by Physics and Astronomy Departments of the University of California at Berkeley (1983–1986), and the Lockhead Dialog® data base. For further discussion of the methodology, including definitions of collision processes, experimental techniques, and units, please consult Ref. 1.

Table 1.1. Inelastic Collision Data for Bromine

Experimental Data for Bromine

State I	State F	Process	Collision Partner	Temp	Meth	v_i	j_i	v_f	j_f	Quant Rept	Data Entry	Units	Est. Error	Ref
B		quench	Br_2		LIF	7,11,14	<15			k	4.2E-10	cm^3s^{-1}	12%	3
B		quench	He		LIF	7,11,14	<15			k	<2E-12	cm^3s^{-1}	12%	3
B	B	R-T	Br_2		LIF	7,11,14	<15			k	6E-10	cm^3s^{-1}	30%	3
X		E-V?	$N_2^*(A^3\Sigma)$	300 K						k	12E-11	cm^3s^{-1}	17%	4
X	X	V-T	CCl_4(liq.)	298 K	FP(ps)					θ_{vib}	8E-11	s		5

Theoretical Data for Bromine

State I	Process	Collision Partners	Method, Comments	Ref
X	V-T+R-T	H,Li^+	VRI/OSA calc'n, 0.08-1.2 eV relative kinetic energy	6
X	T-V	Li^+	Cross sections for 0 --> 1, 0 --> 2 using BSA	7
X	V-T	Ar,Br	Classical trajectory calc'n @ T=2000-3500 K	8
X	dissoc	Ar,Br	Use results of preceding to calc dissoc rates	9

Radiative Lifetimes for Bromine

State I	State F		Meth	Data (s)	Ref
A'		(CCl_4,298 K)	FP	5.5E-9	5

Table 1.4. Inelastic Collision Data for Bromine Iodide

Experimental Data for Bromine Iodide

State I	State F	Process	Collision Partner	Temp	Meth	v_i	j_i	v_f	j_f	Quant Rept	Data Entry	Units	Est. Error	Ref
D		quench	N_2		MEF					k	5E-10	cm^3s^{-1}	20%	10
D		quench	O_2		MEF					k	2.1E-10	cm^3s^{-1}	15%	10
D		quench	CH_4		MEF					k	9.4E-10	cm^3s^{-1}	11%	10

Radiative Lifetimes for Bromine Iodide

State I	State F	Meth	v_i	j_i	Data (s)	Ref
D	X	MEF			27±4E-9	10
B	X	LIF	2	8-33	0.29-0.07E-6	11 (a)
B	X	LIF	3	4-31	0.72-0.28E-7	11 (a)

(a) Supersedes previous data of Wright & Havey [J. Chem. Phys. 68, 864 (1978)].

Table 1.5. Inelastic Collision Data for Chlorine

Experimental Data for Chlorine

State I	State F	Process	Collision Partner	Temp	Meth	v_i	j_i	v_f	j_f	Quant Rept	Data Entry	Units	Est. Error	Ref
X	X	E-V	I*	300 K	IRF					k	<8E-15	cm^3s^{-1}		12
X		E-V?	N_2*($A^3\Sigma$)	300 K						k	7.8E-11	cm^3s^{-1}	20%	4
X	X	E-V	I*	300 K	IRF					k	2.0E-14	cm^3s^{-1}	5%	13 (a)
X	X	E-V	I*	300 K	IRF					k	1.7E-14	cm^3s^{-1}	30%	14 (a)

(a) Actual rate probably <8E-15 cm^3s^{-1}, fast I*-Cl atom quenching observed.[12]

Table 1.6. Inelastic Collision Data for Chlorine Fluoride

Experimental Data for Chlorine Fluoride

State I	State F	Process	Collision Partner	Temp	Meth	v_i	j_i	v_f	j_f	Quant Rept	Data Entry	Units	Est. Error	Ref
X		(a)	I*	298 K	LIF					k	1.1E-13	cm^3s^{-1}	30%	15

(a) Assumed to be reactive, I* + XY → IX + Y, rather than E-V transfer.

Table 1.7. Inelastic Collision Data for Chlorine Iodide

Experimental Data for Chlorine Iodide

State I	State F	Process	Collision Partner	Temp	Meth	v_i	j_i	v_f	j_f	Quant Rept	Data Entry	Units	Est. Error	Ref
B		quench	He		LIF	1	14			σ	0.005E-16	cm^2	40%	16
B		quench	O_2		LIF	1	14			σ	0.6E-16	cm^2	17%	16
B		quench	Kr		LIF	1	14			σ	0.8E-16	cm^2	25%	16
B		quench	CCl_4		LIF	1	14			σ	4.5E-16	cm^2	12%	16
B		quench	CH_2Cl_2		LIF	1	14			σ	5.5E-16	cm^2	8%	16
B		quench	$CHCl_3$		LIF	1	14			σ	6.3E-16	cm^2	10%	16
B		quench	t-$C_2H_2Cl_2$		LIF	1	14			σ	3.8E-16	cm^2	10%	16
B		quench	g-$C_2H_2Cl_2$		LIF	1	14			σ	6.5E-16	cm^2	10%	16
B		quench	c-$C_2H_2Cl_2$		LIF	1	14			σ	8.2E-16	cm^2	12%	16
B		quench	C_6H_6		LIF	1	14			σ	33E-16	cm^2	17%	16
B		quench	He		LIF	2	15			σ	0.11E-16	cm^2	20%	16
B		quench	O_2		LIF	2	15			σ	8.0E-16	cm^2	12%	16
B		quench	Kr		LIF	2	15			σ	19E-16	cm^2	20%	16
B		quench	CCl_4		LIF	2	15			σ	35E-16	cm^2	17%	16
B		quench	CH_2Cl_2		LIF	2	15			σ	38E-16	cm^2	12%	16
B		quench	$CHCl_3$		LIF	2	15			σ	40E-16	cm^2	12%	16
B		quench	t-$C_2H_2Cl_2$		LIF	2	15			σ	41E-16	cm^2	10%	16
B		quench	g-$C_2H_2Cl_2$		LIF	2	15			σ	51E-16	cm^2	10%	16
B		quench	c-$C_2H_2Cl_2$		LIF	2	15			σ	57E-16	cm^2	10%	16
B		quench	C_6H_6		LIF	2	15			σ	79E-16	cm^2	10%	16
X	X	E-V	I*	300 K	IRF					k	1.5E-11	cm^3s^{-1}	30%	14
X	X	E-V	I*	300 K	IRF					k	3.3E-11	cm^3s^{-1}	12%	13
X		E-V?	N_2*($A^3\Sigma$)	300 K						k	8.0E-11	cm^3s^{-1}	20%	4
B		quench	ICl_3		LIF	3	5-52			k	8.7E-10	cm^3s^{-1}	10%	17

Table 1.7. Inelastic Collision Data for Chlorine Iodide (continued)

Radiative Lifetimes for Chlorine Iodide

State I	State F	Meth	v_i	j_i	Data (s)	Est. Error	Ref
B	X	LIF	3	5-52	0.5-1.0E-9		17
B	X	LIF	1	7-55	4.1E-6	5%	18
B	X	LIF	2	7-54	(3.3 to 0.07)E-6 function of j		18

Table 1.8. Inelastic Collision Data for Fluorine

Experimental Data for Fluorine

State I	State F	Process	Collision Partner	Temp	Meth	v_i	j_i	v_f	j_f	Quant Rept	Data Entry	Units	Est. Error	Ref
X		(a)	I*	298 K	LIF					k	<8.7E-14	$cm^3 s^{-1}$		15

(a) Assumed to be reactive, I* + XY → IX + Y, rather than E-V transfer.

Table 1.9. Inelastic Collision Data for Fluorine Iodide

Experimental Data for Fluorine Iodide

State I	State F	Process	Collision Partner	Temp	Meth	v_i	j_i	v_f	j_f	Quant Rept	Data Entry	Units	Est. Error	Ref
B		quench	He	298 K	LIF	all	v'			k	<1.0E-14	$cm^3 s^{-1}$		19
B		quench	N_2	298 K	LIF	all	v'			k	<1.0E-14	$cm^3 s^{-1}$		19
B		quench	SF_6	298 K	LIF	all	v'			k	<1.0E-14	$cm^3 s^{-1}$		19
B		quench	F_2	298 K	LIF	all	v'			k	4E-12	$cm^3 s^{-1}$	25%	19
B		quench	He	298 K	LIF	3-8				σ	<7.9E-20	cm^2		20
B		quench	Ne	298 K	LIF	3-8				σ	<1.6E-19	cm^2		20
B		quench	Ar	298 K	LIF	3-8				σ	<2.2E-19	cm^2		20
B		quench	Kr	298 K	LIF	3-8				σ	<2.9E-19	cm^2		20
B		quench	Xe	298 K	LIF	3-8				σ	<3.1E-19	cm^2		20
B		quench	F_2	298 K	LIF	3				k	3.4E-12	$cm^3 s^{-1}$	15%	20
B		quench	F_2	298 K	LIF	3				σ	7.4E-17	cm^2	15%	20
B		quench	F_2	298 K	LIF	6				k	4.5E-12	$cm^3 s^{-1}$	15%	20
B		quench	F_2	298 K	LIF	6				σ	9.9E-17	cm^2	15%	20
B		quench	F_2	298 K	LIF	7				k	5.2E-12	$cm^3 s^{-1}$	8%	20
B		quench	F_2	298 K	LIF	7				σ	1.1E-16	cm^2	8%	20
B		quench	I_2	298 K	LIF	3-7				σ	9.2E-15	cm^2		20
B		quench	N_2	298 K	LIF	3-8				σ	<1.9E-19	cm^2		20
B		quench	H_2O	298 K	LIF	3-6				σ	3.8E-15	cm^2	30%	20
B		quench	H_2O	298 K	LIF	4				σ	2.8E-17	cm^2	20%	20
B		quench	H_2O	298 K	LIF	5				σ	6.2E-17	cm^2	25%	20
B		quench	H_2O	298 K	LIF	6				σ	1.2E-16	cm^2	25%	20
B		quench	O_2	298 K	LIF	3-8				σ	1.2E-16	cm^2	20%	20
B		quench	O_2	298 K	LIF	3-8				σ	2.1E-17	cm^2	15%	20
B	B	V-T	He	298 K	LIF	4	3			k	6.9E-12	$cm^3 s^{-1}$	10%	19
B	B	V-T	N_2	298 K	LIF	4	3			k	3.5E-12	$cm^3 s^{-1}$	10%	19
B	B	V-T	He	298 K	LIF	3	2			k	5.4E-12	$cm^3 s^{-1}$	10%	19
B	B	V-T	N_2	298 K	LIF	3	2			k	2.5E-12	$cm^3 s^{-1}$	10%	19
B		V-T	Ar/N_2	300 K		1				k	0.6E-12	$cm^3 s^{-1}$	30%	21
B		V-T	Ar/N_2	300 K		2				k	1.6E-12	$cm^3 s^{-1}$	30%	21
B		V-T	Ar/N_2	300 K		3				k	2.6E-12	$cm^3 s^{-1}$	30%	21
B		V-T	Ar/N_2	300 K		4				k	3.5E-12	$cm^3 s^{-1}$	30%	21
B		V-T	Ar/N_2	300 K		5				k	2.6E-12	$cm^3 s^{-1}$	30%	21
B		V-T	Ar/N_2	300 K		6				k	4.1E-12	$cm^3 s^{-1}$	30%	21
B		V-T	Ar/N_2	300 K		7				k	4.0E-12	$cm^3 s^{-1}$	30%	21
B		V-T	Ar/N_2	300 K		8				k	2.9E-12	$cm^3 s^{-1}$	30%	21
B		quench	I_2†	300 K		3-6				k	3.5E-10	$cm^3 s^{-1}$	6%	22
X	B	V-E	HF†	900-2000 K	CL					k	3.6E-14	$cm^3 s^{-1}$		23
X	B	E-E	N_2*	300 K				3-6		k	2.0E-10	$cm^3 s^{-1}$		21
X		(a)	I*	400 K	AA					k	1.3E-11	$cm^3 s^{-1}$	50%	15

(a) Process assumed to reactive, I* + XY → IX + Y, rather than E-V transfer.

Table 1.10. Inelastic Collision Data for Iodine

Experimental Data for Iodine

State I	State F	Process	Collision Partner	Temp	Meth	v_i	j_i	v_f	j_f	Quant Rept	Data Entry	Units	Est. Error	Ref
B		quench	Kr		LIF	8-49				σ	5-12E-16	cm^2	15%	24
B		quench	t-$C_2H_2Cl_2$		LIF	8-49				σ	30-60E-16	cm^2	10%	24
B	B	V-T	He	19 K	LIF	43	12+16	44 to 33		k	[equation]	$cm^3 s^{-1}$		25 (a)
B	B	V-T	H_2	30 K	LIF	43	12+16	44 to 33		k	[equation]	$cm^3 s^{-1}$		25 (a)
X		E-V?	$N_2^*(A^3\Sigma)$	300 K						k	23E-11	$cm^3 s^{-1}$	20%	4
X	D'	E-E	$N_2^*(A^3\Sigma)$	300 K						k	4E-14	$cm^3 s^{-1}$		4
X	X	E-V	I*	300 K	IRF					k	3.0E-11	$cm^3 s^{-1}$	3%	13
X	A'?	E-E	O_2*(benzene solution)	300 K	FP					k	2.3E-18	$cm^3 s^{-1}$	15%	26
X	X	V-T	He	E_T=0-500 meV	BS	0		1,2,3		σ	[graph]	cm^2		27 (b)
X	X	V-T	H_2,D_2	E_T=0-500 meV	BS	0		1,2,3		σ	[graph]	cm^2		28 (b)
B	B	V-T+R-T	He	300 K	LIF	13	41	12	39	k	526E-14	$cm^3 s^{-1}$	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				41	k	556E-14	$cm^3 s^{-1}$	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				43	k	526E-14	$cm^3 s^{-1}$	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				45	k	512E-14	$cm^3 s^{-1}$	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				47	k	470E-14	$cm^3 s^{-1}$	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				49	k	439E-14	$cm^3 s^{-1}$	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				51	k	400E-14	$cm^3 s^{-1}$	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				53	k	357E-14	$cm^3 s^{-1}$	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				55	k	319E-14	$cm^3 s^{-1}$	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				57	k	289E-14	$cm^3 s^{-1}$	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				59	k	258E-14	$cm^3 s^{-1}$	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				61	k	211E-14	$cm^3 s^{-1}$	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				63	k	186E-14	$cm^3 s^{-1}$	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				65	k	170E-14	$cm^3 s^{-1}$	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				67	k	139E-14	$cm^3 s^{-1}$	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				69	k	115E-14	$cm^3 s^{-1}$	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				71	k	95E-14	$cm^3 s^{-1}$	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				73	k	74E-14	$cm^3 s^{-1}$	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				75	k	61E-14	$cm^3 s^{-1}$	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				77	k	45E-14	$cm^3 s^{-1}$	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				79	k	37E-14	$cm^3 s^{-1}$	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				81	k	31E-14	$cm^3 s^{-1}$	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				83	k	31E-14	$cm^3 s^{-1}$	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				85	k	21E-14	$cm^3 s^{-1}$	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				87	k	23E-14	$cm^3 s^{-1}$	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				89	k	18E-14	$cm^3 s^{-1}$	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				91	k	12E-14	$cm^3 s^{-1}$	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				93	k	14E-14	$cm^3 s^{-1}$	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				95	k	11E-14	$cm^3 s^{-1}$	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				97	k	4E-14	$cm^3 s^{-1}$	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				99	k	9E-14	$cm^3 s^{-1}$	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				101	k	8E-14	$cm^3 s^{-1}$	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				103	k	7E-14	$cm^3 s^{-1}$	10-20%	29

(a) For He-$I_2^*(v)$,

$$k(v,v') = (1.70 \pm 0.05\ cm^3 s^{-1}) \times (0.0065)^{|\Delta v-1|}\ v(v-1)(v-2)\ldots(v+\Delta v+1)$$

For H_2-$I_2^*(v)$,

$$k(v,v') = (3.2 \pm 0.1\ cm^3 s^{-1}) \times (0.009)^{|\Delta v-1|}\ v(v-1)(v-2)\ldots(v+\Delta v+1)$$

(b) Cross section σ varies between 0.0 and 0.02E-16 cm^2, energy- and v_f- dependent.

Table 1.10. Inelastic Collision Data for Iodine (continued)

Experimental Data for Iodine

State I	State F	Process	Collision Partner	Temp	Meth	v_i	j_i	v_f	j_f	Quant Rept	Data Entry	Units	Est. Error	Ref
B	B	V-T+R-T	He	300 K	LIF	13	91	12	49	k	78E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				53	k	83E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				55	k	60E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				57	k	87E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				59	k	72E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				61	k	106E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				63	k	103E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				73	k	203E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				75	k	237E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				77	k	281E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				79	k	297E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				81	k	324E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				83	k	340E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				85	k	391E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				87	k	416E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				89	k	421E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				91	k	464E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				93	k	472E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				95	k	451E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				99	k	409E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				101	k	357E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				105	k	307E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				107	k	292E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	He	300 K	LIF				109	k	277E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	Xe	300 K	LIF	13	41	12	39	k	107E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	Xe	300 K	LIF				41	k	109E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	Xe	300 K	LIF				47	k	115E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	Xe	300 K	LIF				49	k	104E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	Xe	300 K	LIF				51	k	112E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	Xe	300 K	LIF				53	k	113E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	Xe	300 K	LIF				55	k	113E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	Xe	300 K	LIF				57	k	120E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	Xe	300 K	LIF				59	k	119E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	Xe	300 K	LIF				61	k	118E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	Xe	300 K	LIF				63	k	116E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	Xe	300 K	LIF				65	k	120E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	Xe	300 K	LIF				67	k	118E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	Xe	300 K	LIF				73	k	113E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	Xe	300 K	LIF				75	k	119E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	Xe	300 K	LIF				77	k	110E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	Xe	300 K	LIF				79	k	103E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	Xe	300 K	LIF				81	k	110E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	Xe	300 K	LIF				83	k	108E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	Xe	300 K	LIF				85	k	95E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	Xe	300 K	LIF				87	k	108E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	Xe	300 K	LIF				89	k	88E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	Xe	300 K	LIF				91	k	94E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	Xe	300 K	LIF				93	k	81E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	Xe	300 K	LIF				95	k	88E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	Xe	300 K	LIF				97	k	84E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	Xe	300 K	LIF				99	k	86E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	Xe	300 K	LIF				101	k	77E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	Xe	300 K	LIF	13	91	12	51	k	50E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	Xe	300 K	LIF				53	k	60E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	Xe	300 K	LIF				55	k	59E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	Xe	300 K	LIF				57	k	69E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	Xe	300 K	LIF				59	k	65E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	Xe	300 K	LIF				61	k	71E-14	cm^3s^{-1}	10-20%	29

Table 1.10. Inelastic Collision Data for Iodine (continued)

Experimental Data for Iodine

State I	State F	Process	Collision Partner	Temp	Meth	v_i	j_i	v_f	j_f	Quant Rept	Data Entry	Units	Est. Error	Ref
B	B	V-T+R-T	Xe	300 K	LIF	13	91	12	63	k	67E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	Xe	300 K	LIF				75	k	78E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	Xe	300 K	LIF				77	k	83E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	Xe	300 K	LIF				79	k	94E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	Xe	300 K	LIF				81	k	94E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	Xe	300 K	LIF				83	k	66E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	Xe	300 K	LIF				85	k	90E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	Xe	300 K	LIF				87	k	85E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	Xe	300 K	LIF				89	k	75E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	Xe	300 K	LIF				91	k	98E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	Xe	300 K	LIF				93	k	99E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	Xe	300 K	LIF				95	k	91E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	Xe	300 K	LIF				99	k	72E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	Xe	300 K	LIF				101	k	100E-14	cm^3s^{-1}	10-20%	29
B	B	V-T+R-T	Xe	300 K	LIF				105	k	96E-14	cm^3s^{-1}	10-20%	29
B	B	V-T	He(?)	300 K?		15	14			σ	3.1E-16	cm^2	10%	30
B	B	V-T	He(?)	300 K?		15	13			σ	0.78E-16	cm^2	10%	30
B	B	dephase	I_2	300 K?	CT	[589.7nm]				σ	590E-16	cm^2	15%	31
B	B	e-loss	I_2	300 K?	CT	[589.7nm]				σ	227E-16	cm^2	10%	31
X		dissoc	MgO(100)	548 K	BS(TOF)									32 (c)
A'		quench	He	300 K	FP					k	9.4E-15	cm^3s^{-1}	10%	33 (d)
A'		quench	Ar	300 K	FP					k	2.8E-14	cm^3s^{-1}	35%	33 (d)
A'		quench	N_2	300 K	FP					k	3.5E-14	cm^3s^{-1}	30%	33 (d)
A'		quench	SF_6	300 K	FP					k	2.4E-13	cm^3s^{-1}	5%	33 (d)
A'		quench	I_2	300 K	FP					k	5.5E-11	cm^3s^{-1}	15%	33 (d)
B		quench	He	298 K	CT	2	59	15	60	σ	1.16E-16	cm^2	6%	34,35,36
B		quench	Ne	298 K	CT	2	59	15	60	σ	3.58E-16	cm^2	6%	34,35,36
B		quench	Ar	298 K	CT	2	59	15	60	σ	10.4E-16	cm^2	5%	34,35,36
B		quench	Kr	298 K	CT	2	59	15	60	σ	22.2E-16	cm^2	2%	34,35,36
B		quench	Xe	298 K	CT	2	59	15	60	σ	48.7E-16	cm^2	2%	34,35,36
B		quench	I_2	298 K	CT	2	59	15	60	σ	115E-16	cm^2	20%	34,35,36
B		V-T+R-T	He	298 K	CT	2	59	15	60	σ	54.3E-16	cm^2	2%	34,35,36
B		V-T+R-T	Ne	298 K	CT	2	59	15	60	σ	89.5E-16	cm^2	3%	34,35,36
B		V-T+R-T	Ar	298 K	CT	2	59	15	60	σ	136E-16	cm^2	3%	34,35,36
B		V-T+R-T	Kr	298 K	CT	2	59	15	60	σ	154E-16	cm^2	2%	34,35,36
B		V-T+R-T	Xe	298 K	CT	2	59	15	60	σ	150E-16	cm^2	2%	34,35,36
B		V-T+R-T	I_2	298 K	CT	2	59	15	60	σ	110E-16	cm^2	40%	34,35,36
B		dephas	He	298 K	CT	2	59	15	60	σ	66E-16	cm^2	10%	34,35,36
B		dephas	Ne	298 K	CT	2	59	15	60	σ	110E-16	cm^2	10%	34,35,36
B		dephas	Ar	298 K	CT	2	59	15	60	σ	160E-16	cm^2	10%	34,35,36
B		dephas	Kr	298 K	CT	2	59	15	60	σ	207E-16	cm^2	10%	34,35,36
B		dephas	Xe	298 K	CT	2	59	15	60	σ	270E-16	cm^2	10%	34,35,36
B		quench	I_2	243-273 K	LIF	14	0-8			σ	190E-16	cm^2	7%	37
B		quench	H_2	243-273 K	LIF	14	0-8			σ	2.5E-16	cm^2	12%	37
B		quench	CO	243-273 K	LIF	14	0-8			σ	15.1E-16	cm^2	4%	37
B		quench	CH_4	243-273 K	LIF	14	0-8			σ	18.0E-16	cm^2	4%	37
B		quench	He	9.4 K	LIF +SSE	11				σ	0.33E-16	cm^2	25%	38

(c) Dissociation measured on collision of I_2 with heated MgO surface.

(d) Temperature dependence of quenching rate measured, 330 K > T > 280 K.

Table 1.10. Inelastic Collision Data for Iodine (continued)

Theoretical Data for Iodine

State I	Process	Collision Partners	Method, Comments	Ref
B	V-T	He	Resonance analysis, T = 0 to 5 K	39
	V-V	I_2	Quantum calculation	40
B	R-T	He	Quasi-classical, collision energy < 0.5 cm^{-1}	41
X	V-T	He,Ar	Scaling law for vibrational energy transfer	42
B	R-T	He	620 K (formula)	43
B	R-T	Xe	550 K (formula)	43
X	V-V+V-T	I_2	Quantum calculation	44
X	V-V+V-T	I_2	Quantum calculation	45
X	V-T	He	VEDW/IOS calculation	46
X	V-T	He	VCC/IOS calculation, cf. to Hall et al. expts. [27]	47
	V-T	He	Semiclassical calculation at low relative K.E.	48
X	V-T	$(Ne)_n$ clusters n=4,8,16	Theoretical calculation, classical dynamics	49
X	V-T	He	Theoretical calculation at very low temperatures using VCC-RIOS	50

Radiative Lifetimes for Iodine

State I	State F	Meth	v_i	j_i		Ref
B	X	43	12,16		(formula) with Γ_{rad} = 0.314 ± 0.018E+6 s^{-1}	51

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